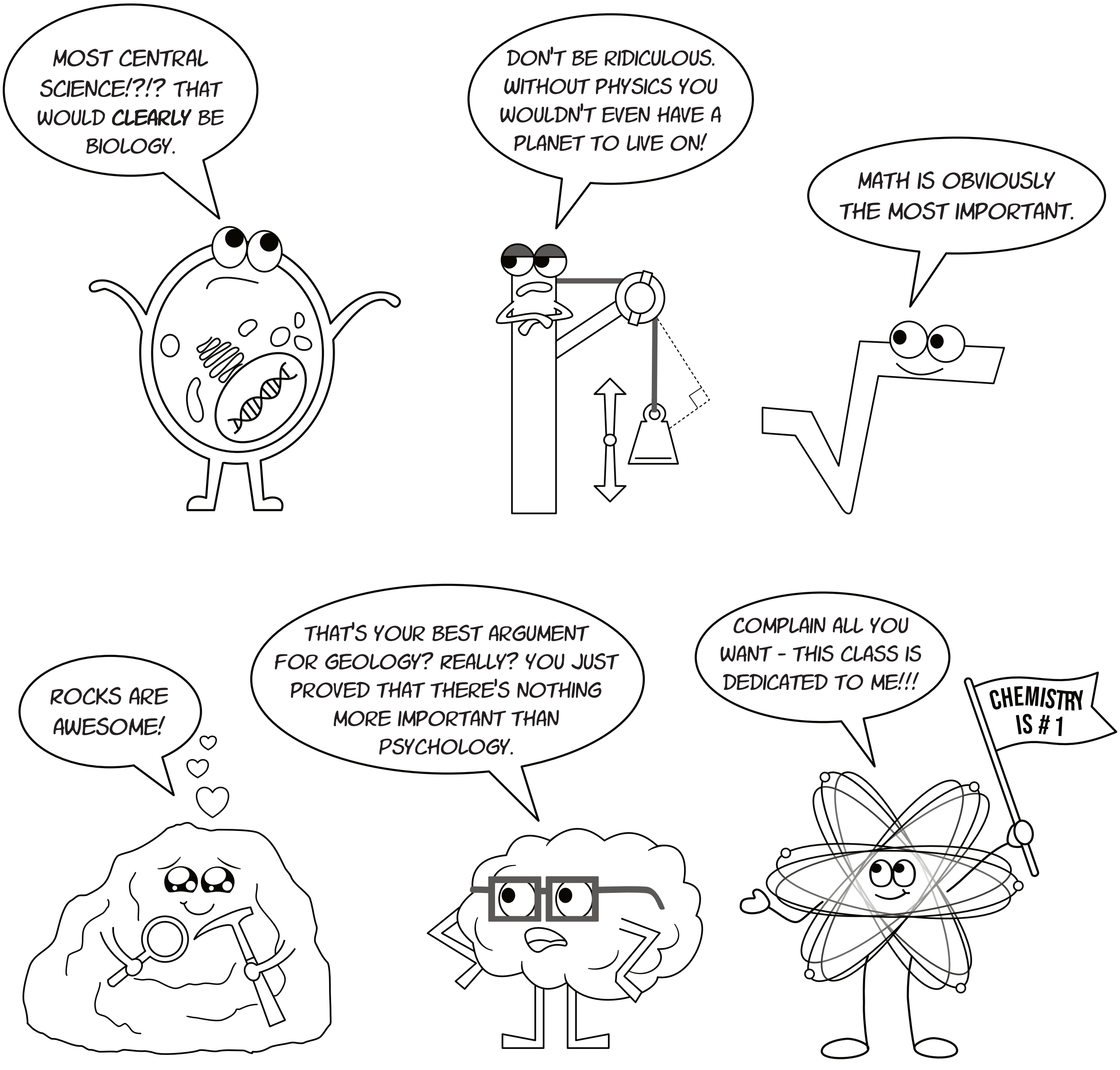
CHEMISTRY

The central and most important branch of science



A course by Science Mom & Math Dad

Chemistry

Lesson	Topic	Pages in the notes
Introduction	Tips for best learning and why chemistry is important!	
1	The story of the atom	1-2
2	Elemental, dear Watson!	3-4
3	Modeling clay orbitals	5-6
4	A noble quest	7-10
5	Why share electrons?	11-12
6	Game show review	
7	Element vs mixture vs compound	15-16
8	What is radioactivity?	17-18
9	Going bananas	
10	Edible experiments	19-20
11	States of matter	21-22
12	Matter batter	
13	Physical reactions	23-24
14	Fizzing experiments	25-26
15	What’s a reaction?	
16	Chemical reactions	27-29
17	More chemical reactions	30
18	Carbon, the building block of life	31-32
19	Toasters and cooking mysteries	33-34
20	Lemon battery	35-36
21	Game show review	
22	Where do fossil fuels come from?	37-39
23	What is fire really?	40
24	Why do leaves change color?	41-42
25	The chemistry of lava	43
26	The chemistry of acids and bases	44

Have questions? Contact jenny@science.mom

You will see boxes like this through out the notes. Use them to draw your favorite moment from class, to write down something cool you learned, or for plain old-fashioned doodling.

YOUR DOODLE SPACE

Supply List for Hands-on Activities:

Lesson 3 - Modeling Clay Orbitals

- Toothpicks
- Modeling clay or play dough (7 different colors)

Lesson 10 - Edible Experiments

- Granulated Sugar (at least 7 cups)
- Kool-aid packets
- Cake pop sticks or string
- A ruler
- 2 pint-size mason jars with lids OR cups and rubber bands
- Coffee filters or paper
- 2 Microwavable popcorn packets

Lesson 14 - Fizzing Experiments

- 6 Alka-Seltzer tablets
- 6 bottles of soda in plastic containers with narrow tops. Any size and type will work, but I recommend 16 oz coke bottles (because Coke is slightly more carbonated than other sodas). You'll use the bottles twice in this experiment and reuse two of them again in the Dec 4th water filtration experiment.
- Baking soda
- 3 packages of Pop Rocks candy
- 6 Balloons (standard 9 inch size)
- A funnel (to help get baking soda inside the balloon)
- Food Coloring
- Vinegar
- Vegetable oil (a whole bottle)
- Safety glasses

Lesson 20 - Lemon or Vinegar Batteries

- Citrus fruit such as lemons OR a potato OR vinegar and an empty ice cube tray
- An LED diode
- Copper penny, wire, or copper sheets
- Galvanized nail or zinc sheets
- Alligator clips
- Scissors or knife

You can get all of these items in a "lemon battery science kit" online for between $5 and $9.

Lesson 30 Frankenseeds

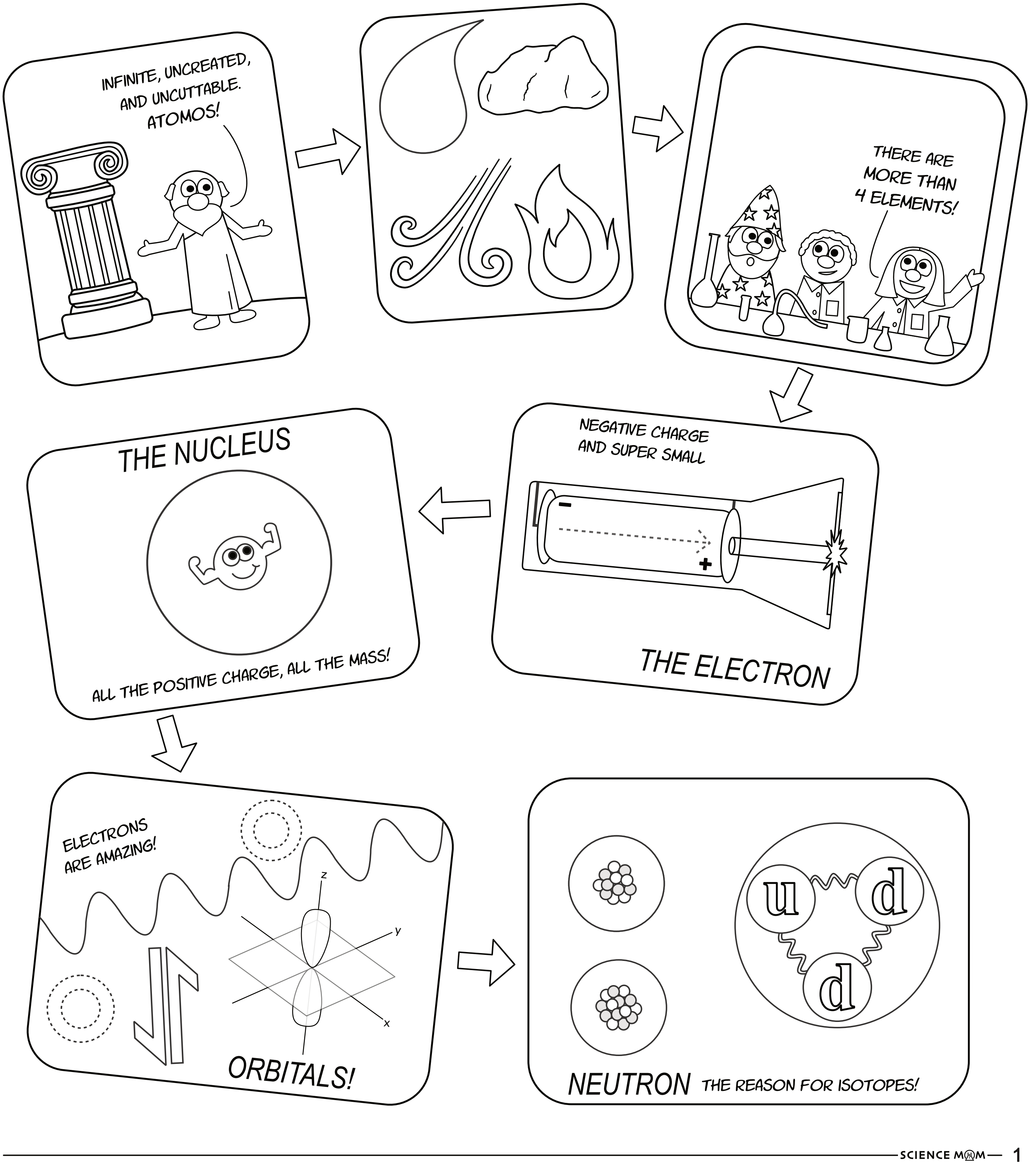
- Cardboard egg carton(s)
- Paper towels
- An empty bread or produce bag
- At least 6 types of seeds from the kitchen (could include rice, beans, lentils, chia seeds, walnuts, sunflower seeds, almonds, peanuts, flax seeds quinoa, or seeds from inside foods like apples, peas, avocados, pears, oranges, kiwis, or cucumbers)

Lesson 43 - DIY Water Filter

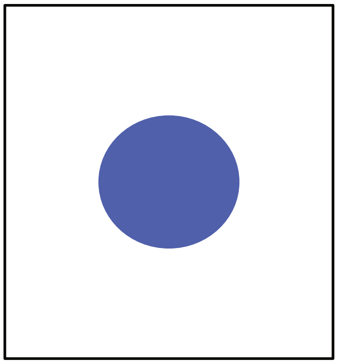
- Two plastic 12 or 16 oz bottles (can reuse the ones from Sept 25)
- Scissors (you might want an adult's help to cut the bottles)
- Sand
- Gravel
- Activated charcoal
- Coffee filters
- A small square of cotton fabric or a couple of cotton balls

The story of the ATOM

WHAT ARE THINGS REALLY MADE OF?



DRAW THE DIFFERENT
MODELS IN THESE BOXES!

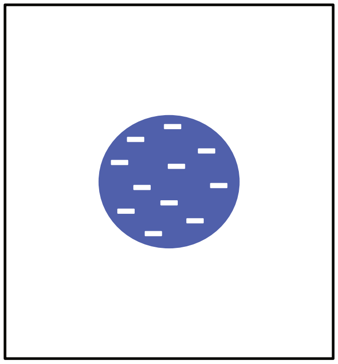


SOLID SPHERE

1803
JOHN DALTON

✓ ELEMENTS ARE MADE OF DIFFERENT ATOMS

⊗ THE ATOM ISN'T THE SMALLEST PARTICLE

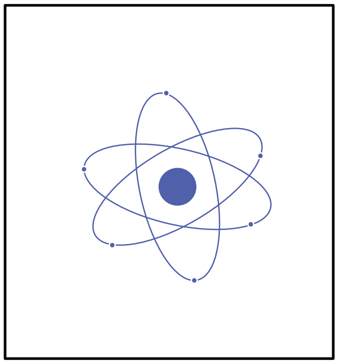


PLUM PUDDING

1904
J.J. THOMPSON

✓ ELECTRONS ARE NEGATIVELY CHARGED

⊗ HAD NO NUCLEUS

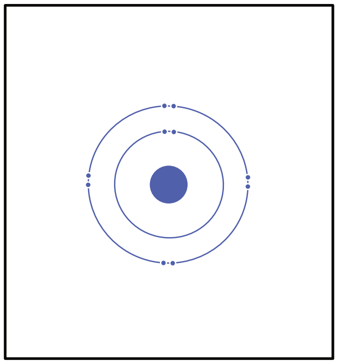


PLANETARY

1911
RUTHERFORD

✓ POSITIVE CHARGE IN THE NUCLEUS

⊗ DIDN'T PREDICT ELECTRON BEHAVIORS



BOHR

1913
NIELS BOHR

✓ ELECTRONS HAVE DISTINCT ENERGY LEVELS

⊗ DID NOT EXPLAIN LARGER ATOMS

YOUR NOTES:

Welcome to the Elemental Cafe

How to order an element:

1. Choose the number of protons*
2. Make it an isotope!
   Adjust the number of neutrons
3. Make it an ion!**
   Adjust the number of electrons

* NUMBERS ABOVE 90 ARE NOT SERVED.

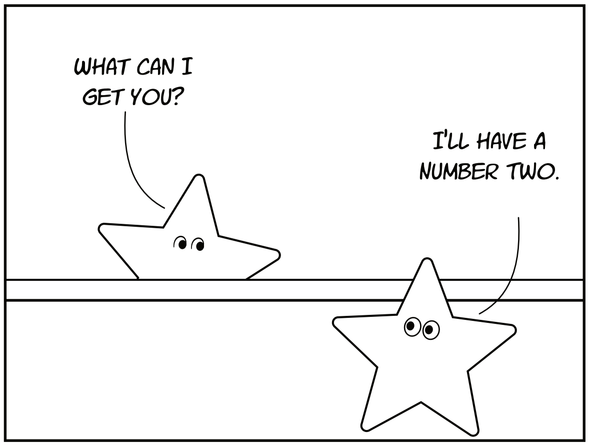
** LIMITED AVAILABILITY.

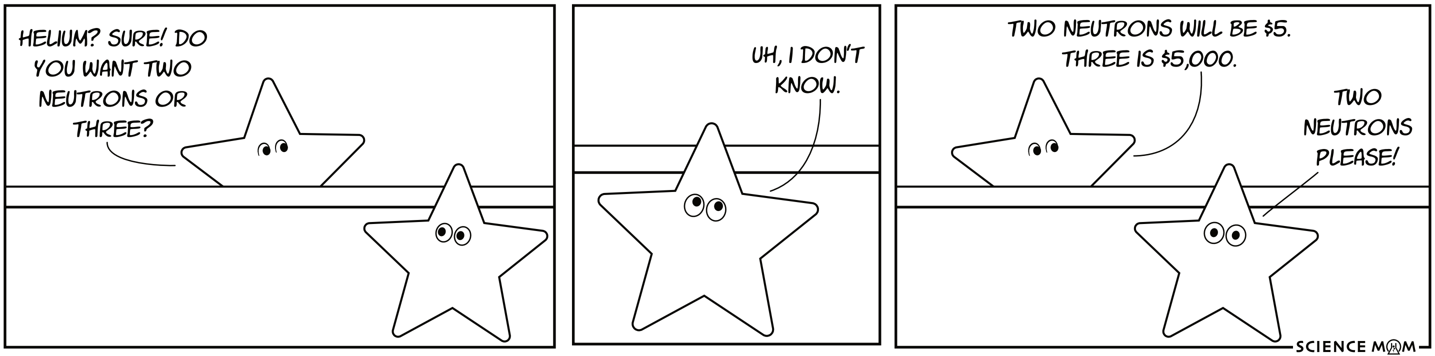
Daily Special

CARBON 14

6 PROTONS
8 NEUTRONS
6 ELECTRONS

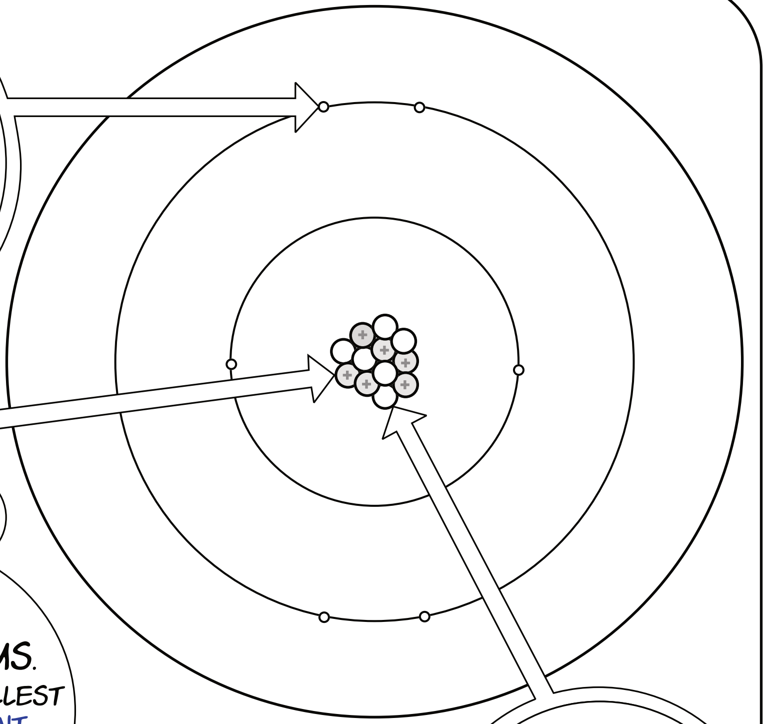
Remarkably stable!





FILL IN THE BLANKS USING THESE WORDS:

element	positive	no	nucleus	orbitals	neutrons
nucleus	protons	nucleus	118	matter	negative



ELECTRONS

HAVE A NEGATIVE CHARGE AND OCCUPY SPACES AROUND THE NUCLEUS KNOWN AS ORBITALS THEY DON'T ADD ANY REAL MASS TO THE ATOM.

PROTONS

HAVE A POSITIVE CHARGE AND EXIST IN THE NUCLEUS OF THE ATOM. THE NUMBER OF PROTONS DETERMINES WHICH ELEMENT THE ATOM IS.

ALL MATTER IS MADE OF **ATOMS**. ATOMS ARE THE SMALLEST PIECE OF AN ELEMENT THAT STILL BEHAVES LIKE THAT ELEMENT.

AN **ELEMENT** CANNOT BE BROKEN INTO SIMPLER SUBSTANCES BY CHEMICAL REACTIONS. THERE ARE 118 KNOWN ELEMENTS.

ATOMS ARE MOSTLY MADE OF EMPTY SPACE. ALL OF THEIR MASS IS IN THE THE NUCLEUS

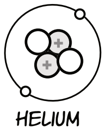
NEUTRONS

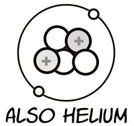
HAVE NO CHARGE AND EXIST IN THE NUCLEUS OF THE ATOM. THEY ARE ABOUT THE SAME SIZE AS PROTONS.

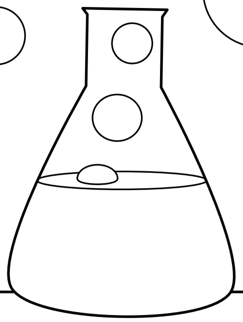
THE **NUCLEUS** OF THE ATOM CONTAINS PROTONS AND NEUTRONS BUT NOT ELECTRONS.

ISOTOPES

ARE VARIATIONS OF THE SAME ELEMENT WITH A DIFFERENT NUMBER OF NEUTRONS.



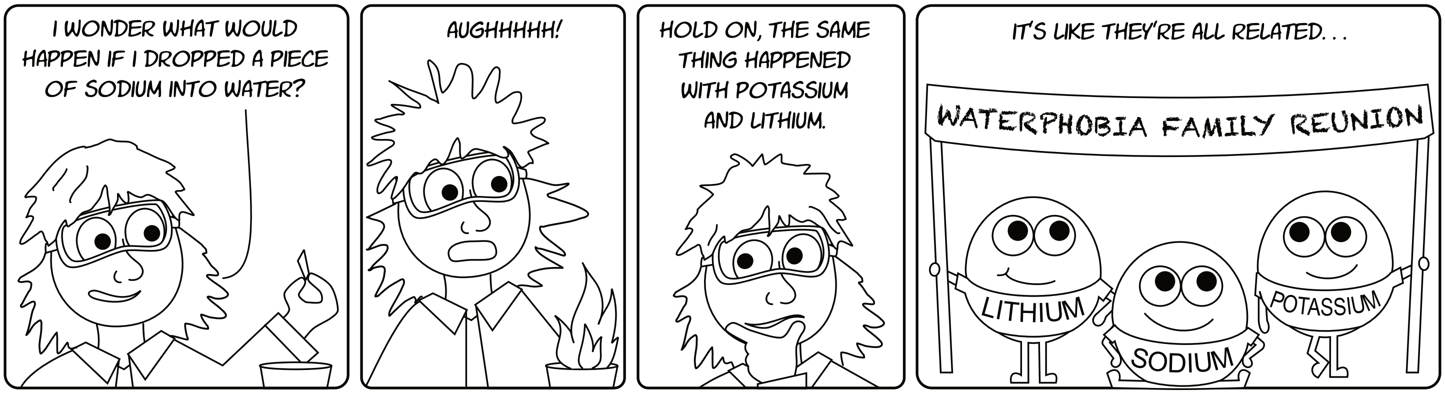


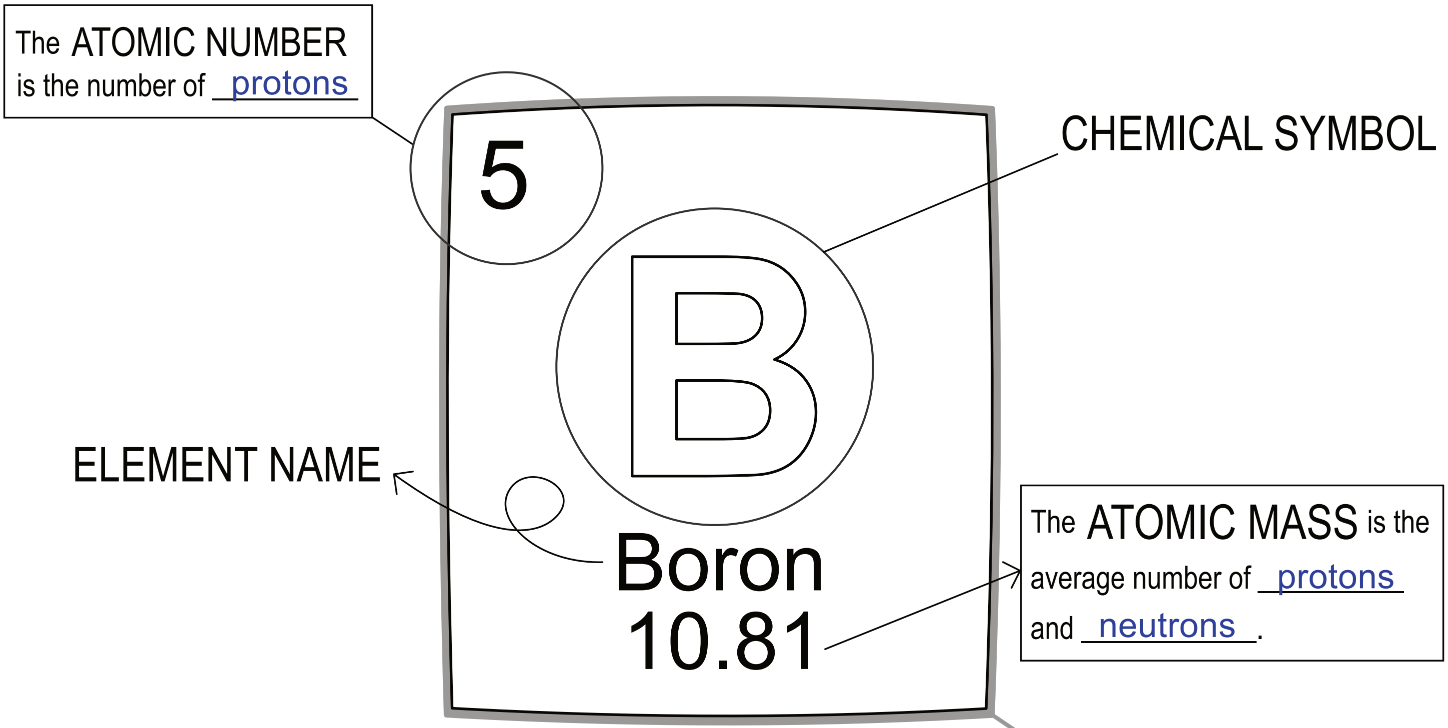


YOUR NOTES:

The PERIODIC table

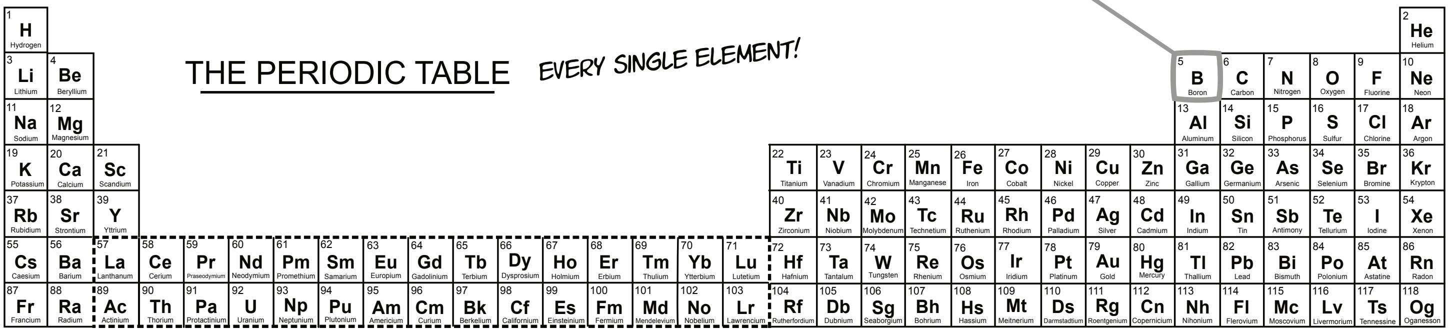
PRETTY MUCH THE COOLEST CHART EVER





The ATOMIC NUMBER is the number of protons

The ATOMIC MASS is the average number of protons and neutrons.



