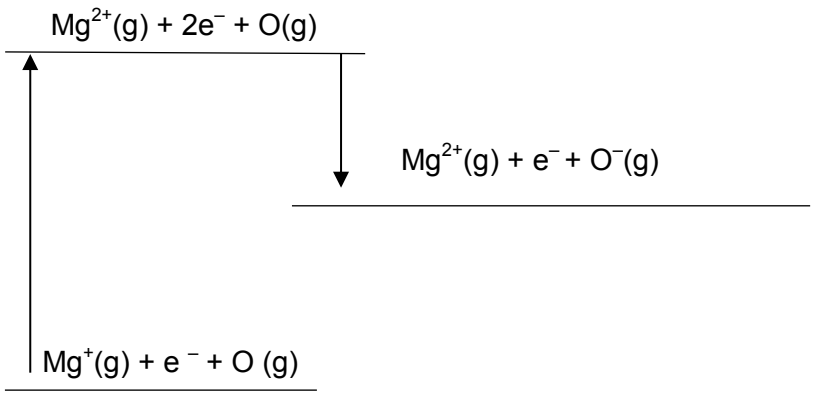
Question	Answers	Additional Comments/Guidance	Mark
01.1		One mark for each level with correct state symbols	1 1 1
01.2	$\Delta_f H = \Delta_a H$ (Mg) + ½ $\Delta_{BD} H$ (O_2) + $\Delta_{1st\ IE} H$ (Mg) + $\Delta_{2nd\ IE} H$ (Mg) + $\Delta_{1st\ EA} H$ (O) + $\Delta_{2nd\ EA} H$(O) + $\Delta_{LE} H$ (MgO) - 602 = 150 + (½ x 496) + 736 +1450 – 142 + 844 + Δ_{LE}H (MgO) $\Delta_{LE} H$ (MgO) = -3888 / -3890 (kJ mol^{-1})	Allow answers to 2sf or more 1 mark for +3888 or +3890 1 mark for -4136 or -4140 (not 496 x 1/2)	1 1 1
Total			6