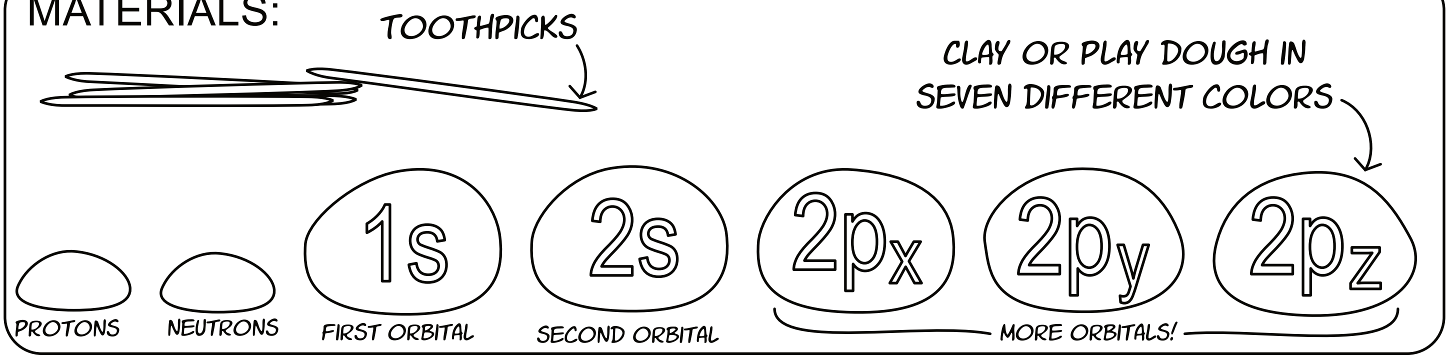
1 H Hydrogen																															2 He Helium
3 Li Lithium	4 Be Beryllium																									5 B Boron	6 C Carbon	7 N Nitrogen	8 O Oxygen	9 F Fluorine	10 Ne Neon
11 Na Sodium	12 Mg Magnesium																									13 Al Aluminum	14 Si Silicon	15 P Phosphorus	16 S Sulfur	17 Cl Chlorine	18 Ar Argon
19 K Potassium	20 Ca Calcium	21 Sc Scandium															22 Ti Titanium	23 V Vanadium	24 Cr Chromium	25 Mn Manganese	26 Fe Iron	27 Co Cobalt	28 Ni Nickel	29 Cu Copper	30 Zn Zinc	31 Ga Gallium	32 Ge Germanium	33 As Arsenic	34 Se Selenium	35 Br Bromine	36 Kr Krypton
37 Rb Rubidium	38 Sr Strontium	39 Y Yttrium															40 Zr Zirconium	41 Nb Niobium	42 Mo Molybdenum	43 Tc Technetium	44 Ru Ruthenium	45 Rh Rhodium	46 Pd Palladium	47 Ag Silver	48 Cd Cadmium	49 In Indium	50 Sn Tin	51 Sb Antimony	52 Te Tellurium	53 I Iodine	54 Xe Xenon
55 Cs Caesium	56 Ba Barium	57 La Lanthanum	58 Ce Cerium	59 Pr Praseodymium	60 Nd Neodymium	61 Pm Promethium	62 Sm Samarium	63 Eu Europium	64 Gd Gadolinium	65 Tb Terbium	66 Dy Dysprosium	67 Ho Holmium	68 Er Erbium	69 Tm Thulium	70 Yb Ytterbium	71 Lu Lutetium	72 Hf Hafnium	73 Ta Tantalum	74 W Tungsten	75 Re Rhenium	76 Os Osmium	77 Ir Iridium	78 Pt Platinum	79 Au Gold	80 Hg Mercury	81 Tl Thallium	82 Pb Lead	83 Bi Bismuth	84 Po Polonium	85 At Astatine	86 Rn Radon
87 Fr Francium	88 Ra Radium	89 Ac Actinium	90 Th Thorium	91 Pa Protactinium	92 U Uranium	93 Np Neptunium	94 Pu Plutonium	95 Am Americium	96 Cm Curium	97 Bk Berkelium	98 Cf Californium	99 Es Einsteinium	100 Fm Fermium	101 Md Mendelevium	102 No Nobelium	103 Lr Lawrencium	104 Rf Rutherfordium	105 Db Dubnium	106 Sg Seaborgium	107 Bh Bohrium	108 Hs Hassium	109 Mt Meitnerium	110 Ds Darmstadtium	111 Rg Roentgenium	112 Cn Copernicium	113 Nh Nihonium	114 Fl Flerovium	115 Mc Moscovium	116 Lv Livermorium	117 Ts Tennessine	118 Og Oganesson

ARRANGING ALL OF THE ELEMENTS BY NUMBER CREATES A REALLY WIDE TABLE. SO THIS BLOCK (THE LANTHANIDES AND ACTINIDES) IS USUALLY SHOWN BELOW THE REST OF THE ELEMENTS.

Hands-on Activity

MODELING CLAY ORBITALS!

MATERIALS:



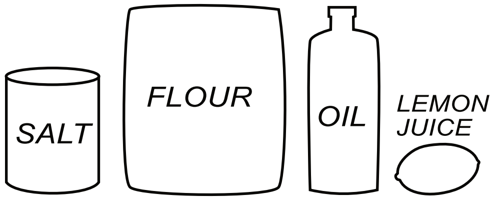
Don't have modeling clay? No problem! Make play dough using this recipe:

PLAY DOUGH

1 cup flour
1/3 cup salt
3/4 cup water
3 Tbsp lemon juice
1 Tbsp cooking oil
Food coloring

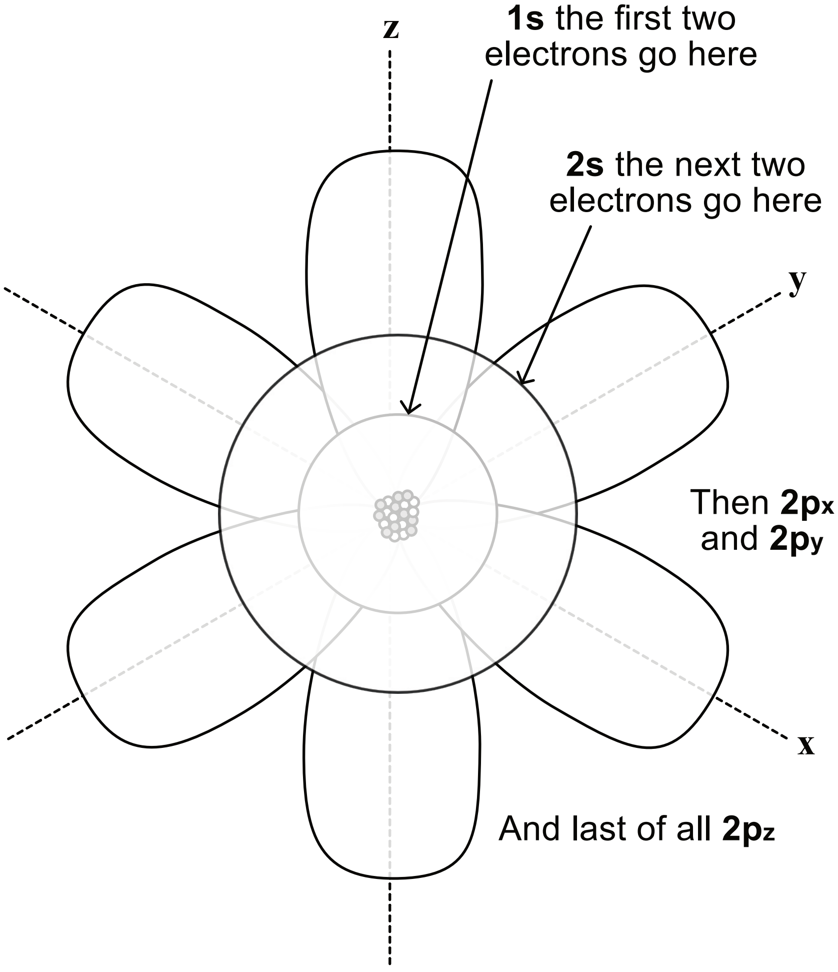
Mix the flour and salt together in a bowl. Heat the water to boiling and add the oil and lemon juice. Then mix all the ingredients together. For best results, mix in a pot over the stovetop until mixture is thick (about 1 minute).

Let sit and cool for a few minutes before kneading. Add another spoonful of flour if the dough is too sticky. Kool-aid drink packets can be used instead of food coloring.



ORBITALS

WHERE AN ELECTRON IS MOST LIKELY TO BE



1s the first two electrons go here

2s the next two electrons go here

Then **2p**x and **2p**y

And last of all **2p**z

~ INSTRUCTIONS ON PAGE 6 ~

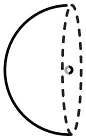
MODELING CLAY ORBITALS CONTINUED...

INSTRUCTIONS:

Shape the colors of clay that represent neutrons and protons into small spheres and put them together to make the nucleus. Remember, the number of protons determines the element! making models of a hydrogen, helium, lithium, carbon, fluorine, and neon atom. Partially-filled orbitals can be represented by moulding half of the orbital. Use toothpicks to attach the p-orbitals.

HYDROGEN

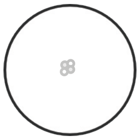
1 PROTON
0 NEUTRONS
1 ELECTRON



WARNING! VERY REACTIVE
ORBITAL INCOMPLETE

HELIUM

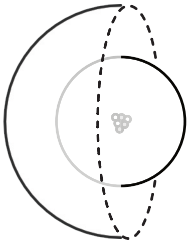
2 PROTON
2 NEUTRONS
2 ELECTRON



CONGRATULATIONS!
YOU ARE REMARKABLY STABLE

LITHIUM

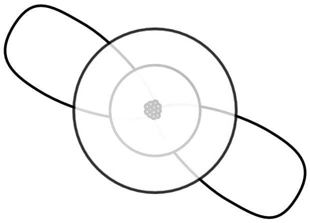
3 PROTON
3 NEUTRONS
3 ELECTRON



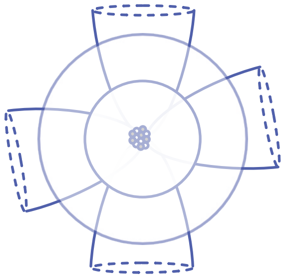
WARNING! VERY REACTIVE
ORBITAL INCOMPLETE

CARBON

6 PROTON
6 NEUTRONS
6 ELECTRON

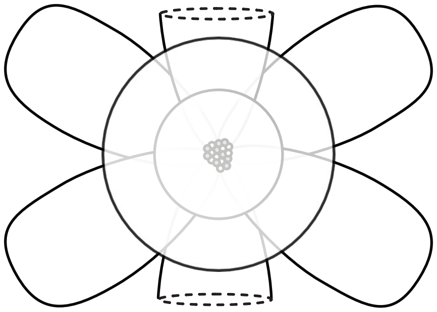


Science mom made a mistake! With the "half full orbitals" that she drew to represent an orbital with just 1 electron, carbon should look like this:



FLUORINE

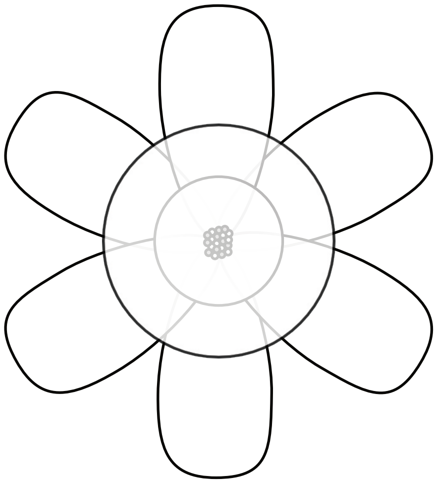
9 PROTON
9 NEUTRONS
9 ELECTRON



WARNING! VERY REACTIVE
ORBITAL INCOMPLETE

NEON

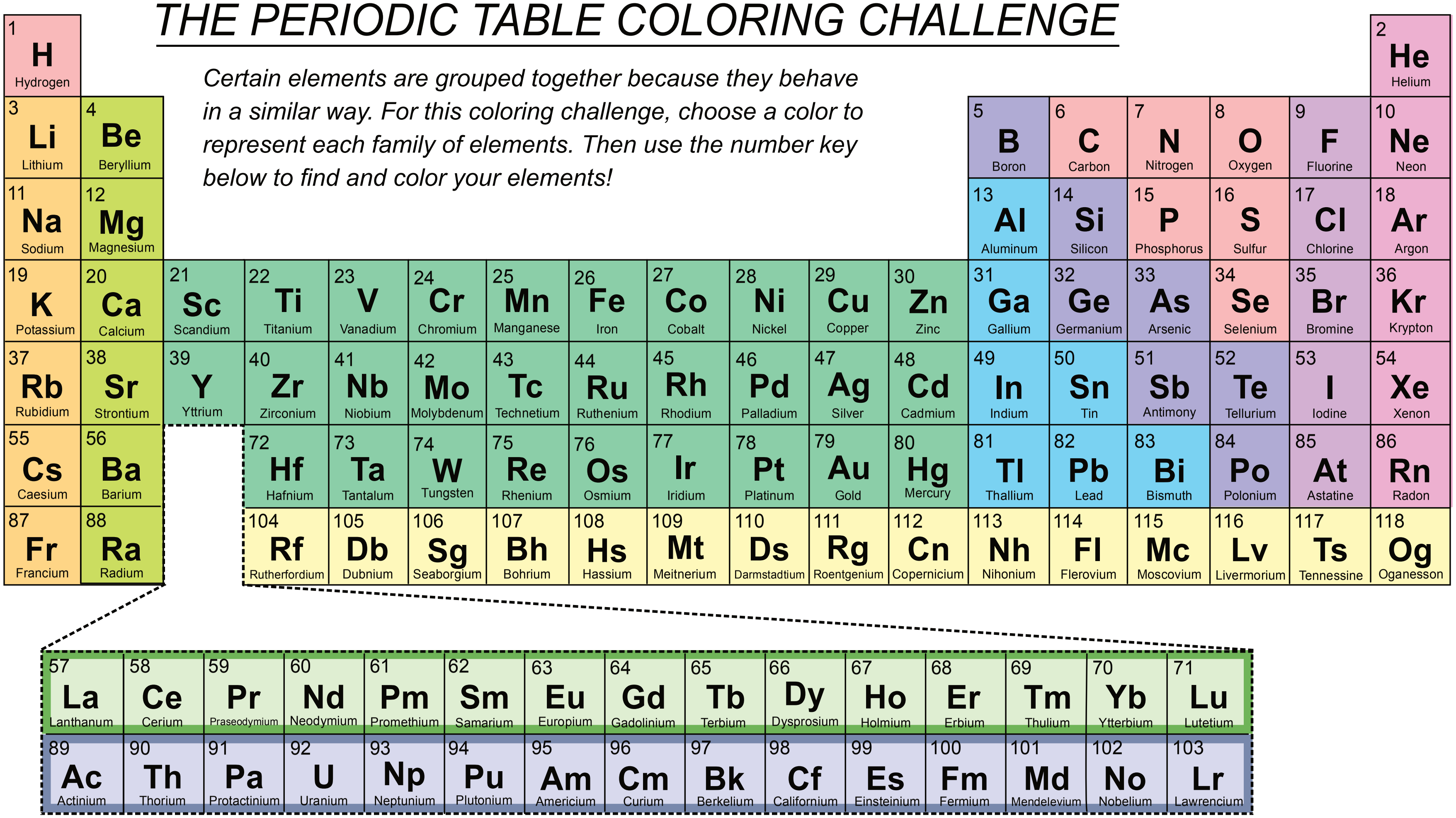
10 PROTON
10 NEUTRONS
10 ELECTRON



CONGRATULATIONS!
YOU ARE REMARKABLY STABLE

THE PERIODIC TABLE COLORING CHALLENGE

Certain elements are grouped together because they behave in a similar way. For this coloring challenge, choose a color to represent each family of elements. Then use the number key below to find and color your elements!



1 H Hydrogen																	2 He Helium
3 Li Lithium	4 Be Beryllium											5 B Boron	6 C Carbon	7 N Nitrogen	8 O Oxygen	9 F Fluorine	10 Ne Neon
11 Na Sodium	12 Mg Magnesium											13 Al Aluminum	14 Si Silicon	15 P Phosphorus	16 S Sulfur	17 Cl Chlorine	18 Ar Argon
19 K Potassium	20 Ca Calcium	21 Sc Scandium	22 Ti Titanium	23 V Vanadium	24 Cr Chromium	25 Mn Manganese	26 Fe Iron	27 Co Cobalt	28 Ni Nickel	29 Cu Copper	30 Zn Zinc	31 Ga Gallium	32 Ge Germanium	33 As Arsenic	34 Se Selenium	35 Br Bromine	36 Kr Krypton
37 Rb Rubidium	38 Sr Strontium	39 Y Yttrium	40 Zr Zirconium	41 Nb Niobium	42 Mo Molybdenum	43 Tc Technetium	44 Ru Ruthenium	45 Rh Rhodium	46 Pd Palladium	47 Ag Silver	48 Cd Cadmium	49 In Indium	50 Sn Tin	51 Sb Antimony	52 Te Tellurium	53 I Iodine	54 Xe Xenon
55 Cs Caesium	56 Ba Barium		72 Hf Hafnium	73 Ta Tantalum	74 W Tungsten	75 Re Rhenium	76 Os Osmium	77 Ir Iridium	78 Pt Platinum	79 Au Gold	80 Hg Mercury	81 Tl Thallium	82 Pb Lead	83 Bi Bismuth	84 Po Polonium	85 At Astatine	86 Rn Radon
87 Fr Francium	88 Ra Radium		104 Rf Rutherfordium	105 Db Dubnium	106 Sg Seaborgium	107 Bh Bohrium	108 Hs Hassium	109 Mt Meitnerium	110 Ds Darmstadtium	111 Rg Roentgenium	112 Cn Copernicium	113 Nh Nihonium	114 Fl Flerovium	115 Mc Moscovium	116 Lv Livermorium	117 Ts Tennessine	118 Og Oganesson

57 La Lanthanum	58 Ce Cerium	59 Pr Praseodymium	60 Nd Neodymium	61 Pm Promethium	62 Sm Samarium	63 Eu Europium	64 Gd Gadolinium	65 Tb Terbium	66 Dy Dysprosium	67 Ho Holmium	68 Er Erbium	69 Tm Thulium	70 Yb Ytterbium	71 Lu Lutetium
89 Ac Actinium	90 Th Thorium	91 Pa Protactinium	92 U Uranium	93 Np Neptunium	94 Pu Plutonium	95 Am Americium	96 Cm Curium	97 Bk Berkelium	98 Cf Californium	99 Es Einsteinium	100 Fm Fermium	101 Md Mendelevium	102 No Nobelium	103 Lr Lawrencium

🟥 **Nonmetals:** These Elements do not conduct electricity. 1,6,7,8,15,16,34

🟧 **Alkali Metals:** All of these react explosively with water. 3,11,19,37,55,87

🟩 **Alkali Earth metals:** These all also reactive elements and especially like to react with oxygen. 4,12,20,38,56,88

🟩 **Transition metals:** These are good conductors of heat and electricity. And there are a lot of them! 21, 22, 23, 24 25, 26, 27, 28, 29, 30, 39, 40, 41, 42, 43, 44, 45, 46, 47, 48, 72, 73, 74, 75, 76, 77, 78, 79, 80

🟦 **Metals:** These are great conductors heating electricity and in their solid form they can are shiny and ductile. 13, 31, 49, 50, 81, 82, 83

🟪 **Metaloids:** these elements are semiconductors! 5,14,32,33,51,52,84

🟪 **Halogens:** These are very reactive elements.9,17,35,53,85

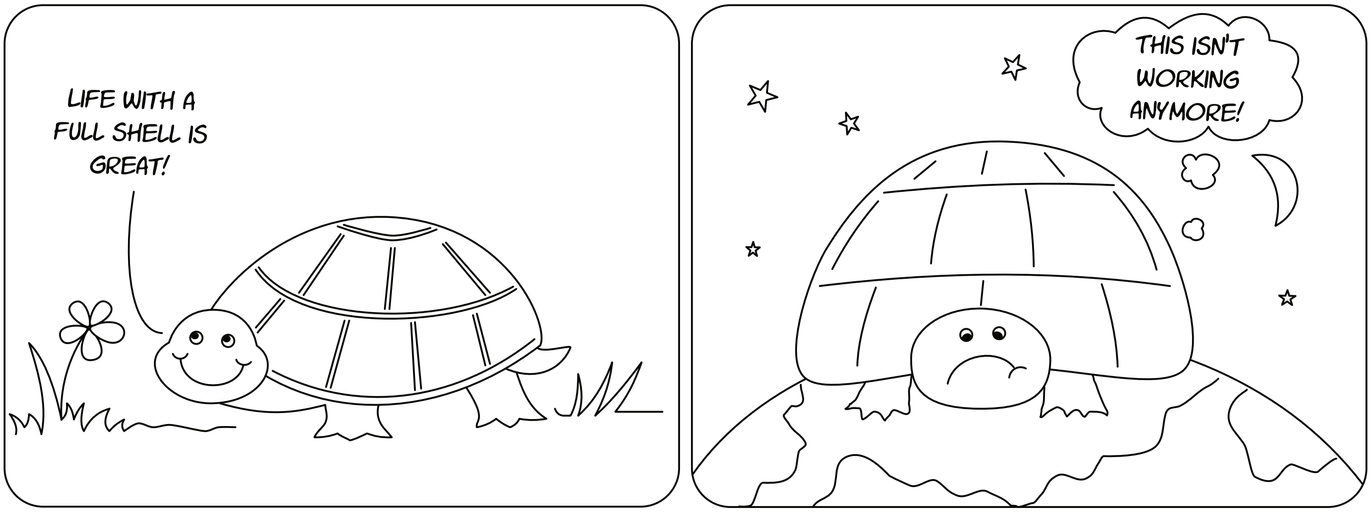
🟥 **Noble gases:** These elements have a full show of electrons and are not very reactive. 2,10,18,36,54,86

🟨 **Transactinides:** Super big elements with more than 104 protons! These have been created artificially in laboratories, but are not found in nature.104-118

🟩 **Lanthanides:** Called the rare earth elements.57-71

🟦 **Actinides:** These are all radioactive. 89-103

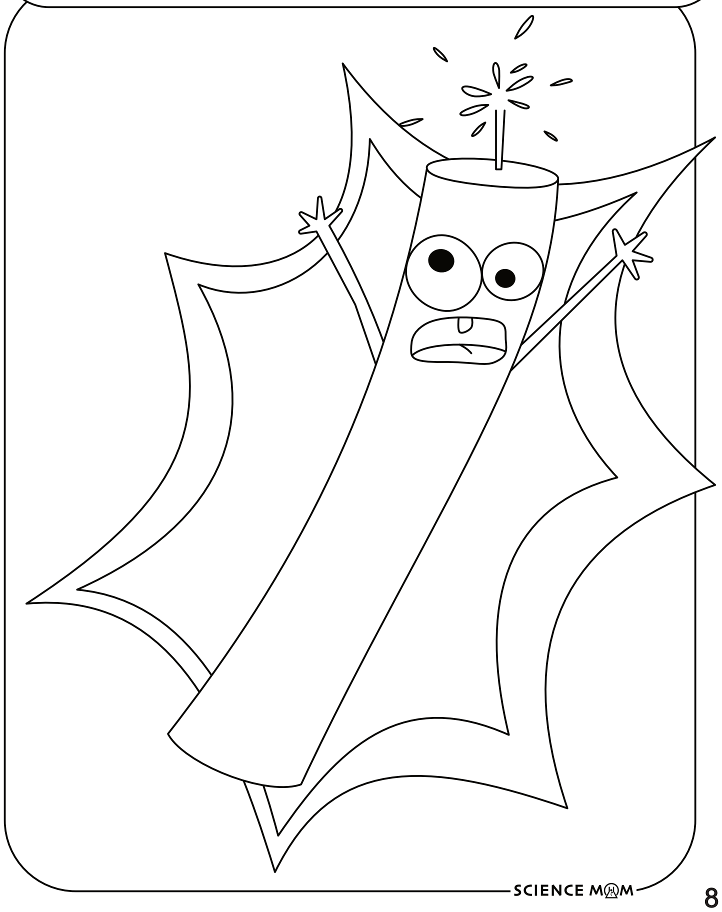
A FULL SHELL OF ELECTRONS IS LIKE A HAPPY TURTLE - UNLESS IT GETS TOO BIG.



An ____element____ with a full shell is stable. It is not interested in reacting with other elements. But if it gets too large, then that “turtle” is no longer very happy, even though it has a full shell.

The elements with ___full___ shells of electrons are in the column called the noble gases. Next to the noble gases are the ____halogens____. If these elements *gain* one more electron, then they have a full shell. If the ___alkali___ ___metals___ *lose* one electron, then they have a full shell. Both groups or families of elements are very ___reactive___. They want to ___bond___ with other elements and fill their shells!

AN ALMOST-FULL SHELL OF ELECTRONS IS LIKE A FIRECRACKER READY TO EXPLODE!



FILL IN THE BLANKS USING THESE WORDS:

reactive	alkali	element	full
metals	halogens	periodic	bond

Your notes: ______________________________

____

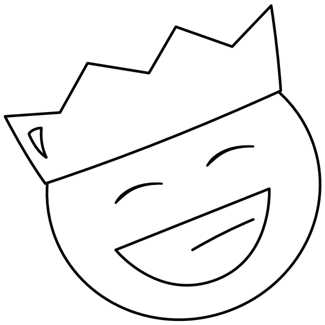
____

____

____

____

A NOBLE Quest!



The elements in this family are called the "Noble gases." At room temperature, they are all are colorless, odorless, and tasteless. They hardly ever form bonds or react with anything! Can you draw lines to match each element with it's fact box?

2
He
Helium
4

10
Ne
Neon
20.18

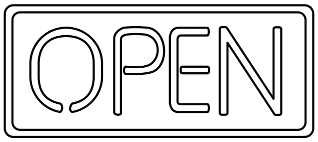
18
Ar
Argon
39.95

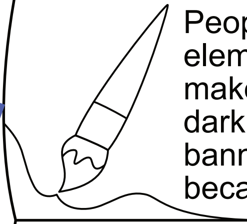
36
Kr
Krypton
83.80

54
Xe
Xenon
131.24

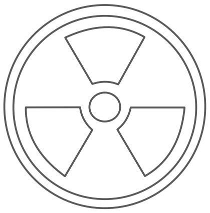
86
Rn
Radon
222

When electricity passes through this colorless gas, it can glow a bright red-orange color. It's often used in signs.



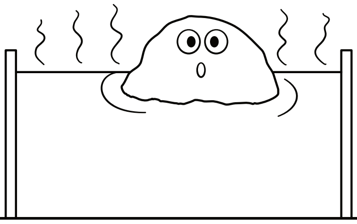


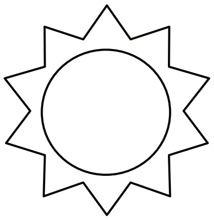
People used to mix this element into paints to make them glow in the dark, but then it was banned from paint because it's radioactive.



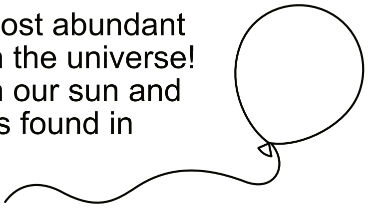
This gas is sometimes used in high-powered lasers.

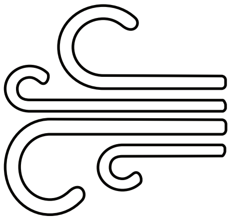
This gas glows bright white when electricity passes through it. In its (very cold) liquid form, it's so dense that granite would float on it!





Second most abundant element in the universe! Created in our sun and sometimes found in balloons.





This gas, the 3rd most abundant in our atmosphere, makes up 1% of the air you breath.

1%

The Halogens

REACTIVE & DANGEROUS

The elements in this family are called the "halogens." At room temperature, the first two (fluorine and chlorine) are gasses with strong unpleasant smells. Breathing too much of them is toxic and they are all flammable and corrosive (will destroy or damage other substances). Can you draw lines to match each element in this family with its fact box?

9
F
Fluorine
19.0

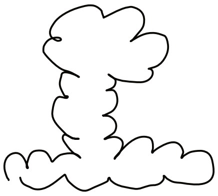
17
Cl
Chlorine
35.45

35
Br
Bromine
79.80

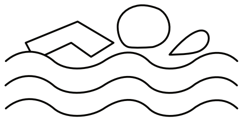
53
I
Iodine
126.91

85
At
Astatine
210

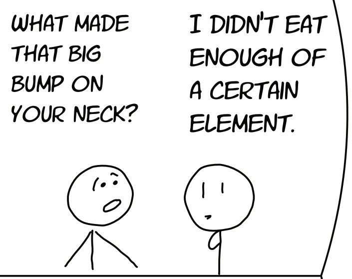
Rarest of the naturally occurring elements, this radioactive element will only be around for a few hours, but while it is, it can do impressive damage. If you had more than few molecules of this element, they'd immediately be vaporized by the heat of their own radioactivity.



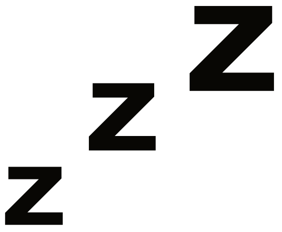
Used to treat water in swimming pools.



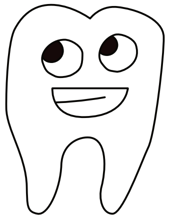
Lack of this element in a person's diet can cause a goiter: a huge overgrown thyroid that sticks out the from the neck.



Salts of this element were once used as sedatives.



Can strengthen tooth enamel.

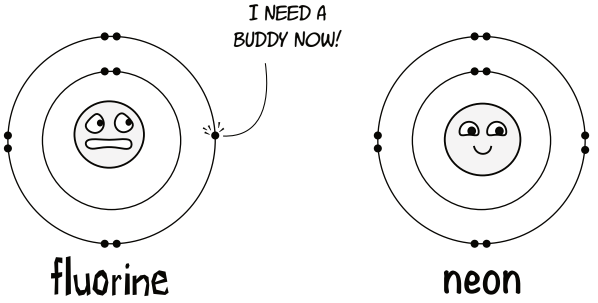


Artwork by MATH DAD! ☺

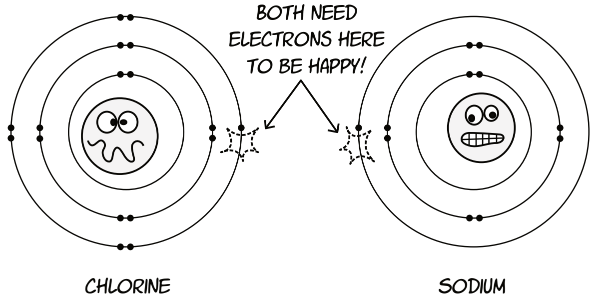
CHEMICAL BONDS

SHARING
ELECTRONS
MAKES ATOMS
HAPPY!

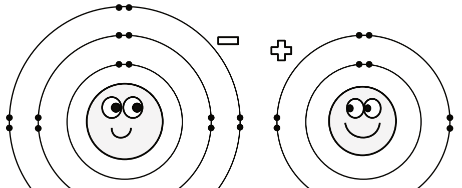
Electrons really like to be in pairs. Fluorine, which is super reactive, has nine electrons, leaving one of them unpaired. Neon, a nonreactive noble gas, has ten electrons, each of them paired in different “shells” or orbitals around the nucleus.



By themselves, chlorine and sodium are both “unhappy” because they have unpaired electrons.



But if sodium gives it’s lonely electron to chlorine, then they’re both happy. They’ve formed an ionic bond! Other atoms solve the same problem by sharing electrons.



Your notes: ______________________________

____

____

____

____

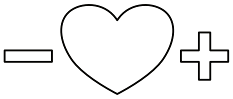
____

____

____

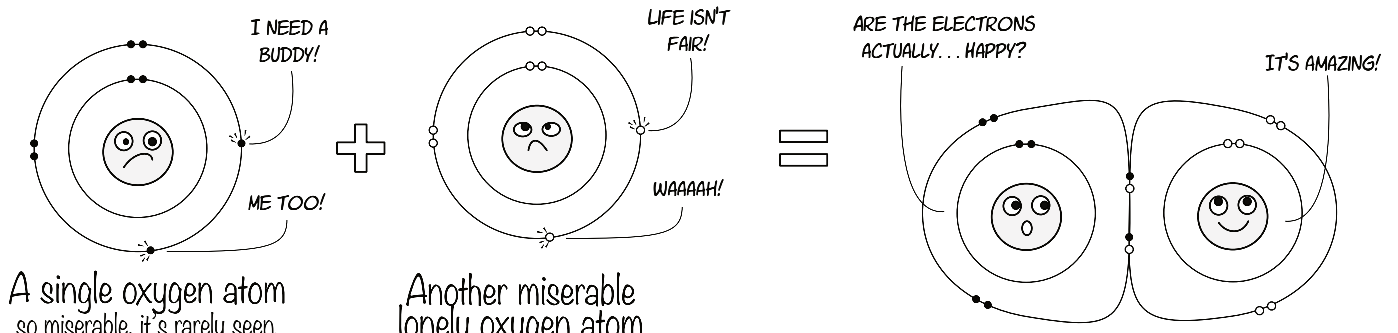
____

IONIC BOND: A CHEMICAL BOND WHERE AN ELECTRON IS TRANSFERRED FROM ONE ATOM TO ANOTHER. THIS CREATES **IONS** WITH OPPOSITE CHARGES. AND OPPOSITES ATTRACT!



YOUR DOODLE SPACE

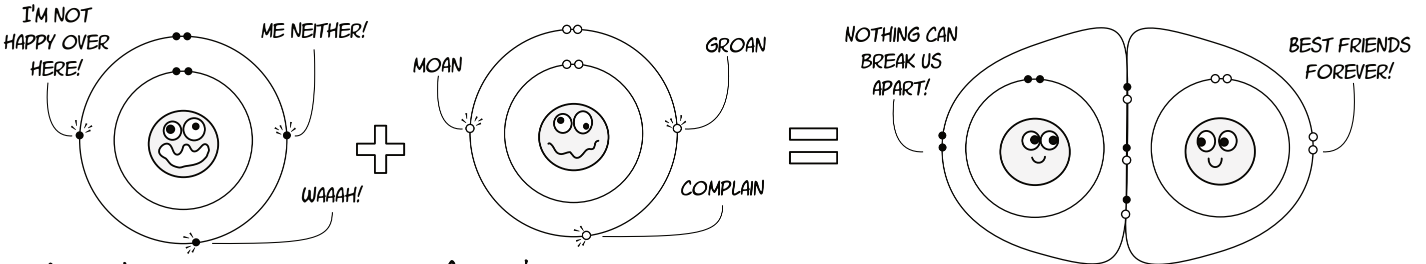
Draw your favorite moment from class or write a cool fact!



A single oxygen atom
so miserable, it's rarely seen. When it does go out, it does a lot of damage.

Another miserable lonely oxygen atom
We've drawn the electrons differently here so you can better appreciate what happens next!

Oxygen gas - formula O_2
Since they share two pairs of electrons, we call this a "double" bond!

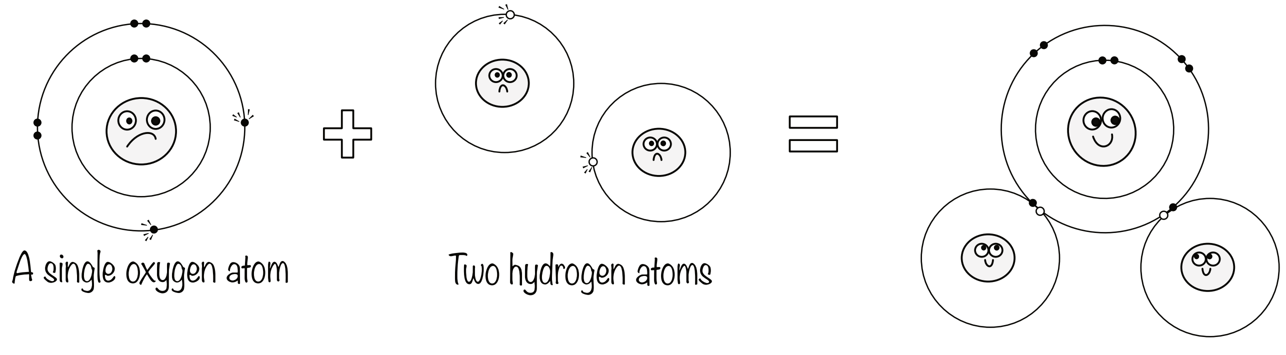


A single nitrogen atom

A single nitrogen atom

N_2 - atmospheric nitrogen
Since they share two pairs of electrons, we call this a "double" bond!

One oxygen + 2 hydrogens = H_2O



A single oxygen atom

Two hydrogen atoms

A molecule of water!

COVALENT BOND: A CHEMICAL BOND WHERE ELECTRONS ARE SHARED BETWEEN TWO ATOMS. SOMETIMES THE ELECTRONS ARE SHARED EQUALLY, AND OTHER TIMES ONE ATOM (WE'RE TALKING ABOUT YOU, OXYGEN!) WILL BE A BIT GREEDY.

Your notes: ______________________________

____

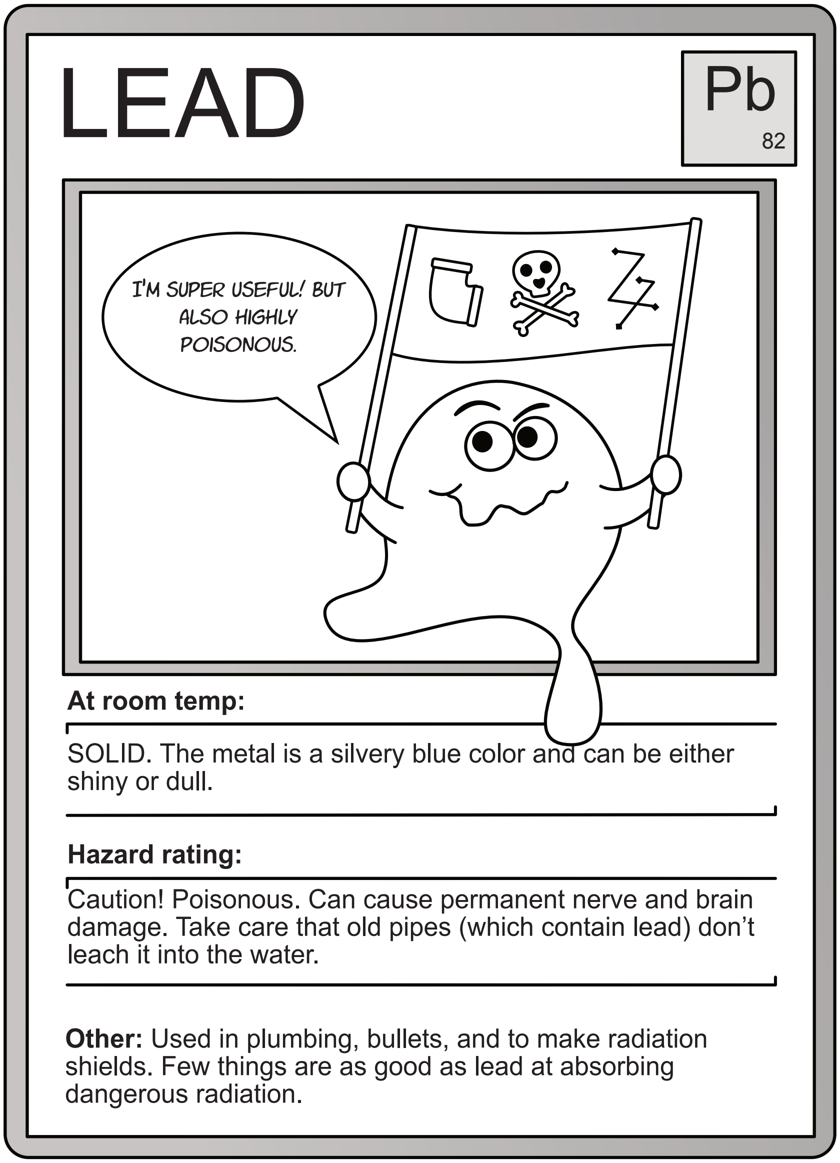
____

____

____

Make your own Element Cards!

Choose four elements to study. Research them and draw cards for them on the blank templates on the next page (you can print more pages to make more if you'd like!) Be sure to look up the chemical symbol and atomic number of your element. Research how your element behaves at room temperature and give it a hazard rating too. Then draw an avatar. It can look like anything! Be creative and have fun designing your cards.



LEAD

Pb
82

At room temp:

SOLID. The metal is a silvery blue color and can be either shiny or dull.

Hazard rating:

Caution! Poisonous. Can cause permanent nerve and brain damage. Take care that old pipes (which contain lead) don't leach it into the water.

Other: Used in plumbing, bullets, and to make radiation shields. Few things are as good as lead at absorbing dangerous radiation.

NAME →

← CHEMICAL SYMBOL

YOUR ELEMENT'S AVATAR! IT CAN LOOK LIKE ANYTHING YOU'D LIKE. BE CREATIVE AND HAVE FUN! →

← IS YOUR ELEMENT A SOLID, LIQUID, OR GAS AT ROOM TEMPERATURE?

IN ITS PURE FORM, IS THIS ELEMENT COMPLETELY SAFE, MILDLY DANGEROUS, TOXIC, OR RADIOACTIVE? →

← A NOTE ABOUT HOW HUMANS USE THE ELEMENT, WHERE IT'S FOUND, OR ANY OTHER COOL FACT YOU DISCOVERED.

At room temp:

Hazard rating:

Other:

At room temp:

Hazard rating:

Other:

At room temp:

Hazard rating:

Other:

At room temp:

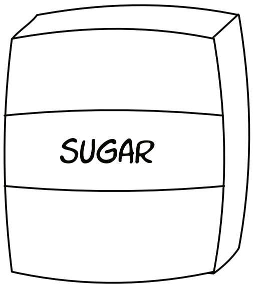
Hazard rating:

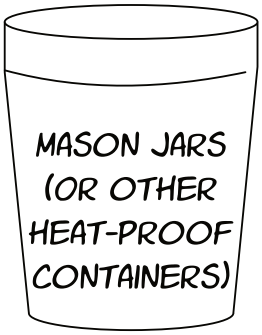
Other:

Hands-on Activity

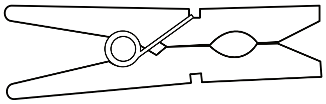
EDIBLE EXPERIMENTS - ROCK CANDY!

MATERIALS:





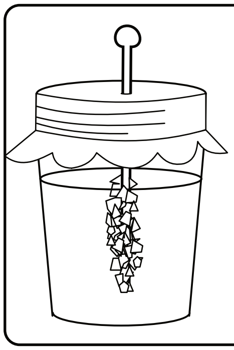
CLOTHES PINS



ROCK CANDY STICK,
SKEWER OR STRING

KOOL-AID PACKETS
(0.13 OUNCES
UNFLAVORED)

*Don't have Kool-Aid?
No problem! Use food
coloring to color the crystals.*



COFFEE FILTER METHOD

ROCK CANDY

2 pint sized mason jars
(Or other heat-proof containers)
1 cup water
3 ½ cups sugar
2 Kool-Aid packets

① Prepare the sticks or string by getting them wet and rolling them in dry sugar.

② Bring the water to a boil, then add the sugar and stir well. Reduce the heat and continue cooking until the solution turns clear and all the sugar dissolves.

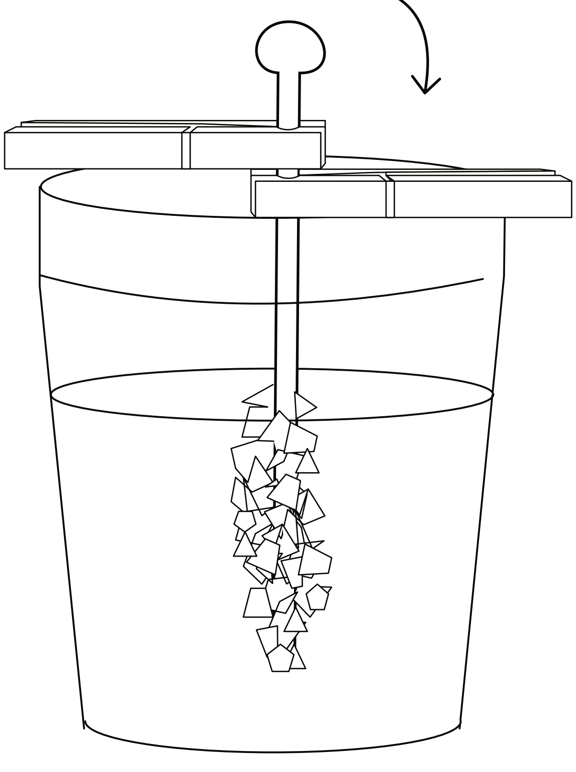
③ Pour the powder from one Kool-Aid packet into each mason jar.

④ Very carefully, pour the sugar solution into the mason jars and stir well to make sure that the Kool-Aid mixes in.

⑤ Use the clothes pins or coffee filter to suspend the stick in the center of the jar.

⑥ Let the jars sit for 2 to 8 days. Crystal formation takes time, be patient!

CLOTHES PIN METHOD



The science behind the treat:

When sugar dissolves into water it forms a MIXTURE - the sugar is still there and the water is still there. New molecules have NOT been formed. But the sugar molecules are attracted to the water and visa versa. When the water is HOT, it can hold more sugar than when it is cool. If you add as much sugar as the water can "carry" when it's hot, then as it cools the sugar will "come out" of the water and you'll see crystals form. If the sugar crystals grow slowly, you end up with larger crystals. If the sugar crystals grow quickly, they're smaller.

YOUR DOODLE SPACE

Draw your favorite moment from class or write a cool fact!

Troubleshooting tips:

What if there are no crystals on your stick? First, did you "seed" it by getting it wet and rolling it in dry sugar before-hand? This really helps! Second, sometimes the crystals take DAYS (up to 7 or 10) to form. If your first batch isn't working, you can try again and increase the amount of sugar (add an extra cup). The hardness of your water and measuring error can make a difference. If you don't see crystals after 14 days, probably best to try again with a fresh batch and add some extra sugar this time.

Do you think you could also make salt crystals using the same recipe? Why or why not?

Share your opinion! Any answer works, but here are some potential ideas:

Arguments in favor of salt crystals: salt can also dissolve in water. If hot water holds more salt than cold water, then crystals should form after the solution cools.

Arguments against salt crystals: Salt is a different molecule than sugar. It may behave differently. You can't dissolve as much salt in water as sugar.

How did your crystals turn out? Did you see a difference between the size and shape of the crystals in different jars? How long did it take before your rock candy started growing?

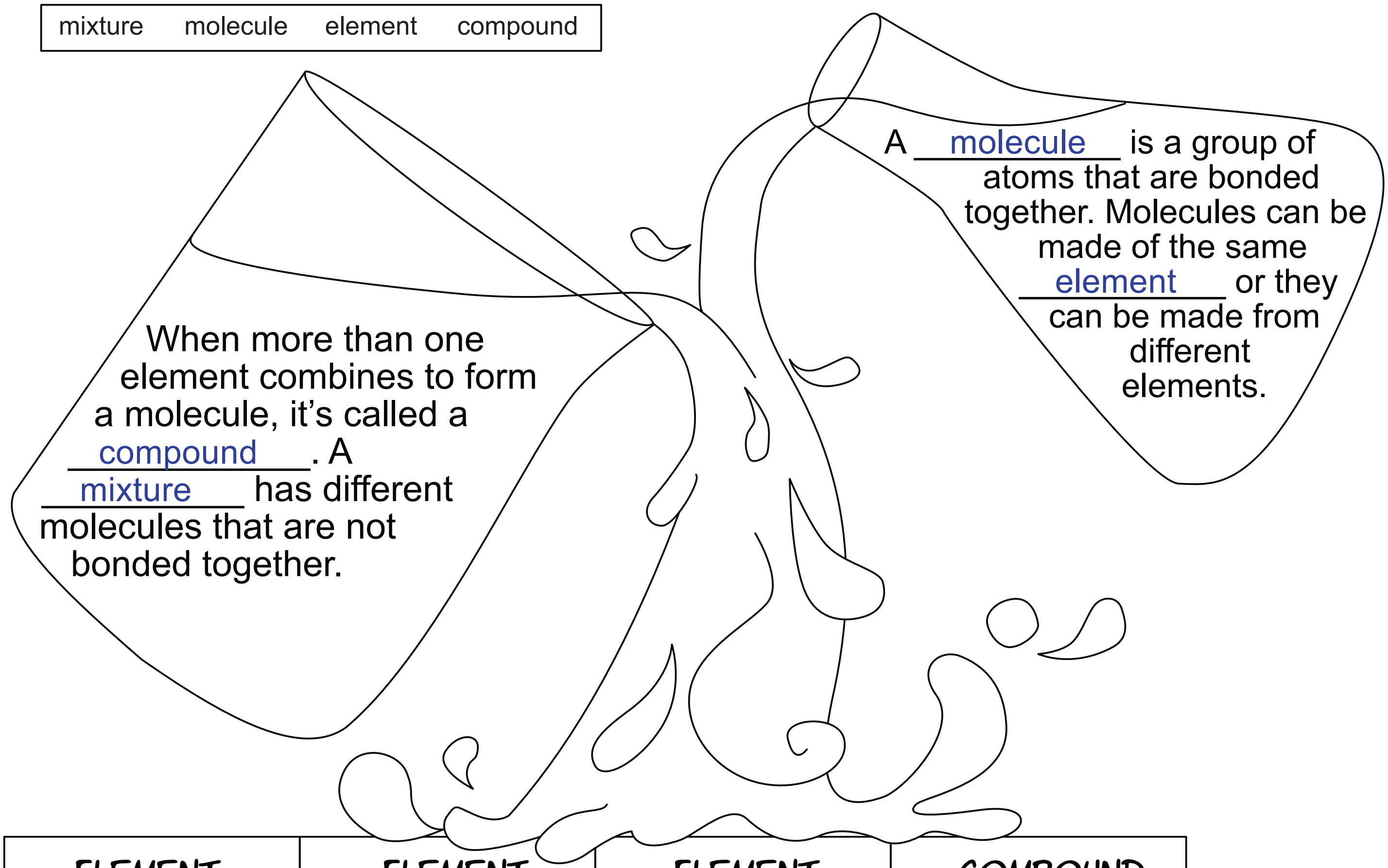
Crystals that form slowly tend to be larger than crystals that form quickly.

Sometimes it takes up to 2 weeks before they form. If you don't see crystals. In a few days time, you can always try a second batch with a higher concentration of sugar.

Element vs Mixture vs Compound

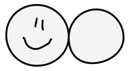
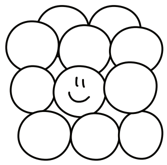
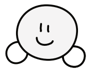
FILL IN THE BLANKS USING THESE WORDS:

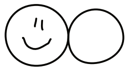
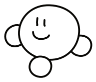
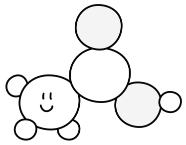
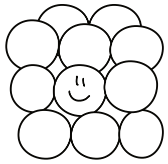
mixture molecule element compound



When more than one element combines to form a molecule, it's called a __compound__. A __mixture__ has different molecules that are not bonded together.

A __molecule__ is a group of atoms that are bonded together. Molecules can be made of the same __element__ or they can be made from different elements.

ELEMENT	ELEMENT	ELEMENT	COMPOUND
			
A MOLECULE OF HELIUM	A MOLECULE OF OXYGEN	A MOLECULE OF GOLD	A MOLECULE OF WATER

element	compound	compound	element
			
A MOLECULE OF NITROGEN	A MOLECULE OF AMMONIA	A MOLECULE OF VINEGAR	A MOLECULE OF IRON

WRITE DOWN WHETHER EACH OF THESE IS AN ELEMENT OR COMPOUND!

Artwork by MATH DAD ☺

Element, Isotope, or Ion?

Each of the atoms below is going to gain or lose protons, neutrons, or electrons. Write down what the atom will be after that change!

1 **H** Hydrogen								2 **He** Helium
3 **Li** Lithium	4 **Be** Beryllium		5 **B** Boron	6 **C** Carbon	7 **N** Nitrogen	8 **O** Oxygen	9 **F** Fluorine	10 **Ne** Neon
			13 **Al** Aluminum	14 **Si** Silicon	15 **P** Phosphorus	16 **S** Sulfur	17 **Cl** Chlorine	18 **Ar** Argon
28 **Ni** Nickel	29 **Cu** Copper	30 **Zn** Zinc	31 **Ga** Gallium	32 **Ge** Germanium	33 **As** Arsenic	34 **Se** Selenium	35 **Br** Bromine	36 **Kr** Krypton
46 **Pd** Palladium	47 **Ag** Silver	48 **Cd** Cadmium	49 **In** Indium	50 **Sn** Tin	51 **Sb** Antimony	52 **Te** Tellurium	53 **I** Iodine	54 **Xe** Xenon
78 **Pt** Platinum	79 **Au** Gold	80 **Hg** Mercury	81 **Tl** Thallium	82 **Pb** Lead	83 **Bi** Bismuth	84 **Po** Polonium	85 **At** Astatine	86 **Rn** Radon

CARBON 12
6 protons
6 neutrons
6 electrons

→ +2 neutrons →

CARBON 14
6 protons
8 neutrons
6 electrons

HYDROGEN 1
(Also called protium)
1 protons
0 neutrons
1 electrons

→ -1 electron →

HYDROGEN ION
+
1 protons
0 neutrons
0 electrons

HELIUM 4
2 protons
2 neutrons
2 electrons

→ +1 proton →

LITHIUM 5
UNSTABLE
+
3 protons
2 neutrons
2 electrons

HELIUM 4
2 protons
2 neutrons
2 electrons

→ -1 proton →

HYDROGEN 3
STABLE, BUT VERY RARE
-
1 proton~~s~~
2 neutrons
2 electrons

FLUORINE 18
9 protons
9 neutrons
9 electrons

→ +1 proton →

NEON 19
UNSTABLE
+
10 protons
9 neutrons
9 electrons

BORON 10
5 protons
5 neutrons
5 electrons

→ +1 neutron →

BORON 11
I'M STABLE!
5 protons
6 neutrons
5 electrons

NITROGEN 14
7 protons
7 neutrons
7 electrons

→ +1 proton →

OXYGEN 15
I'M UNSTABLE!
+
8 protons
7 neutrons
7 electrons

OXYGEN 16
8 protons
8 neutrons
8 electrons

→ +1 electron →

OXYGEN 16
I'M CHARGED!
-
8 protons
8 neutrons
9 electrons

COPPER 63
29 protons
34 neutrons
29 electrons

→ +1 proton →

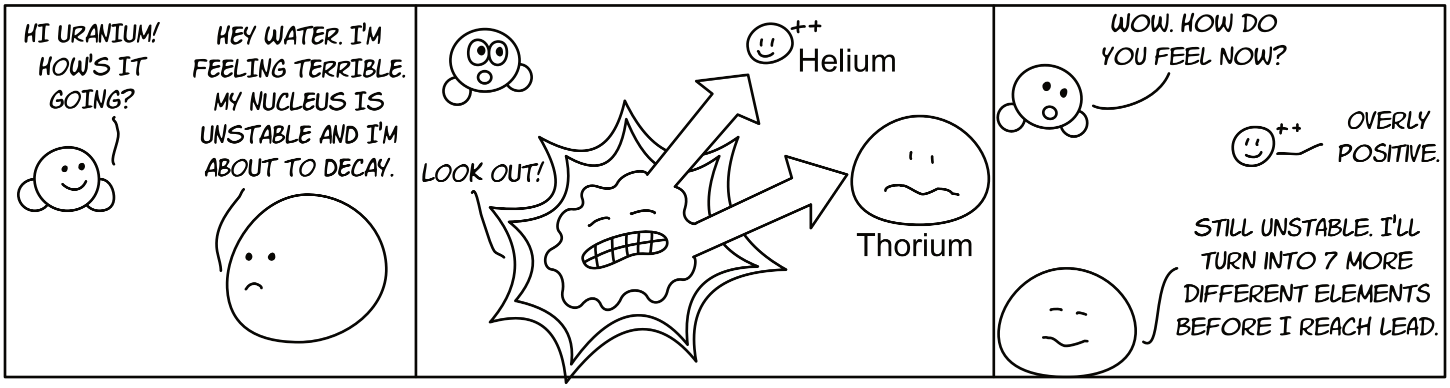
ZINC 64
KINDA STABLE.
+
30 protons
34 neutrons
29 electrons

COPPER 63
29 protons
34 neutrons
29 electrons

→ +1 neutron →

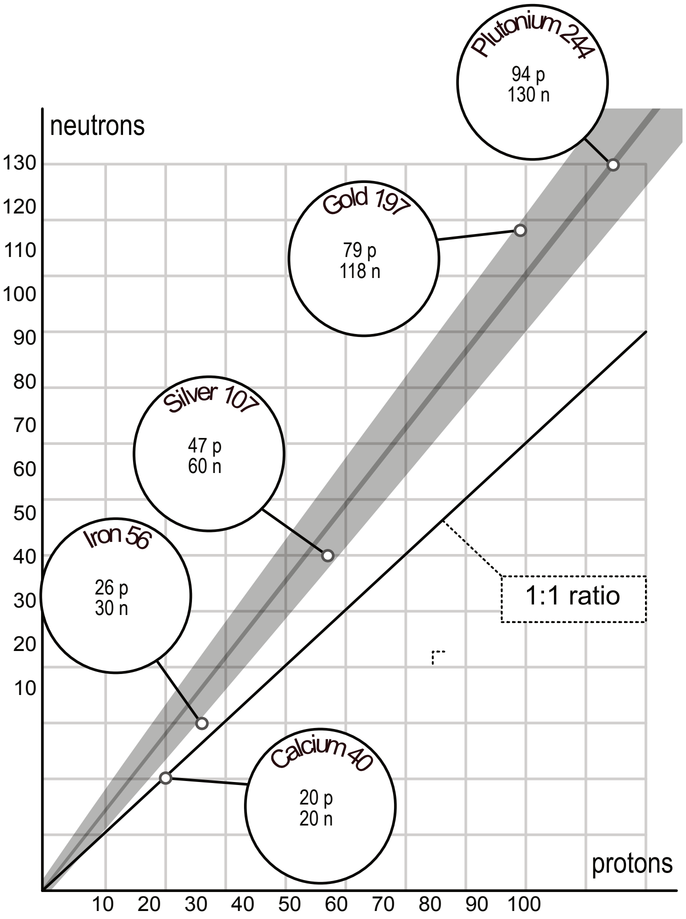
COPPER 64
UNSTABLE
29 protons
35 neutrons
29 electrons

What is RADIOACTIVITY?



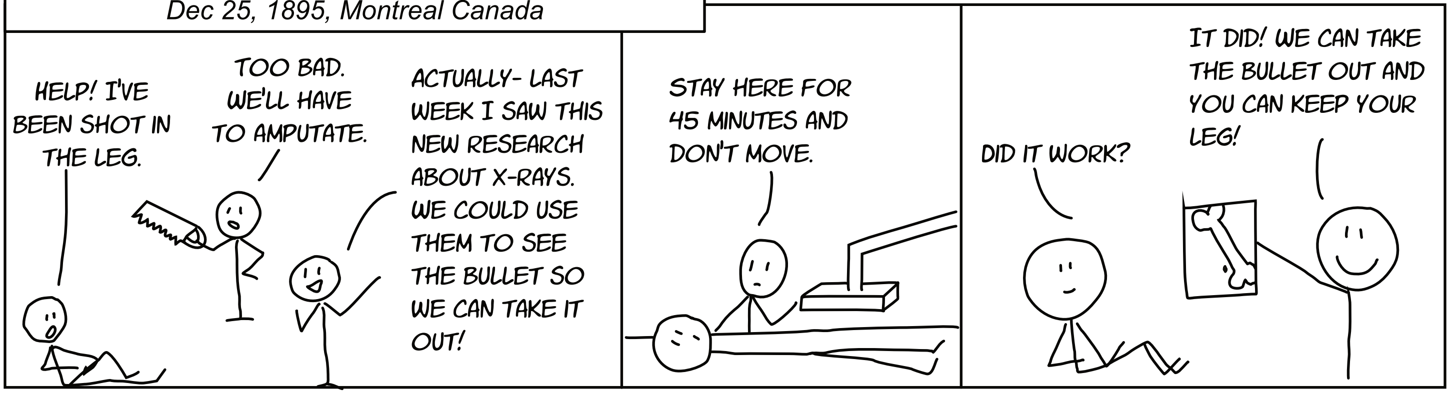
Unstable atoms decay. They split apart to form new elements. You might think that an equal number of protons and neutrons would be the most stable situation, but look at this graph and you'll see that's not the case! Hydrogen is most stable with no neutrons. Larger elements, like gold, need many more neutrons than protons.

Your notes: ______________________________



THE STORY OF THE FIRST MEDICAL X-RAY

Dec 25, 1895, Montreal Canada

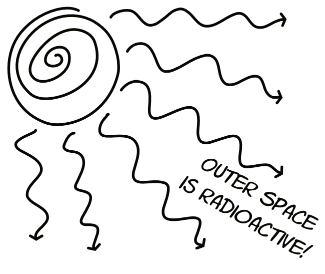


Artwork by MATH DAD ☺

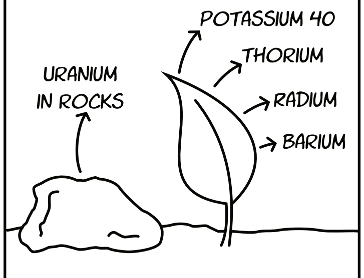
Radiation is

NATURAL

THAT DOESN'T MEAN IT'S GOOD FOR YOU.



THE SUN AND COSMIC RAYS

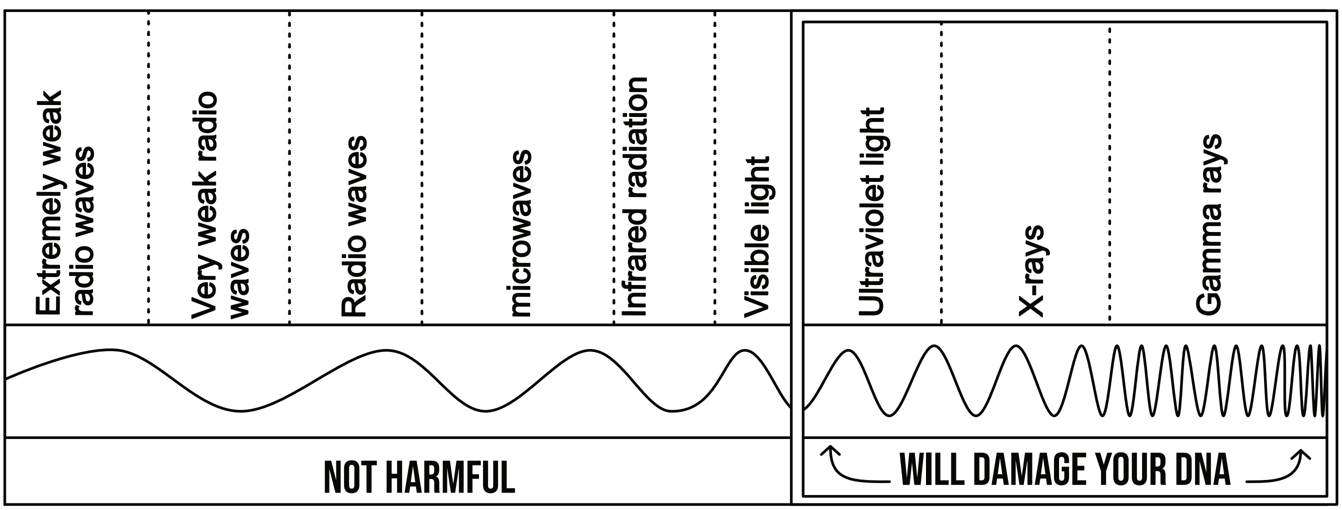


SOIL AND PLANTS

IT'S THE DOSE THAT MAKES THE POISON

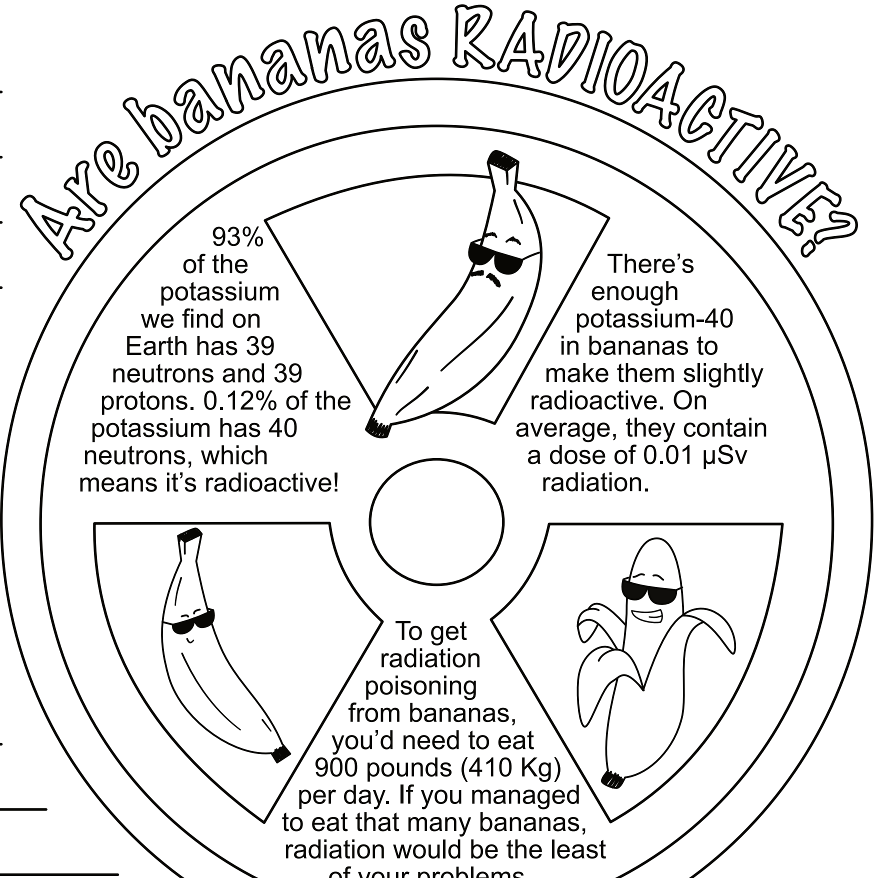
ANYTHING CAN BE TOXIC IF THERE IS TOO MUCH OF IT- EVEN WATER OR OXYGEN. WITH RADIATION, THE THING THAT REALLY MATTERS IS **HOW MUCH.**

All of these types of energy are called "radiation"



Your notes: ______________________

Are bananas RADIOACTIVE?



93% of the potassium we find on Earth has 39 neutrons and 39 protons. 0.12% of the potassium has 40 neutrons, which means it's radioactive!

There's enough potassium-40 in bananas to make them slightly radioactive. On average, they contain a dose of 0.01 µSv radiation.

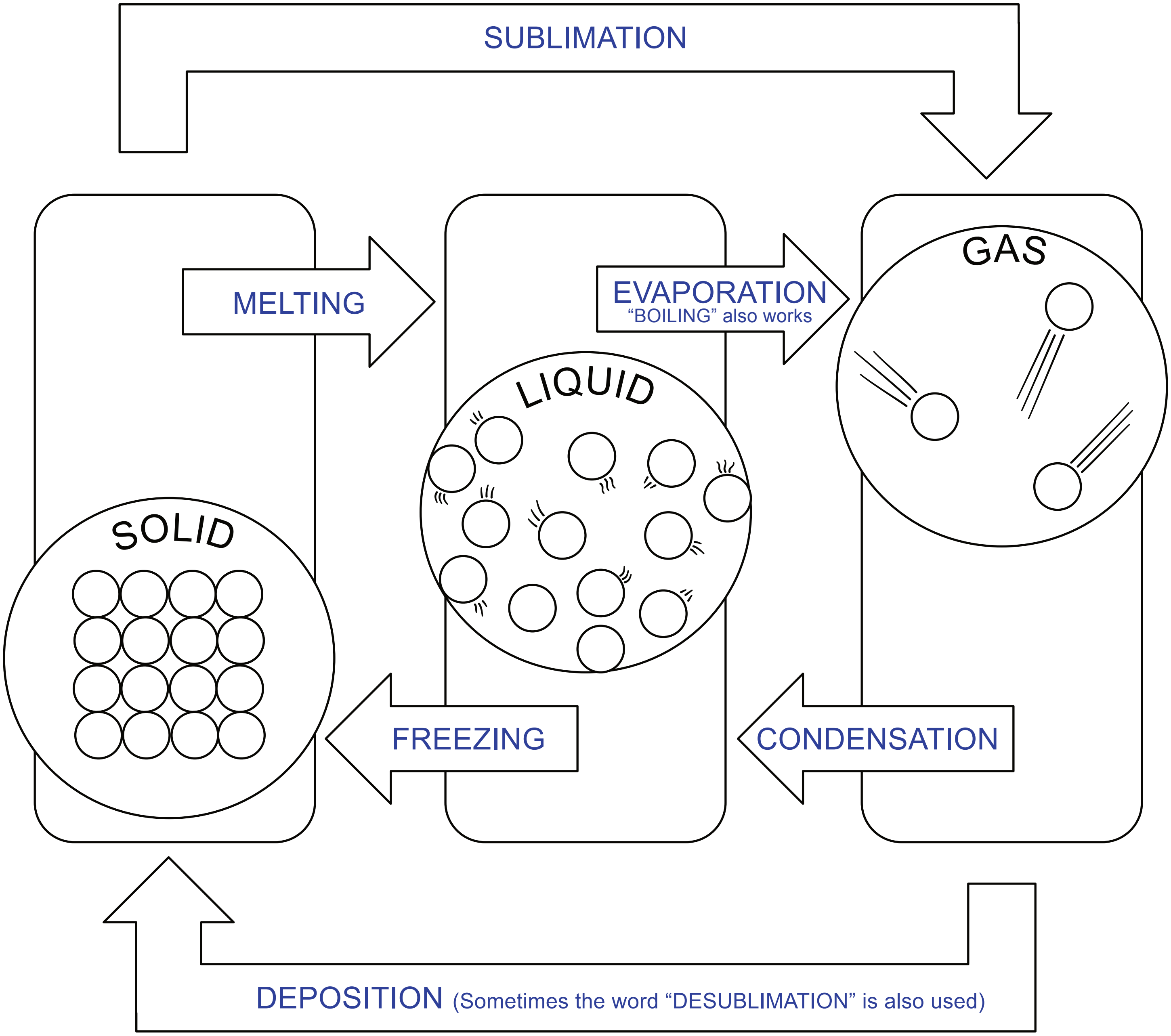
To get radiation poisoning from bananas, you'd need to eat 900 pounds (410 Kg) per day. If you managed to eat that many bananas, radiation would be the least of your problems.

CONSERVATION and states of matter

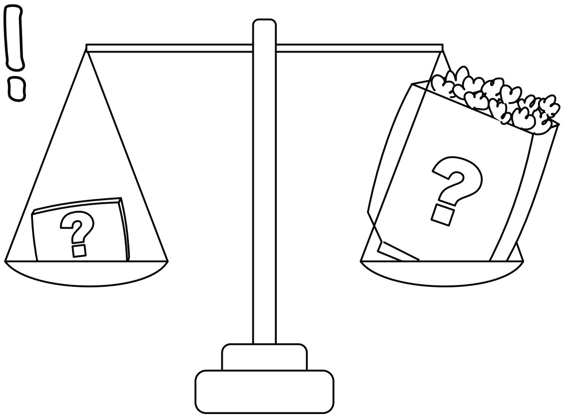
LABLE THE ARROWS WITH THESE WORDS:

sublimation
freezing
evaporation
melting
condensation
deposition

Solids keep their shape and volume. Liquids take the shape of their container, but the volume will stay the same. Gasses are super flexible! They will expand to fill whatever space they are in. Usually, solids are more dense that liquids, and liquids are more dense than gasses, but there is one compound where this rule doesn't hold! Solid water is less dense than liquid water. This is why ice floats.



Popping Predictions!



WHICH WILL WEIGH MORE? A BAG OF POPCORN THAT IS POPPED OR UNPOPPED? WRITE YOUR PREDICTION HERE:

No such thing as a wrong answer here!
Just make a guess and then experiment
to find out if your guess was correct!

YOU CAN TRY THIS YOURSELF BY MAKING A SCALE! ATTACH TWO UNPOPPED BAGS OF POPCORN TO EACH SIDE OF A RULER AND BALANCE IT. THEN POP ONE OF THE BAGS AND REATTACH IT. DOES THE RULER TIP MORE TO ONE SIDE OR THE OTHER? WHICH SIDE IS HEAVIER?

Record the weights that Science Mom & Math Dad measure during class:

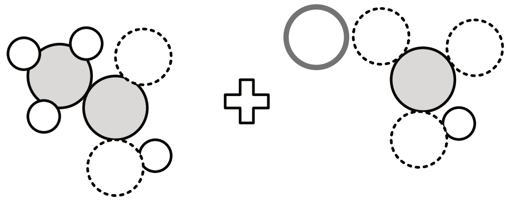
BEFORE POPPING:	*AFTER POPPING:*

YOUR DOODLE SPACE

Draw your favorite moment from class or write a cool fact!

Baking soda + vinegar

Count how many of each atom there are in each of the boxes. Record your observations in the charts below!



$C_2H_4O_2$

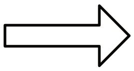
VINEGAR
(officially called
acetic acid)

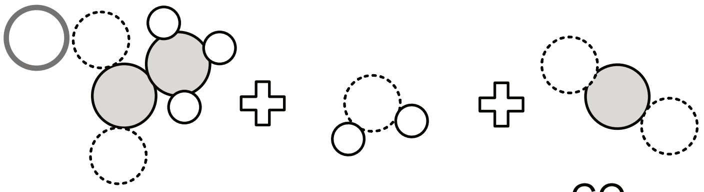
$NaHCO_3$

BAKING SODA
(officially called
sodium bicarbonate)

REACTANTS

How many sodium atoms?	◯	1
How many carbon atoms?	●	3
How many oxygen atoms?	◌	5
How many hydrogen atoms?	○	5





$C_2H_4O_2$

SODIUM
ACETATE
(A salt)

H_2O

WATER
(Can officially be
called "dihydrogen
monoxide")

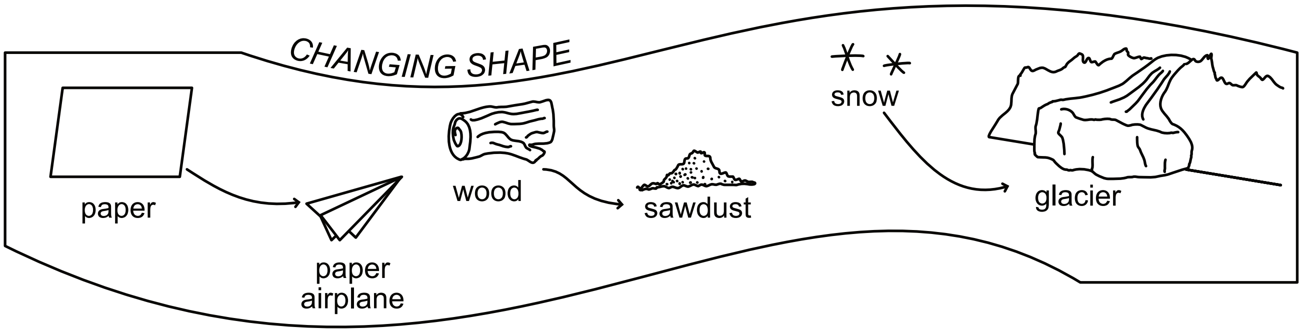
CO_2

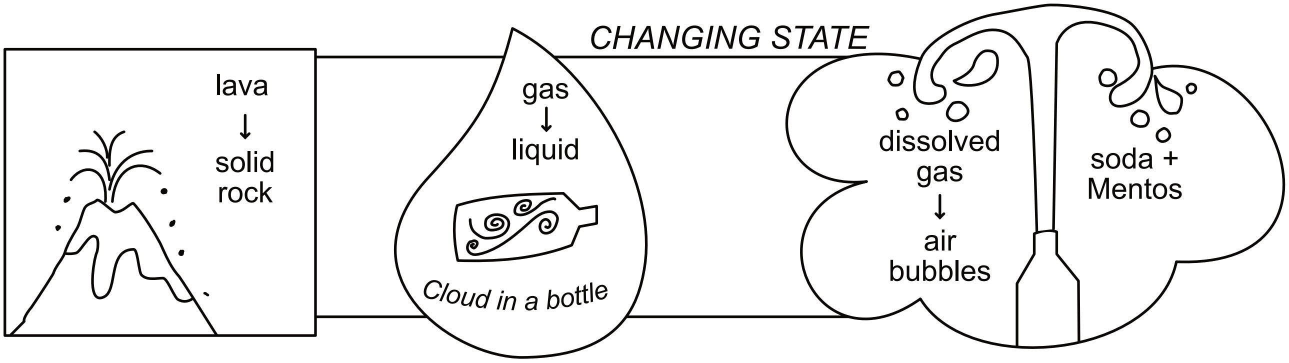
CARBON
DIOXIDE
(A gas! This is
what makes the
bubbles.)

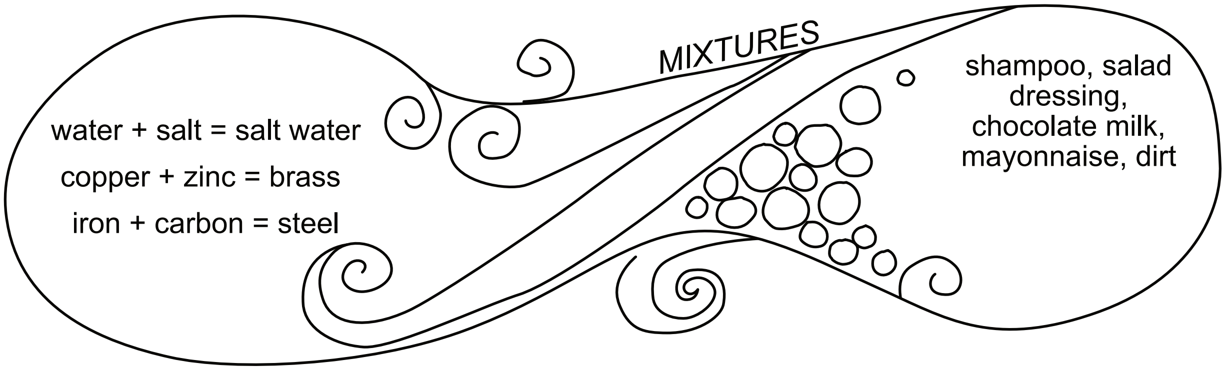
PRODUCTS

How many sodium atoms?	◯	1
How many carbon atoms?	●	3
How many oxygen atoms?	◌	5
How many hydrogen atoms?	○	5

Physical Changes







Your notes:

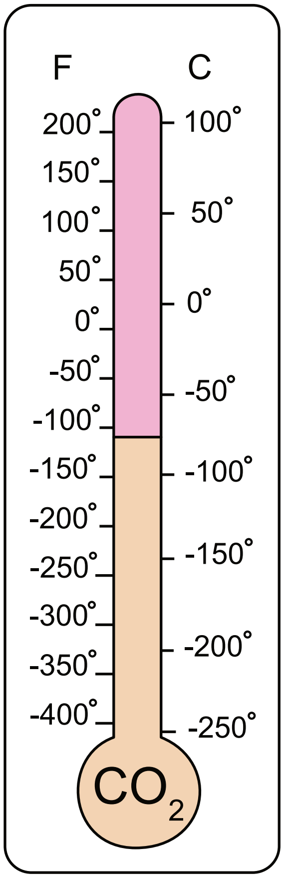
Liquids are rare and actually kind of weird

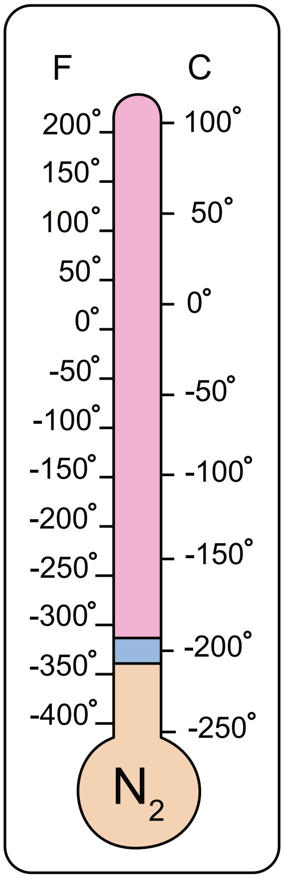
YOUR DOODLE SPACE

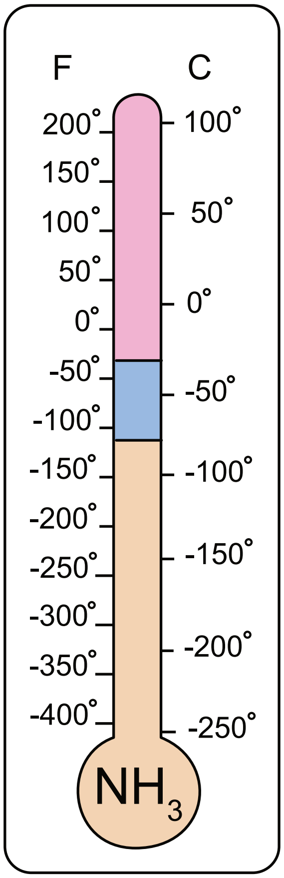
Draw your favorite moment from class or write a cool fact!

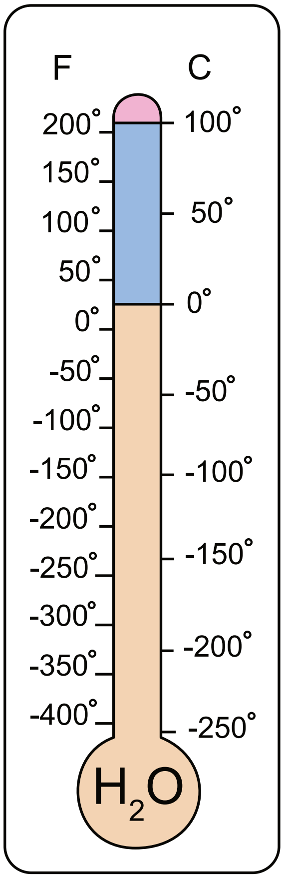
Choose colors to represent solid, liquid, and gas, then color the thermometers to show when the substance will exist in which state of matter. (Hint: CO_2 doesn't exist as a liquid on Earth unless you increase the pressure a lot!)

🟧 SOLID
🟦 LIQUID
🟪 GAS





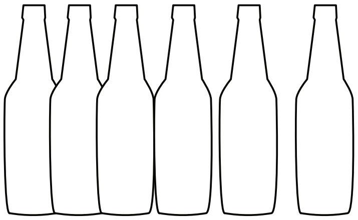




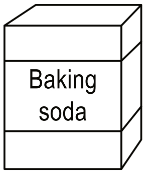
Hands-on Activity

FIZZING FUN!

MATERIALS:

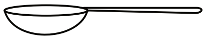


6 BOTTLES OF SODA IN PLASTIC CONTAINERS WITH NARROW TOP

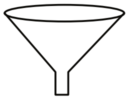


BAKING SODA

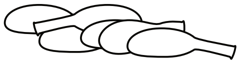
3 PACKAGES OF POP ROCKS CANDY



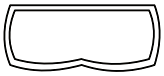
A MEASURING SPOON



A FUNNEL



6 BALLOONS



EYE PROTECTION

First, blow up each of the balloons once or twice to stretch them out, letting the air back out afterward.

Record your observations here:

1	Place a balloon over the top of the bottle so that it is firmly in place. Put on eye protection. Then shake the bottle and record what happens to the balloon.	
2	Use the funnel to pour 2 TBL of baking soda in the balloon. Attach the balloon securely around the mouth of the soda bottle and then tip the ballon so that the baking soda pours from the balloon into the bottle.	
3	Repeat the procedure with the Pop Rocks in a new balloon added to a new bottle of soda. Record your observations.	
4	Mix baking soda and Pop Rocks together in a new balloon and put it over a new bottle of soda. Record your observations.	

With the last two bottles, experiment! You get to decide what to try:

5		
6		

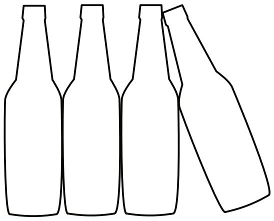
QUESTION: *Why did the balloons expand?* ______________________________

Write your answer on another piece of paper!

Hands-on Activity

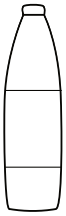
MORE FIZZING FUN!

MATERIALS:



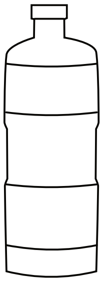
4 EMPTY BOTTLES

6 ALKA-SELTZER TABLETS

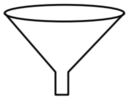


VINEGAR

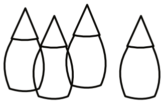
WATER



OIL



A FUNNEL



FOOD COLORING

1. Pour the same amount of water into each of three of your bottles: hot water in one, room temperature water in the second, cold in the third. The level of the water should take up about ¼ of the volume of the bottle.
2. Pour the same amount of vinegar into the fourth bottle.
3. Use your funnel to SLOWLY pour vegetable oil into all 4 bottles until they are mostly full. You may want to wait a few minutes for the oil and water to separate after this step.
4. Add 4 drops of food coloring to each bottle. Watch and observe how it interacts with the oil versus the water.
5. Break a seltzer tablet in half and add to each bottle, at the same time if possible. Watch and record your observations, especially how long the tablets took to react.
6. After all the bubbles have stopped, repeat the reaction. Record your observations and answer the questions.
7. With the final two bottles - YOU get to decide what to do! Which experiment will you try? What changes will you make?

QUESTIONS:

HOW DID THE TEMPERATURE OF THE WATER AFFECT THE LAVA LAMP? WHICH ONE WAS THE MOST DRAMATIC? WHICH LASTED THE LONGEST?

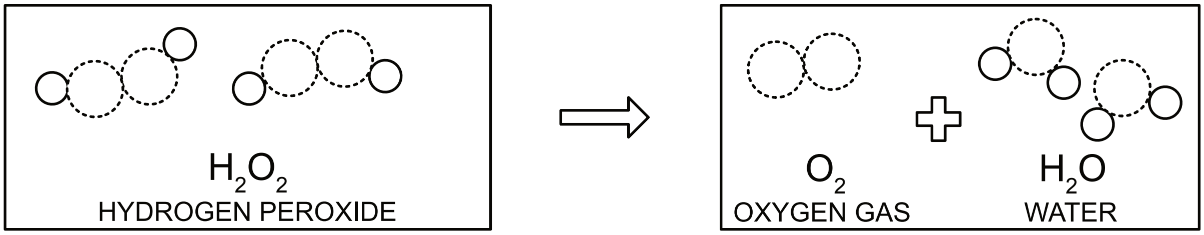
HOW DID THE VINEGAR LAMP COMPARE TO THE WATER LAMPS?

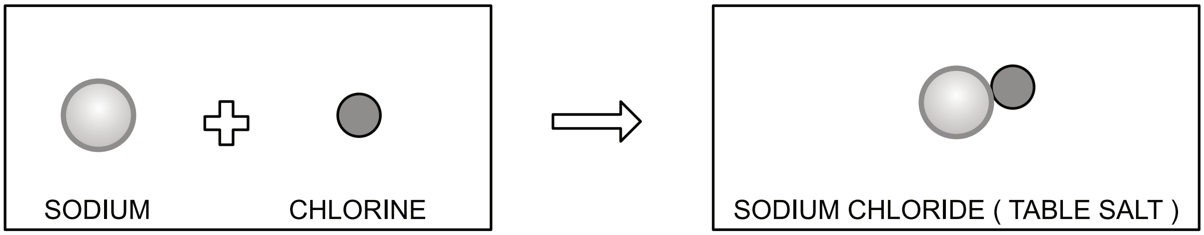
Chemical Reactions

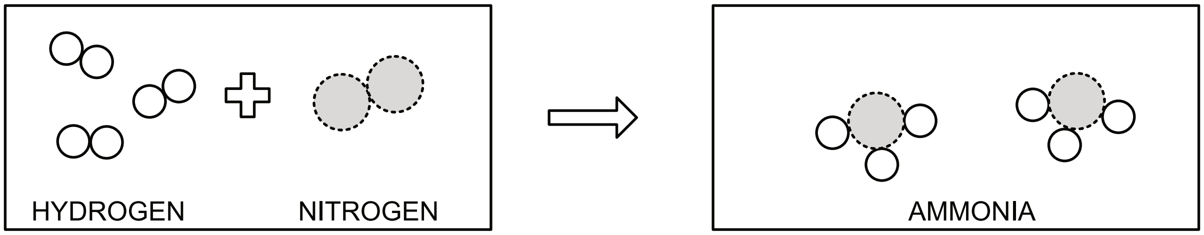
FILL IN THE BLANKS USING THESE WORDS:

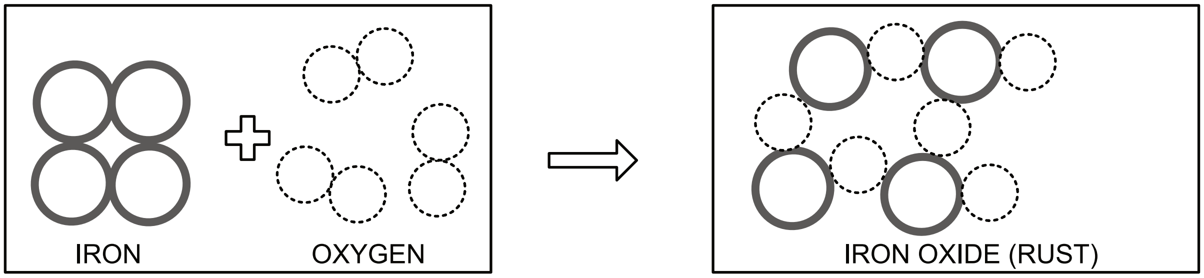
products chemical reactants
physical molecules

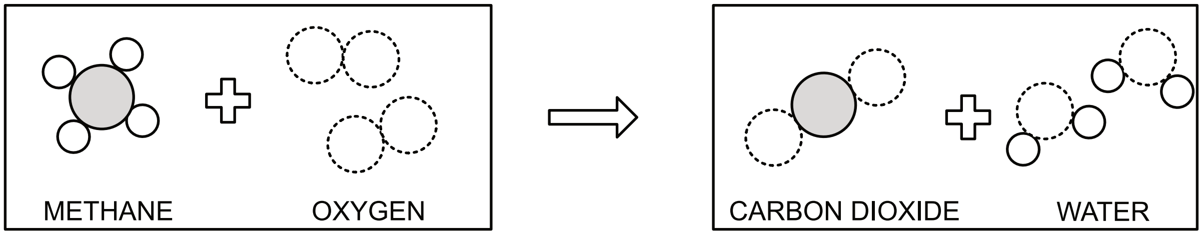
In a chemical reaction, new molecules are formed. The molecules that existed BEFORE the reaction are called the reactants. The molecules that exist AFTER the reaction are called the products. In a physical reaction, matter might change its shape or state, but no new molecules are formed.



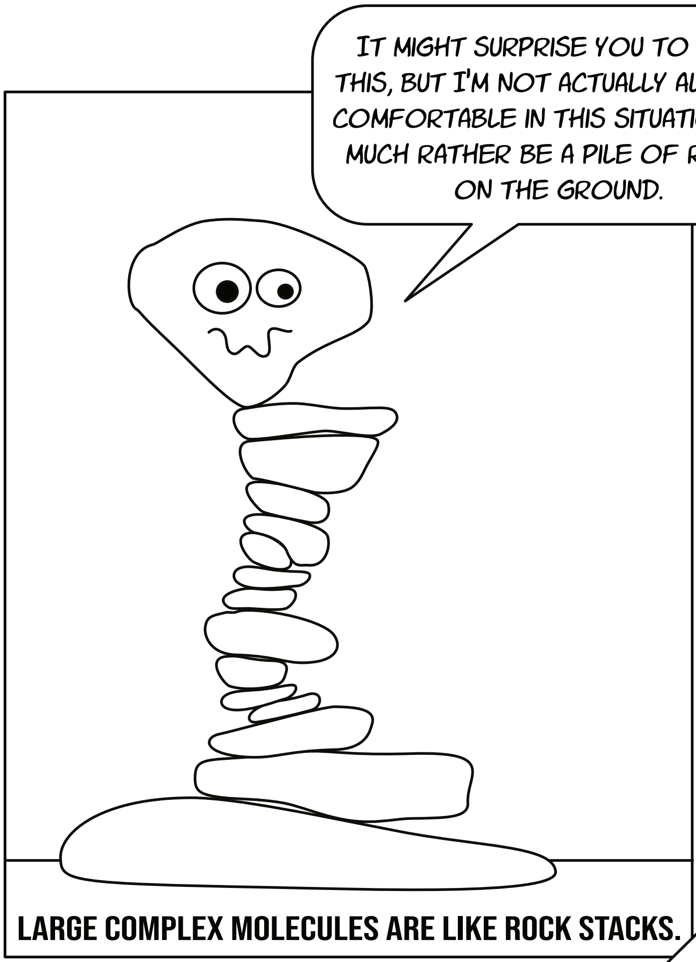




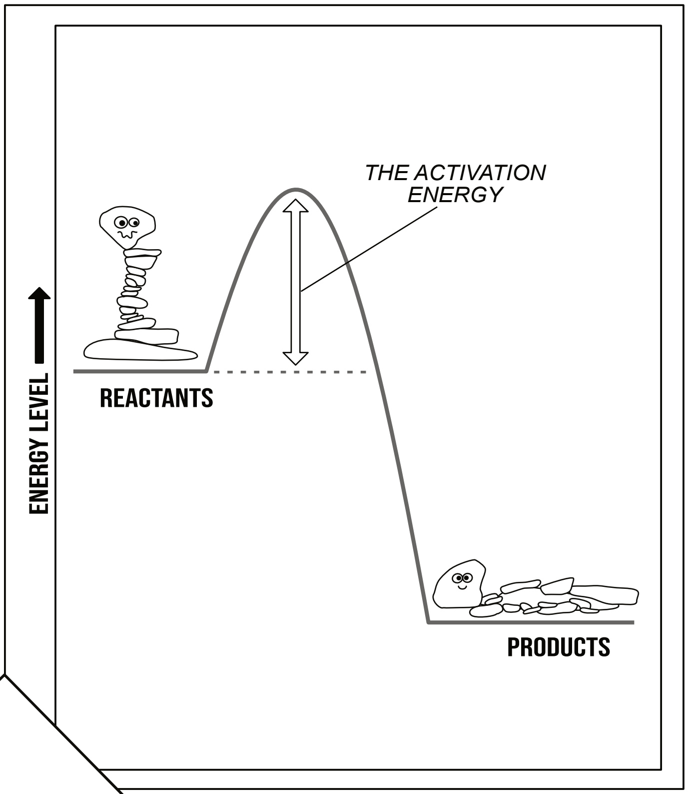




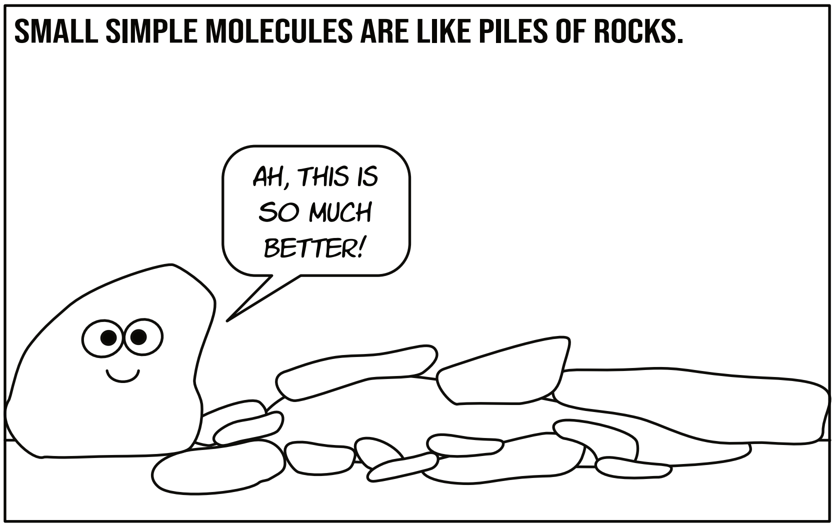
Why do things react?



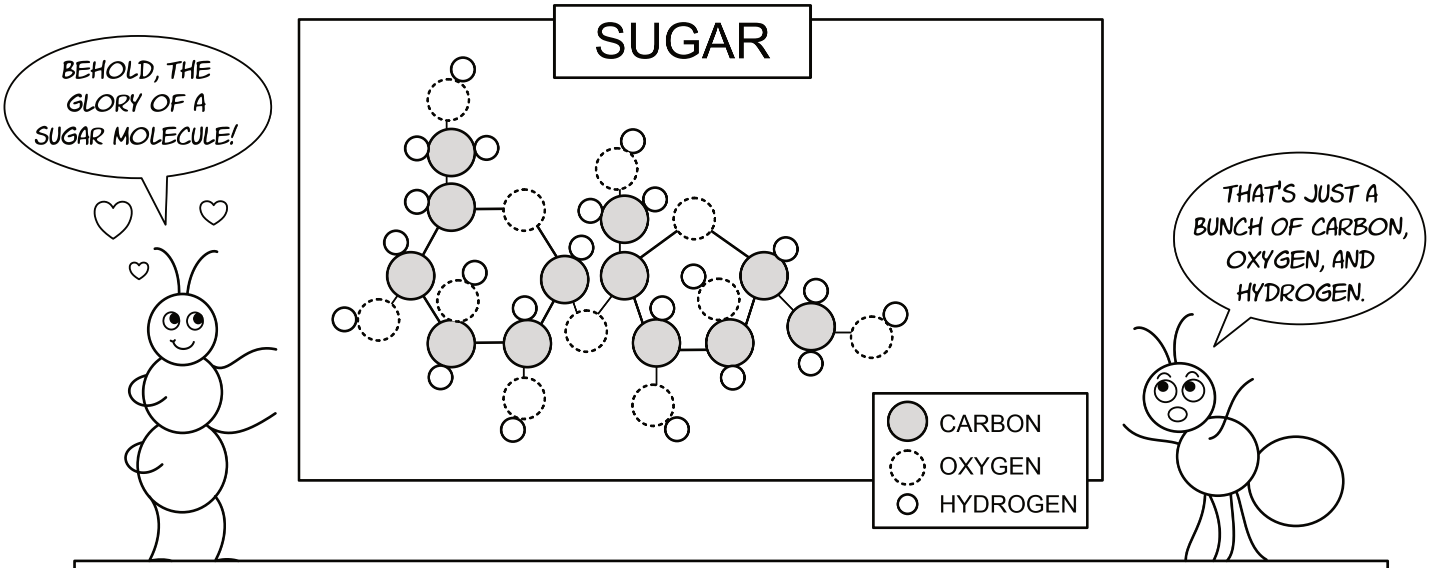
LARGE COMPLEX MOLECULES ARE LIKE ROCK STACKS.



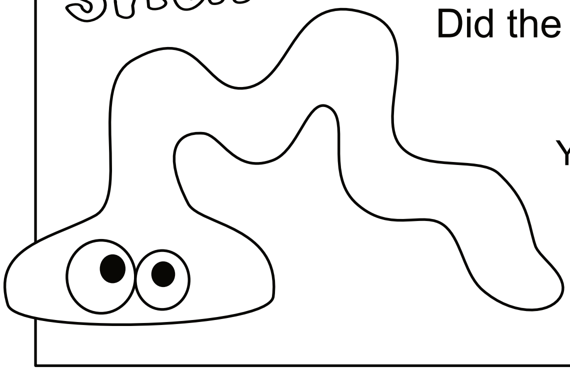
Your notes:



SMALL SIMPLE MOLECULES ARE LIKE PILES OF ROCKS.



sugar snake

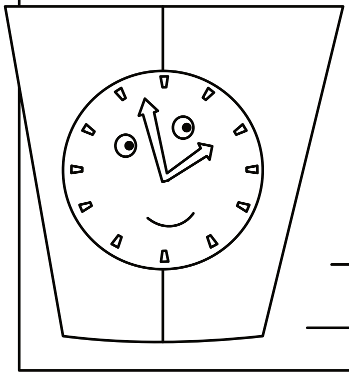


What did the sugar turn into? A dark tube of carbon!

Did the reaction create heat? [✓ YES] NO ?

Your notes:

Iodine clock

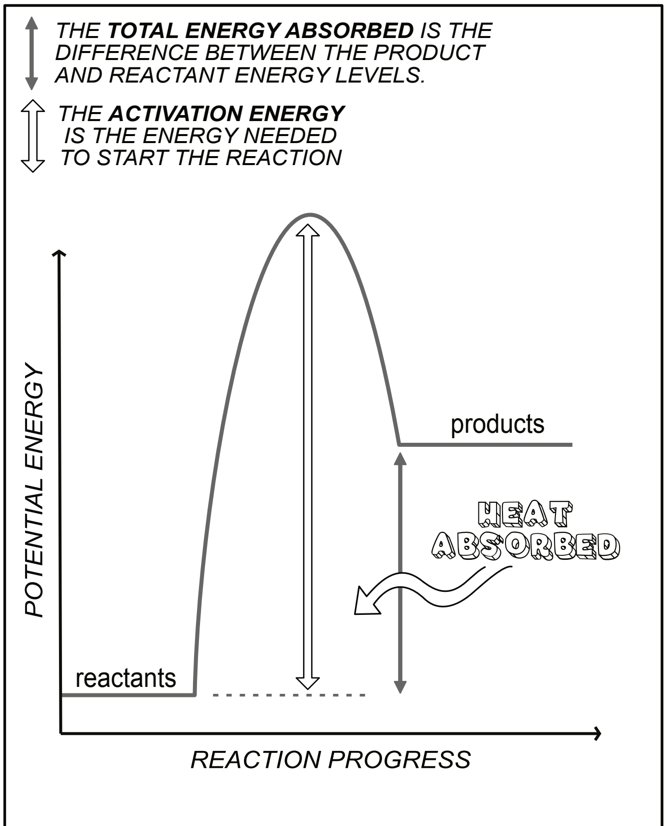


What made the color change? Iodine reacting with starch.

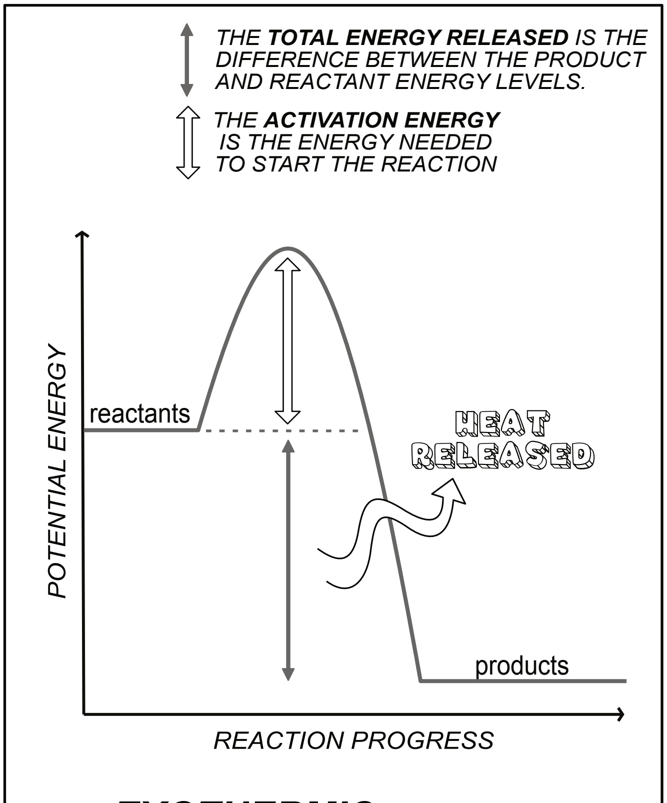
Did the reaction create heat? [✓ YES] NO ?

Your notes:

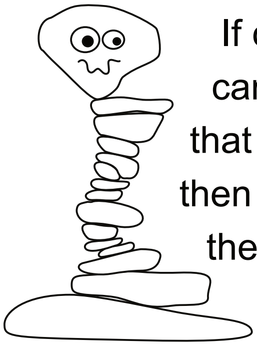
How much energy?



ENDOTHERMIC *REACTION*



EXOTHERMIC *REACTION*

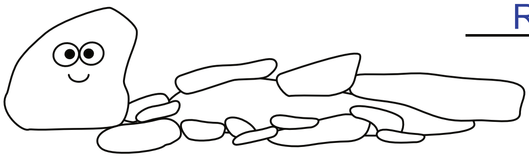


If complex molecules are like carefully balanced rock stacks that want to turn into rock heaps, then how do the stacks get built in the first place? It takes energy! If energy is put INTO a REACTION and the PRODUCTS have a **higher** energy state than the reactants, this is an ENDOTHERMIC reaction. It **absorbs** ENERGY. Evaporating water or dissolving ammonium chloride in water are examples of these "energy-requiring" reactions. When these reactions happen, the temperature DROPS!

FILL IN THE BLANKS USING THESE WORDS:

rises
reaction
drops
energy
products
endothermic
~~determines~~
reactants
exothermic
releases

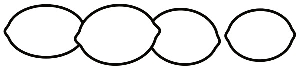
If a reaction PRODUCES energy, and the products have a **lower** energy state than the REACTANTS, this is an EXOTHERMIC reaction. It RELEASES energy. A burning match and rusting metal are examples of these "energy-producing" reactions. When these reactions happen, the temperature RISES!



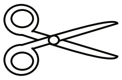
Hands-on Activity

LEMON BATTERY

MATERIALS:

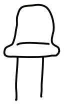


LEMONS (OR POTATOES, OR VINEGAR IN AN ICE CUBE TRAY)

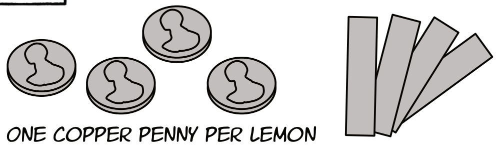


SCISSORS OR KNIFE TO MAKE HOLES IN THE LEMON OR POTATO

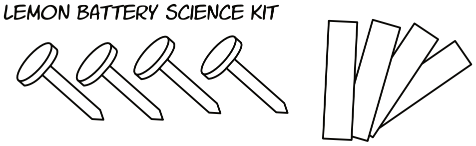
ALLIGATOR CLIPS OR COPPER WIRE



LED DIODE



ONE COPPER PENNY PER LEMON OR ONE PIECE OF COPPER FROM A LEMON BATTERY SCIENCE KIT

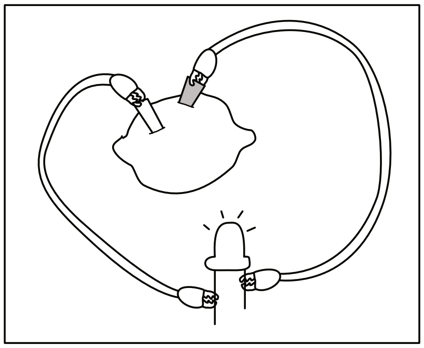


ONE GALVANIZED NAIL PER LEMON OR ONE PIECE OF ZINC FROM A LEMON BATTERY SCIENCE KIT

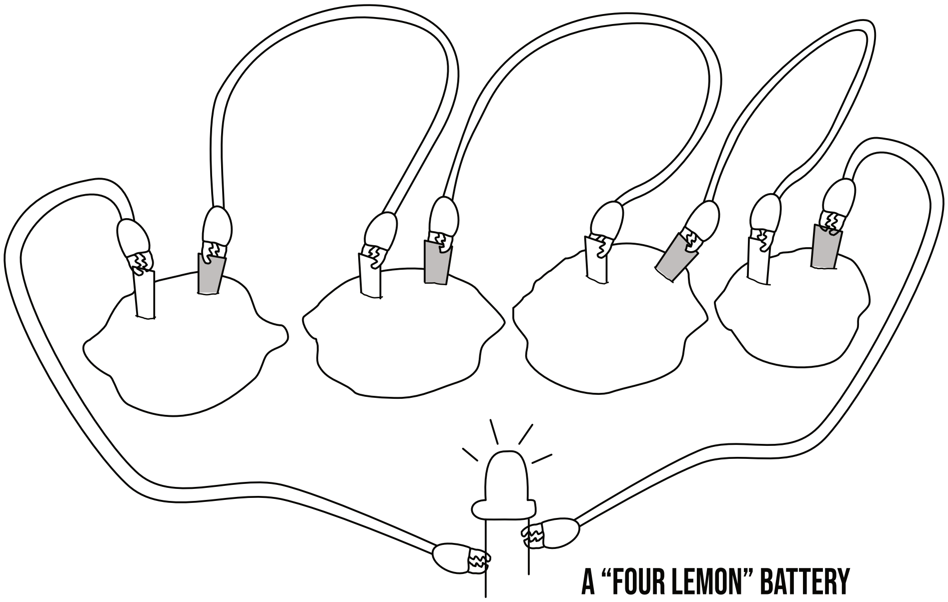
INSTRUCTIONS:

Prepare two lemons or more (The more you have the stronger your battery is. For getting an LED light to light up, we recommend at least two. Potatoes or vinegar in an ice cube tray can be used instead of lemons.)

1. Squeeze and roll the lemons for several minutes. The individual segments of the lemon need to break up enough that a current can run from one end to the other.
2. Make two slits on either side of the lemon and insert the penny or copper into one slit and the galvanized nail or zinc into the other. Make sure that enough of the metal is sticking out of the lemon that you'll be able to attach the alligator clips or wire. Repeat with the remaining lemons.
3. To build the circuit between the lemons, attach one alligator clip around the zinc from the first lemon and connect it to the copper in the next lemon. If using multiple lemons, continue this pattern with each of the lemons.



A “SINGLE LEMON” BATTERY



A “FOUR LEMON” BATTERY

CONTINUED ON NEXT PAGE:

LEMON BATTERY CONTINUED

For electricity to flow through the wires, the circuit needs to form a loop. If you connect the copper in the first lemon to the zinc in the last lemon, then you will have an electric current flowing through the wires - but this current is so small you won't be able to feel it or see it.

4 Attach the ends to the LED light or clock you are trying to power. Touch the wire attached to the first penny or copper to the **long leg** of your LED light. Simultaneously touch the wire attached to the nail of the last lemon to the short leg of the LED light. If you need help differentiating the long leg from the short look for a "flat spot" on the bottom edge of light. That is where you will find the short leg.

If your first attempt doesn't work, try adjusting the number of lemons or vinegar cells you are using.

What happens if you try powering the light with 1 lemon versus 2?

2 lemons provide more charge. The light will (most likely) not turn on with 1 lemon, but will light up if two lemons are used.

What would happen if you had attached the copper wire from one penny to another penny and one nail to another nail instead of following the coin-nail-coin-nail pattern?

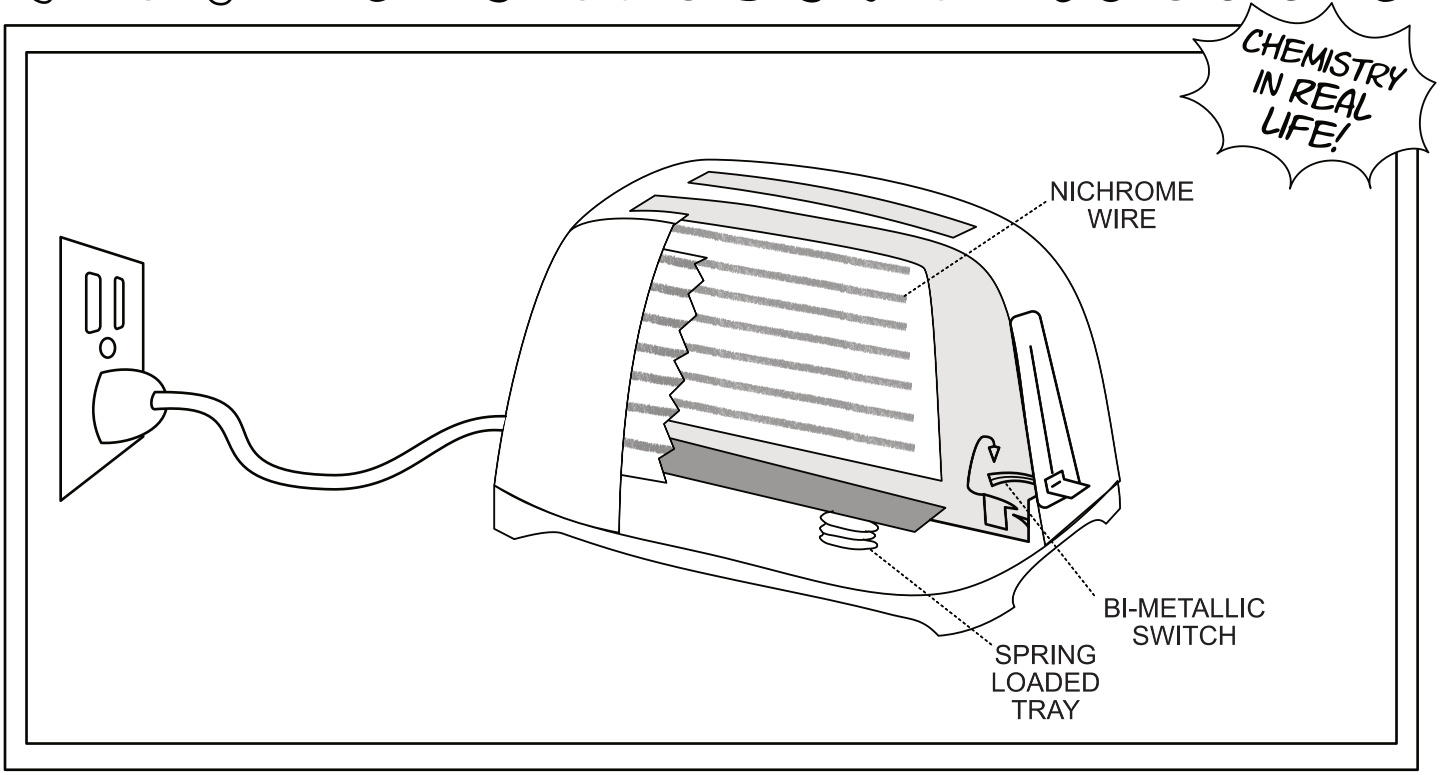
No light! You will only get a battery if the copper is connected to zinc.

What else do you think would serve as a good materials for this experiment? Are there any other conductors (the alligator clips or copper wire) or electron sources (copper and zinc) that you could use?

The copper and zinc were anodes and cathodes in our experiment, and with real life batteries there are lots of different anodes and cathodes. The materials in alkaline batteries (standard AA and AAA batteries) are different from lithium batteries, etc.

Sometimes it can be tricky to get a lemon battery to work. Did you run into any trouble with your experiment? If so, what did you try?

HOW a TOASTER works



FILL IN THE BLANKS USING THESE WORDS:

light heat glow chemical sugars
filaments energy reactions proteins

When the toaster is turned on, _ENERGY__ passes from the outlet to the toaster in the form of electricity. The electric current passes through thin __FILAMENTS__ that are uniformly spaced around the toaster slot. The filaments are specially designed to __HEAT__ up when electricity passes through them. They get so hot that they __GLOW__ bright red! The electrical energy has been converted into heat and __LIGHT__. The steady supply of heat causes _CHEMICAL__ __REACTIONS__ to happen on the surface of the bread. The heat causes __PROTEINS__ and __SUGARS__ to combine together, forming new molecules that change the color and flavor of the bread, turning it into delicious toast.

Your notes: ____

____

____

____

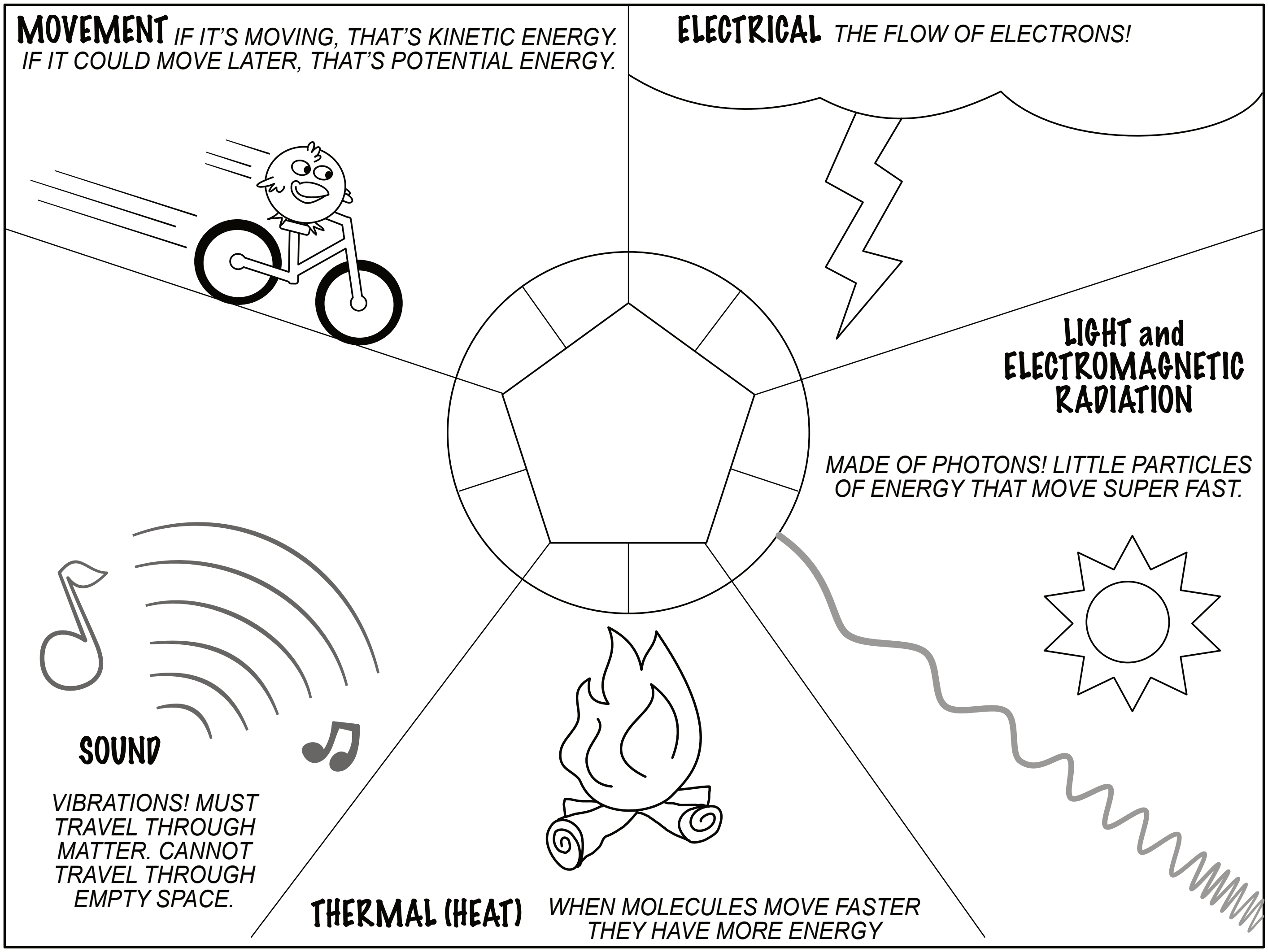
____

ENERGY

Chemical reactions can create electricity (this is how batteries work!), light and heat, sound (think fireworks), and movement too. The LAW OF CONSERVATION tells us that energy cannot be created or destroyed, instead it's transferred from one form to another.

YOUR DOODLE SPACE

Draw your favorite moment from class or write a cool fact!



MOVEMENT *IF IT'S MOVING, THAT'S KINETIC ENERGY. IF IT COULD MOVE LATER, THAT'S POTENTIAL ENERGY.*

ELECTRICAL *THE FLOW OF ELECTRONS!*

LIGHT and ELECTROMAGNETIC RADIATION

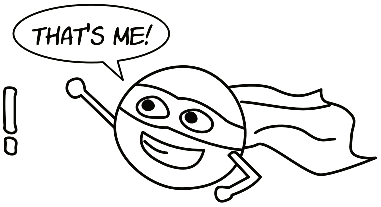
MADE OF PHOTONS! LITTLE PARTICLES OF ENERGY THAT MOVE SUPER FAST.

SOUND

VIBRATIONS! MUST TRAVEL THROUGH MATTER. CANNOT TRAVEL THROUGH EMPTY SPACE.

THERMAL (HEAT) *WHEN MOLECULES MOVE FASTER THEY HAVE MORE ENERGY*

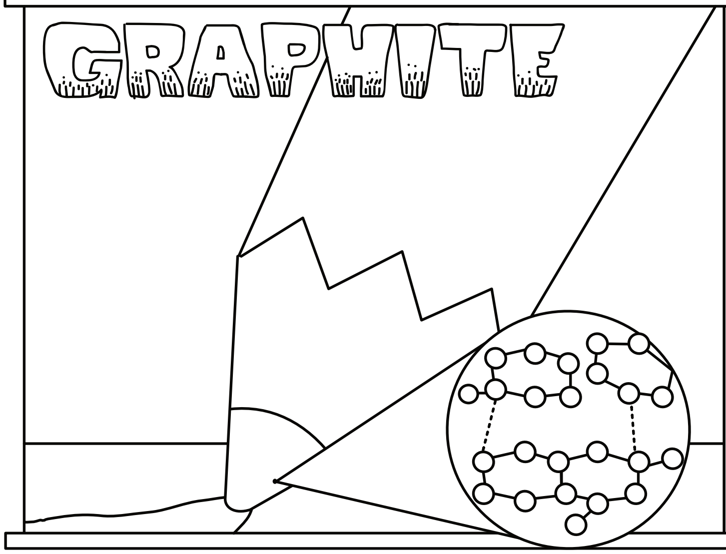
CARBON THE BUILDING BLOCK OF LIFE!

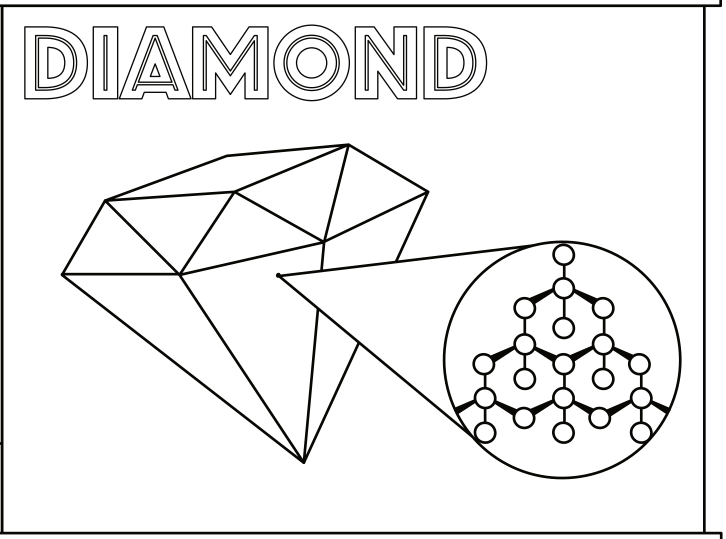


FILL IN THE BLANKS USING THESE WORDS:

oxygen
four
backbone
abundant
electrons
unpaired
graphite
carbon

Carbon is the second most ABUNDANT element in the human body. (The most abundant element is OXYGEN.) It's the BACKBONE of all the molecules that cells are made of. Because it has four ELECTRONS that are UNPAIRED, carbon likes to form FOUR bonds with other atoms. Soft black pencil lead called GRAPHITE is made of carbon. The hard clear crystal of a diamond is made of CARBON too.

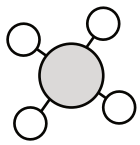




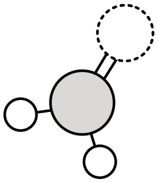
YOUR DOODLE SPACE

Draw your favorite moment from class or write a cool fact!

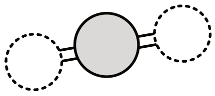
CARBON LIKES TO FORM 4 BONDS:



4 single bonds

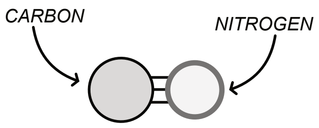


1 double & 2 single



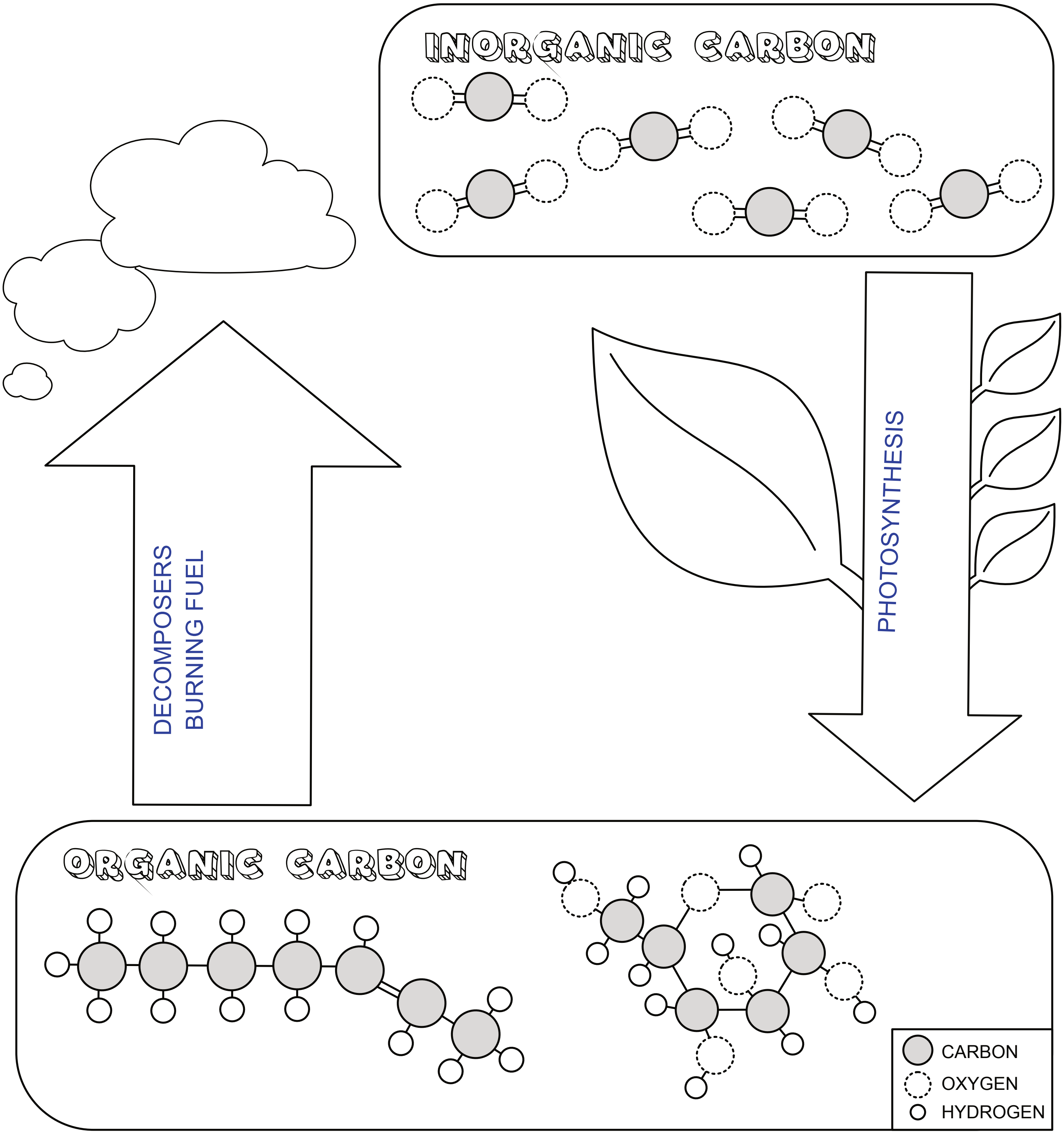
2 double bonds

But sometimes it only has three bonds:



LOOK OUT! THAT'S CYANIDE, WHICH IS VERY POISONOUS.

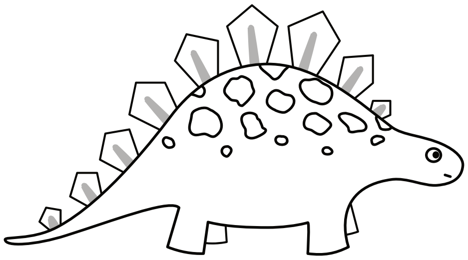
The Carbon Cycle



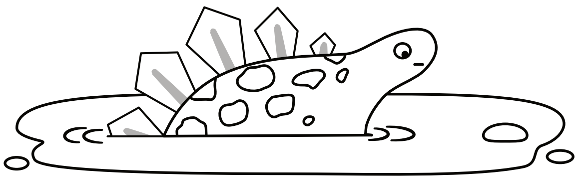
Carbon really just has two forms: carbon dioxide (inorganic) and everything else (organic). When organic carbon is eaten or burned, energy is released and the carbon is converted into carbon dioxide. When algae or plants perform photosynthesis, carbon dioxide is converted back into an organic form. The same carbon atoms can travel in a huge circle from gas to organic matter to gas and back again. This is called the carbon cycle.

WHAT'S A FOSSIL FUEL?

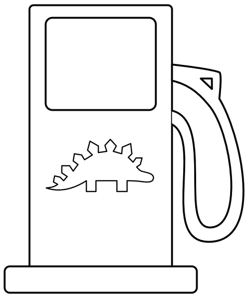
STATES OF DINO-MATTER



SOLID

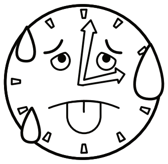


LIQUID



GAS

You've probably seen jokes that credit dinosaurs as the source of gasoline, but this isn't quite accurate. Fossil fuels like petroleum, oil, and natural gas come from organic matter that lived during the Carboniferous period, which occurred several million years before the first dinosaurs walked on Earth. The carbon in gasoline once existed in plants, algae, invertebrates, and fish, but not dinosaurs.

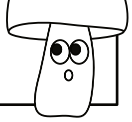


How long will fossil fuels last?
The answer is 50 years or forever,
depending on who you ask.
On this thing more people agree:
the more fuel we burn, the warmer the planet will be.

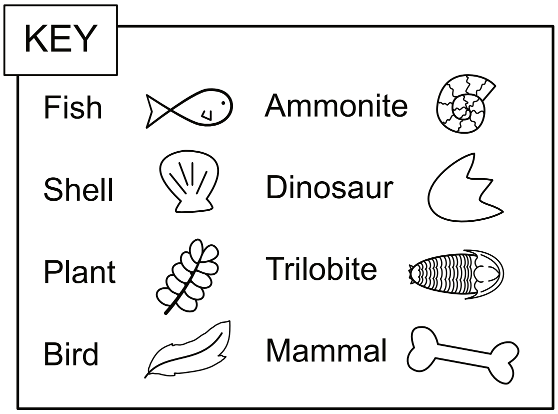
YOUR DOODLE SPACE

Draw your favorite moment from class or write a cool fact!

During the Carboniferous period, fungi hadn't yet developed the ability to break down cellulose, the main ingredient of wood. Without these decomposers, an enormous amount of plant material accumulated.



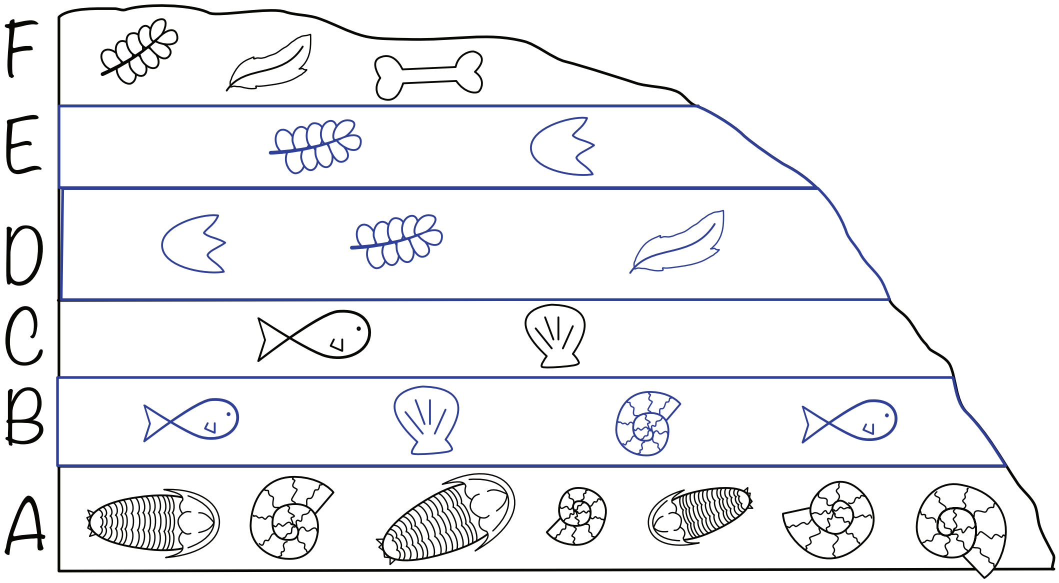
Your notes:



ROCK LAYERS

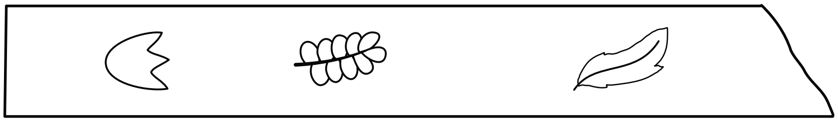
You have joined a team of stratigraphers and paleontologists who are studying the layers of rock and fossils of this site!

Your job is to complete the timeline and rock layer chart by studying the information available. What do the fossils in each layer tell you about each period of time? According to the timeline, what fossils would you find in each missing layer?



Cut out the rock layers and timeline boxes below. Can you paste them over the correct question marks?

Ammonites and trilobites are abundant in the shallow waters.



Mammals and birds live in the area.

Ammonites and trilobites have gone extinct.

ROCK LAYERS continued...

TIMELINE

Layer A

Ammonites and trilobites are abundant in the shallow waters.

Layer B

Fish and shellfish are now abundant in the waters.

Layer C

Ammonites and trilobites have gone extinct.

Layer D

Dinosaurs make their first appearance.

Layer E

Lack of fish fossils indicates that the water has dried up.

Layer F

Mammals and birds live in the area.

Which is the oldest layer? Layer A with the ammonites and trilobites

What layers indicate that water was present? What evidence do you have to support this?

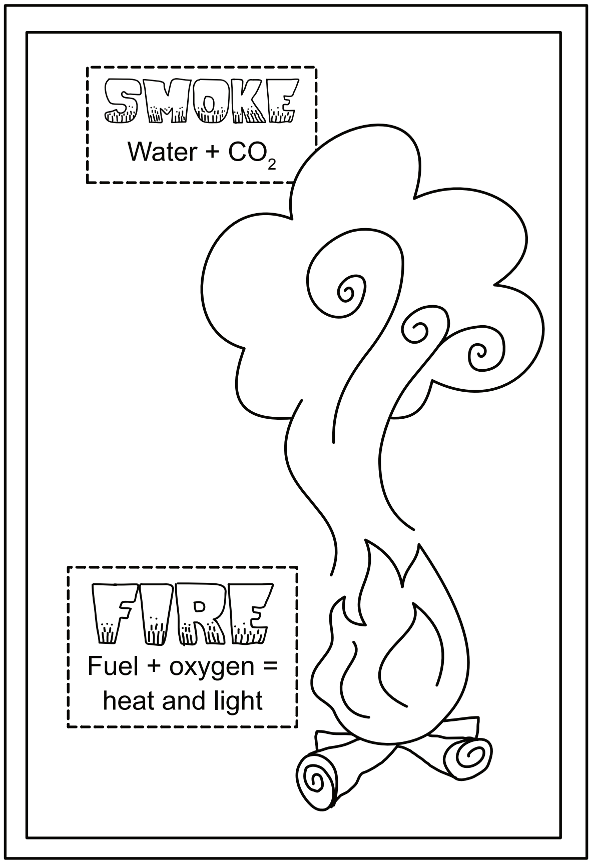
Layers A, B, and C because all of the animal fossils in those layers are animals that lived underwater.

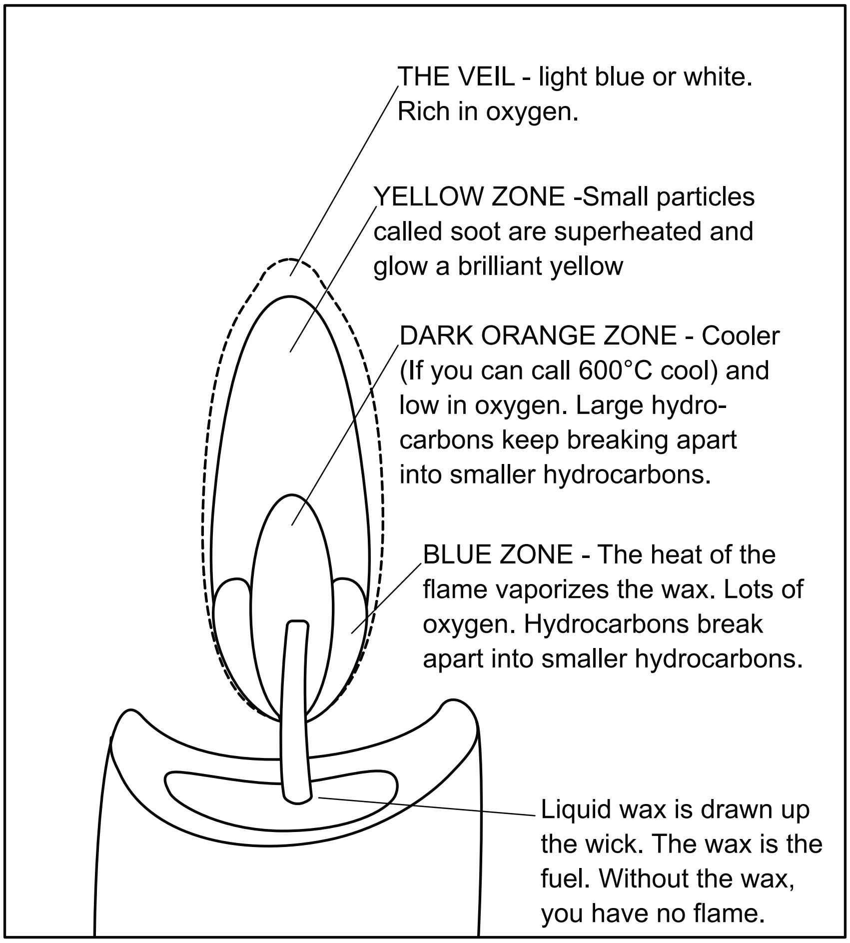
Why weren't mammal bones found in layer B?

Layer B is so old that mammals didn't yet exist!

Cut out the rock layers and timeline boxes on the other side of this paper and see if you can match them over the correct question marks!

What is FIRE? All about combustion




THE VEIL - light blue or white. Rich in oxygen.

YELLOW ZONE -Small particles called soot are superheated and glow a brilliant yellow

DARK ORANGE ZONE - Cooler (If you can call 600°C cool) and low in oxygen. Large hydrocarbons keep breaking apart into smaller hydrocarbons.

BLUE ZONE - The heat of the flame vaporizes the wax. Lots of oxygen. Hydrocarbons break apart into smaller hydrocarbons.

Liquid wax is drawn up the wick. The wax is the fuel. Without the wax, you have no flame.

FILL IN THE BLANKS USING THESE WORDS:

oxygen combustion water
hydrocarbons carbon

Fire is the result of a chemical reaction called ____COMBUSTION____. Three things must be present for fire: __OXYGEN__, fuel, and heat. When a fuel like wood meets oxygen, the ____HYDROCARBONS____ in the wood combine with oxygen to form __CARBON__ dioxide and __WATER__. Water is one of the main ingredients of smoke. It is also the main ingredient of clouds. If a forest fire gets large enough, it can produce a pyrocumulus cloud: a cloud so big that it makes rain and lightning.

Your notes: ____

____

____

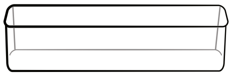
____

____

Hands-on Activity

Build a levee

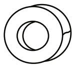
MATERIALS:



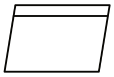
TUPPERWARE OR GLASS CONTAINER

WATER

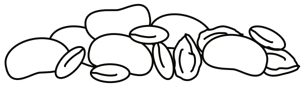
FINE-GRAINED BUILDING MATERIAL (SUCH AS 1 CUP FLOUR MIXED WITH 1 TBSP COCOA POWDER)



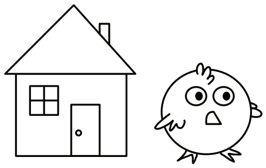
DUCT TAPE OR ELECTRICAL TAPE



PLASTIC BAG



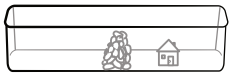
COARSE-GRAINED BUILDING MATERIAL SUCH AS NUTS, DRIED BEANS, OR DRIED FRUIT.



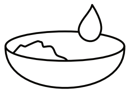
(OPTIONAL) SMALL TOYS TO REPRESENT THE TOWN

INSTRUCTIONS:

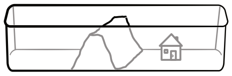
1. First, arrange a small pile of "large particle size" material in the center of your Tupperware container. Slowly pour in 1/2 cup of water on one side of your levee and observe what happens. Does the water stay on one side of the levee or does it leak through to the other side?



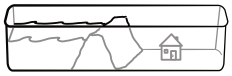
2. Next, prepare your "small particle size" material. If using flour and cocoa powder, mix one cup of flour with 1 tbsp of cocoa powder and 1/2 cup hot water. Add the water slowly and mix well, kneading it into a stiff dough.



3. Form half of your dough into a levee and arrange some small objects on the other side to represent the town. If desired, you can press some of the large particle-size material into the sides of the levee to reinforce it.



4. Slowly pour water on the side of the Tupperware that is opposite of your town and observe. It should be keeping all the flood waters away from the town!



5. Make your levee fail. You could create a small hole with a toothpick, or keep adding water until the water pours over the top.

6. Dry out your container and use the other half of the dough to make a new levee!

A FEW FACTS ABOUT LEVEES

A levy is more than a big pile of dirt, although at first glance that's pretty much what it looks like!

For a levy to work well, it has to be made out of the right material and have the correct slope. In general, finer materials such as clay and silt will do a better job of holding back water than coarse materials like sand, gravel, or rocks.

There are two main ways that levies can fail: one is by being overtopped (the water flows over the top of the levee and then begins to erode it) The other is by breaching (basically a hole forms in the levee and then a big portion of it breaks). Before a levee breaches, there will often be a "sand boil." Water will begin flowing through a weaker spot in the levee and out the other side.

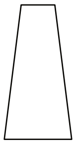
If you live in the United States, you might think that cities along the Mississippi River are the only ones that have levies. But there are levees in all 50 states and more than 40% of the US population lives in a county with at least one levee. When they work, we hardly notice them. When they fail, the flooding can be catastrophic.

Build a levee continued...

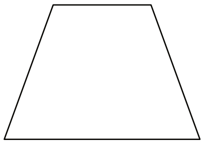
Would your flour levee hold the water back indefinitely/forever or would the water eventually leak through? What could you do to make this levee stronger?

The water would eventually leak through. Adding in cement or other material to make it water proof would help the levee last longer.

Which slope would make the strongest levee? Super steep, medium, or broad? Explain why:



SUPER STEEP



MEDIUM

BROAD / GENTLE SLOPE

A broad/gentle slope creates a better levee because there is more earth or other material blocking the water.

Pretend you are in charge of building two real-life levees, what are some different considerations to take into account for designing an urban vs rural levee? Would you need to you do anything differently for the urban levee (protects a city area with stores, houses, and other buildings) versus an agricultural levee (protects fields)?

Fields of land can recover from flooding easily, but houses and buildings receive a lot of damage (or are washed away entirely). There is less risk of damage if a rural levee breaks.

What are some other natural earth process that affect humans? What kind of designs and solutions have we come up with to cope with them?

Earthquakes: Engineering buildings to be able to withstand shaking.

Tsunamis: posting signs for tsunami evacuation areas and moving schools, nursing homes, and other vulnerable buildings to higher ground.

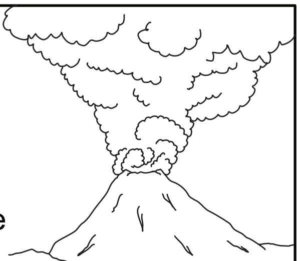
Wildfires: prescribed burns or controlled logging to reduce fuel loads in forests.

Hurricanes/severe storms: weather forecasting and evacuation orders.

The chemistry of LAVA

Oxygen and silicon are the most abundant elements in lava.

They combine to form silica, and the amount of silica determines what type of lava you have!

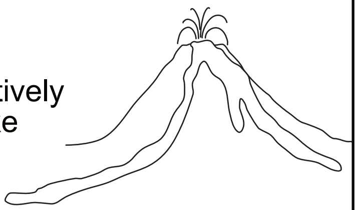


HIGH SILICA

VS

LOW SILICA

Lots of silica produces pale rocks like rhyolite. When melted, this type of lava is super thick and tends to be explosive, like the eruption of Mount St Helens.



Low silica makes dark basalt rock and is relatively runny when melted, like the slowly oozing pahoehoe lava in Hawai'i.

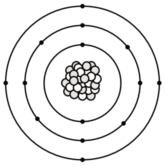
How hot is lava?

It all depends on the type of rock and what minerals it contains. Liquid rock (usually called lava) can be as cool as 700 °C and as hot as 1,200 °C. *That's 1300 - 2200° Fahrenheit!*

Your notes: ______________________

Silicon

Element

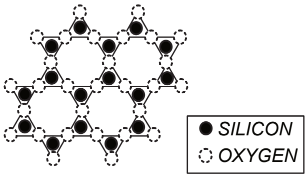


14 PROTONS
14 NEUTRONS
14 ELECTRONS

The second most abundant element on Earth after oxygen. Very rarely found in pure form, it loves to bond with oxygen. Widely used in electronics, especially computer chips.

Silica

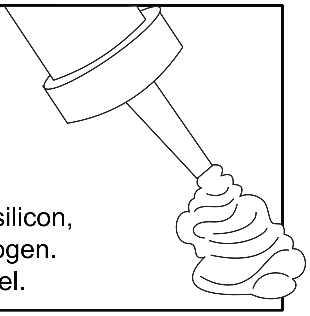
Compound



A compound of silicon and oxygen, most often a crystal of SiO_2. Quartz is silica. The mineral is also found in sand, glass, and many other rocks.

Silicone

Compound



A polymer (long chain) of silicon, oxygen, carbon, and hydrogen. Can be a solid, liquid, or gel.

YOUR DOODLE SPACE

Draw your favorite moment from class or write a cool fact!

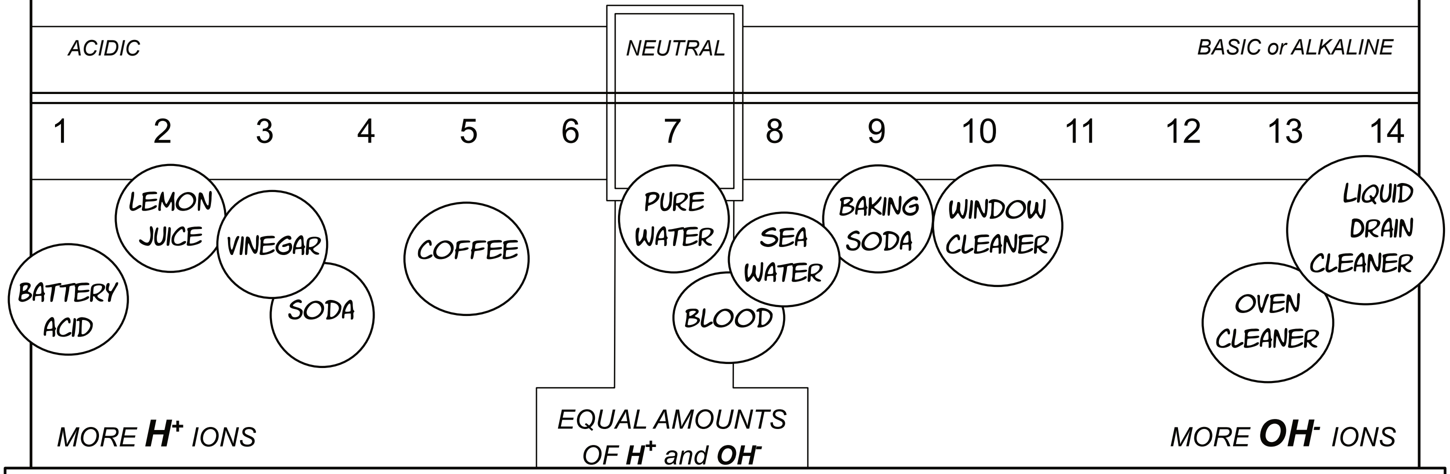
ACIDS & BASES

FILL IN THE BLANKS USING THESE WORDS:

basic
OH^-
pure
acidic
H^+
~~molecules~~
hydrogen
water

pH stands for "potential of HYDROGEN" or "power of hydrogen." It is a scale used to measure how ACIDIC or basic a solution is. At room temperature, PURE WATER is neutral with a pH of 7. At neutral pH, the amount of hydrogen ions (H^+) equals the number of hydroxide ions (OH^-). In an acidic solution, there are more H^+ ions than OH^- ions. In a BASIC solution, there are more OH^- ions than H^+ ions.

The pH scale



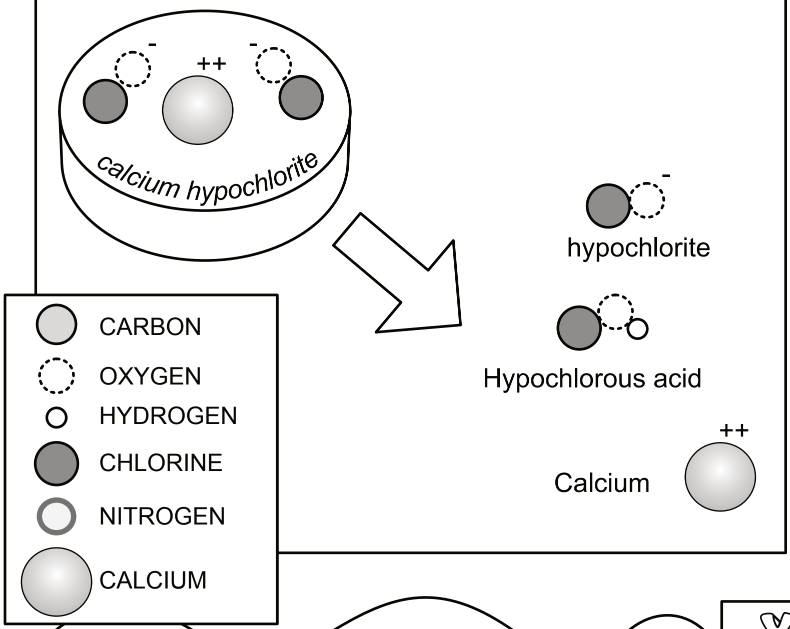
YOUR DOODLE SPACE

Draw your favorite moment from class or write a cool fact!

Your notes: ______

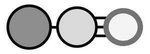
Why do we CHLORINATE SWIMMING POOLS?

HYPOCHLORITE AND HYPOCHLOROUS ACID ARE THE DISINFECTANTS THAT ELIMINATE HARMFUL BACTERIA, ALGAE, AND FUNGAL DISEASES FROM A SWIMMING POOL.

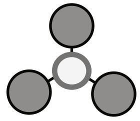


The chemistry of swimming pools is fascinating! To keep a pool safe for people but inhospitable to bacteria and algae, the pH, salts, water hardness, and chlorine levels have to be just right.

As UV light shines down, water evaporates, and people swim in the water, chemical reactions happen and everything changes! You don't need a degree in chemistry to keep your pool healthy – but you do need to check its chemicals frequently to maintain the right balance!



CYANOGEN CHLORIDE



TRICHLORAMINE

NEVER PEE IN A POOL! URINE CONTAINS URIC ACID. THE NITROGEN IN URIC ACID REACTS WITH CHLORINE TO FORM TRICHLORAMINE AND CYANOGEN CHLORIDE, BOTH OF WHICH ARE POISONOUS AND HARMFUL TO YOUR HEALTH.

Your notes: ______________________________

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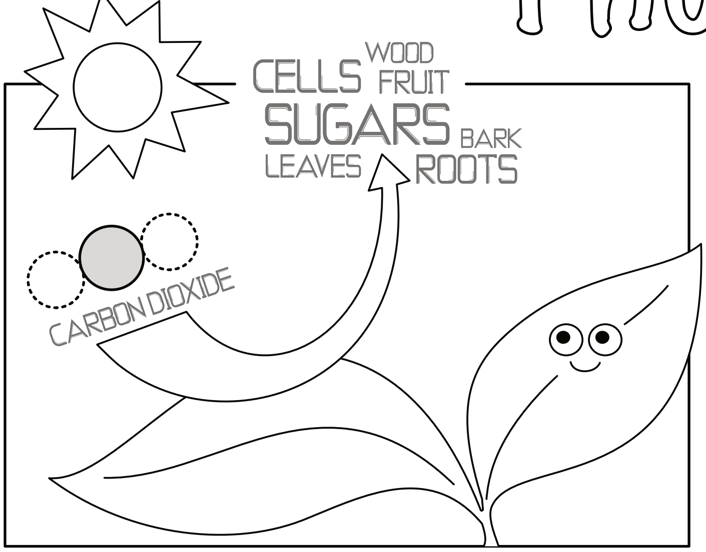
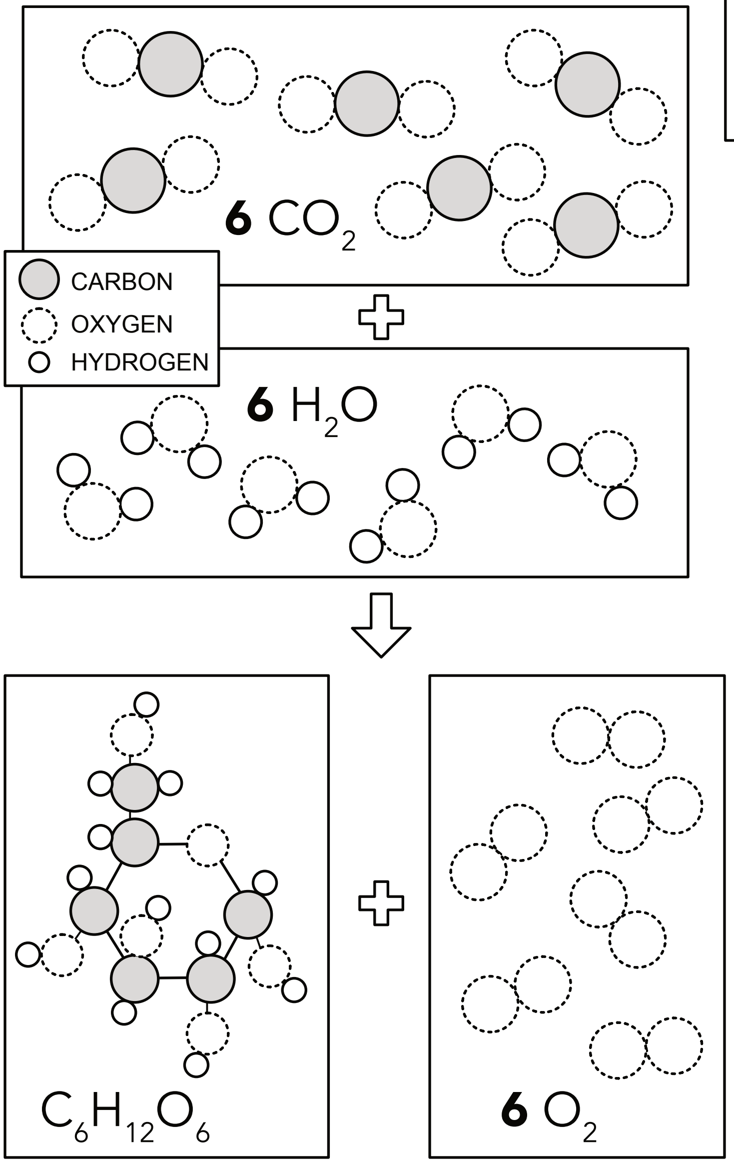
YOUR DOODLE SPACE

Draw your favorite moment from class or write a cool fact!

Photosynthesis



THE CHEMICAL EQUATION FOR
PHOTOSYNTHESIS



YOUR DOODLE SPACE

Draw your favorite moment from class or write a cool fact!

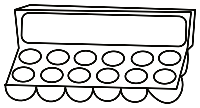
A lot of people think that plants grow out of soil - that atoms in the soil becomes the plant. Actually, most of the plant comes from AIR. More than 98% of the plant's mass comes from carbon dioxide and water.

Your notes: ________________________________

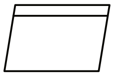
Hands-on Activity

FRANKENSEEDS!

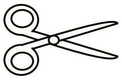
MATERIALS:



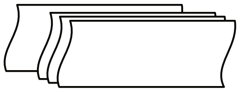
EGG CARTON



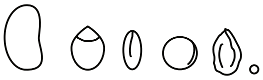
ZIPLOCK BAGS



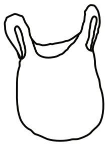
SCISSORS OR KNIFE TO CUT THE EGG CARTON



PAPER TOWELS



AT LEAST 6 TYPES OF SEEDS FROM YOUR KITCHEN



EMPTY BREAD OR PRODUCE BAG

WATER

Choose six types of seeds from your kitchen and make predictions about which ones will sprout and which ones will not. If you expect that they will sprout, draw pictures of what you expect your plants to look like. Will they have two small round leaves or will they look more like a blade of grass? Will the seed split when it germinates? What shape will the leaves have?

1. Cut your egg carton in half so that you have two containers, each with six pockets.
2. Moisten two paper towels with water. Place a wet paper towel along the inside of each of the egg carton halves. Then put your seeds in the cartons. Arrange the cartons to be as identical as possible with three types of each seed in each pocket.
3. Next get two more paper towels wet and place them inside two ziplock bags. Place three of each of your seeds on these paper towels as well.
4. Cover the egg cartons with empty produce bags or bread bags to ensure that they stay moist. Place one in the fridge and one by a window.
5. Tape one of your plastic bags to a window so that it gets some sunlight. Place the other plastic bag in a different location. You can choose to put it in the fridge (cold and dark) or to place it somewhere that has less light.
6. Check on your seeds everyday and record your observations. Make sure the paper towels do not dry out (add water as needed) and that the seeds do not get too wet (they should not be covered in water).
7. After one week, move the seeds that were in the fridge to a location with light and warmer temperatures. After 2 weeks compare your predictions to the results that you observed.

Is it a seed?

	yes	no
Rice		
Popcorn		
Bean		
Basil		
Peppercorn		
Peach pit		
Pinecone		
Cashew		
Peanut		
Tapioca pearls		
Hazelnut		
Almond		
Rosemary		
Nutmeg		
Coconut		
Fig		

IF YOU'RE CARVING PUMPKINS FOR HALLOWEEN, THEIR SEEDS WOULD BE GREAT TO USE IN YOUR EXPERIMENT!

FRANKENSEEDS CONTINUED...

Label your 6 types of seeds A through F. Two or three daya after you plant your seeds, start tracking whether or not they have germinated. Put an x in the box on the first day you see germination (a small rootlet coming from the seed). Draw a leaf on the first day you see green cotyledons or leaves growing from your seed! After two weeks, move the seeds from your fridge to a windowsill. Keep them moist and keep tracking their progress (another piece of paper will be needed to continue your chart).

WEEK ONE

CARTON FROM THE WINDOWSILL:

	SUN	MON	TUE	WED	THURS	FRI	SAT
A							
B							
C							
D							
E							
F							

CARTON FROM THE FRIDGE:

SUN	MON	TUE	WED	THURS	FRI	SAT

WEEK TWO

	SUN	MON	TUE	WED	THURS	FRI	SAT
A							
B							
C							
D							
E							
F							

SUN	MON	TUE	WED	THURS	FRI	SAT

How long did the seeds take to sprout? Which seeds sprouted and which seeds did not? Why do you think the seeds did not sprout?

If the seeds are alive they should sprout within 3-7 days if they are in a moist and warm environment. Some grains (like white rice) have had parts of the seed removed and because of that they cannot sprout. If the seed has been roasted or cooked, it will not sprout.

Why do plants need water?

All living things need water, but plants also use water in photosynthesis.

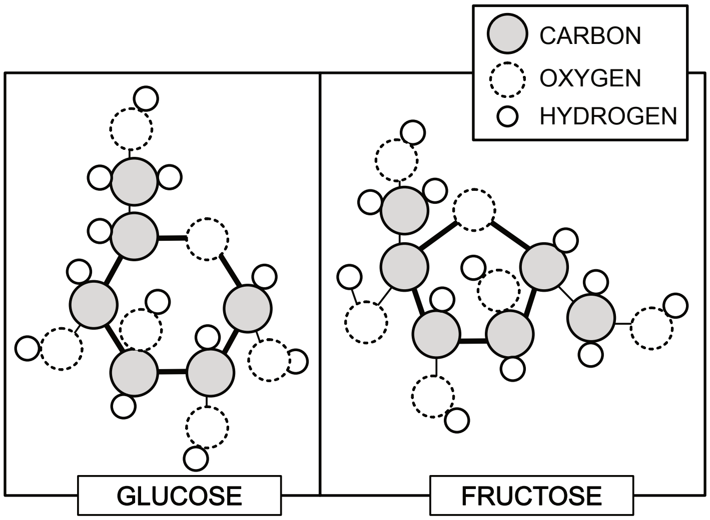
Why do plants need air?

Air is where plants get their carbon from! They convert CO2 into sugars.

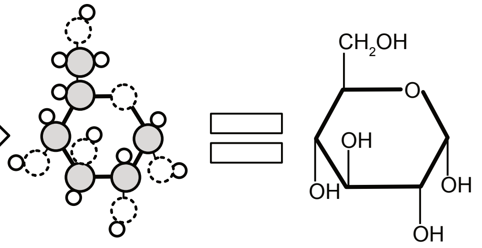
Why do plants need soil?

They actually don't! Most plants can be grown without soil. But the soil provides a place for them to grow as well as nutrients.

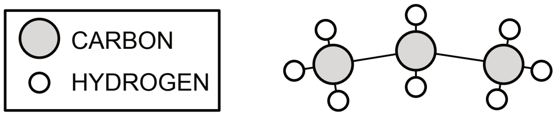
SUGARS

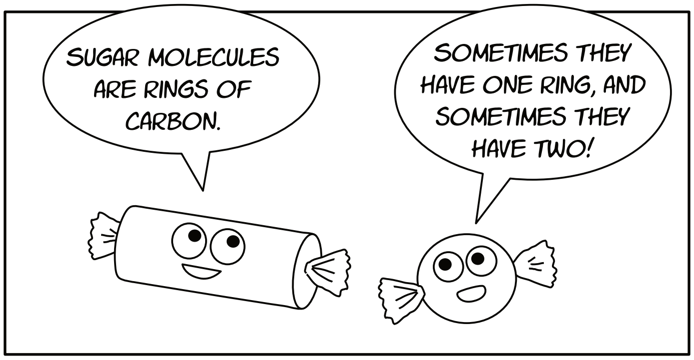


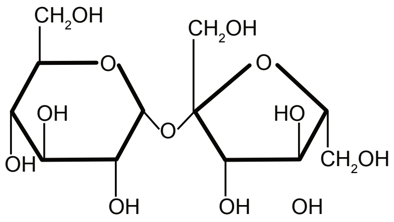
TWO WAYS OF DRAWING THE SAME MOLECULE!



People got tired of drawing out all the hydrogens and carbons, so they came up with a great idea: stick figure carbon chains!  *= 3 carbons & 8 hydrogens. Drawn out with circles for atoms, it would look like this:*







SUCROSE = 1 glucose + 1 fructose

YOUR DOODLE SPACE

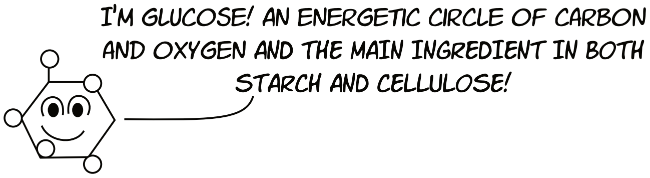
Draw your favorite moment from class or write a cool fact!

HFCS EXPLAINED:

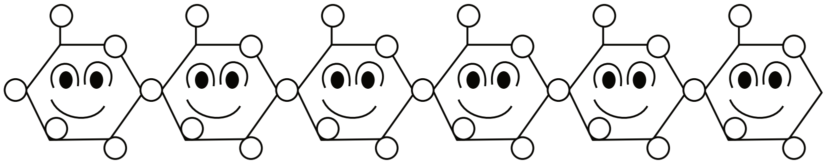
HFCS stands for High Fructose Corn Syrup, an artificial sweetener made by converting glucose to fructose. Why would people designing special chemical reactions to increase the amount of fructose in corn syrup? Because glucose doesn't taste very sweet! Pretty much all of the taste and sweetness of regular sugar (sucrose) comes from the fructose. Increasing the amount of fructose increases the sweetness.

Your notes: ______________________

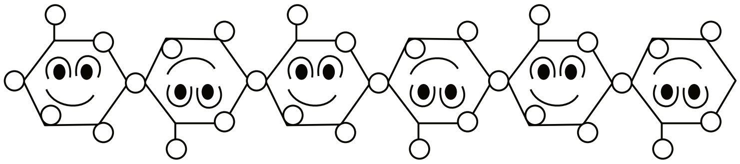
Why can't we eat wood?



Can you spot the difference between starch and cellulose?



String glucose together like this, and you get **starch**—*A big ingredient in things like potatoes and corn and rice and wheat.*



String glucose together like this, and you get **cellulose**—*the main ingredient in things like leaves and straw and wood.*

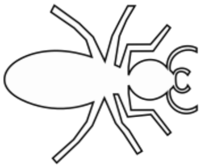
A termite can eat a piece of wood and get energy from it. A cow can eat grass and get energy from that. But if you eat wood or grass it's called *fiber.* Your body can't digest it and it passes straight on through. Have you ever wondered why? Why can you live for weeks on a diet of potatoes, but not newspapers or twigs?

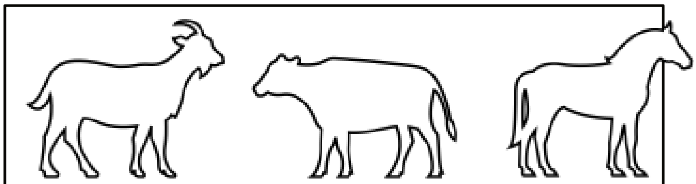
Cellulose and starch are both polymers made of the same building unit: glucose. The difference between them is HOW the glucose molecules are linked together. In starch, all the molecules are facing the same way. We call this an alpha linkage. In cellulose, every other glucose is flipped upside down. We call this a beta linkage. When you eat starch, your body can break that alpha linkage apart so each of your cells can eat the glucose. But beta linkages are tricky. They can only be broken by bacteria and fungi. NOT A SINGLE ANIMAL can do it. So then how in the world do termites eat wood? How do horses cows, goats, and sheep eat grass? (Look at the next page to find out!)

Your notes: ____

The Termite

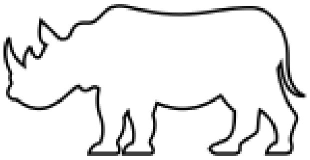
Termites have special bacteria living in their stomachs that digest cellulose for them, breaking it apart into glucose. Give a termite an antibiotic, and it would starve to death no matter how much wood it ate. The termite can only digest wood with the help of its special "termite gut microbes."





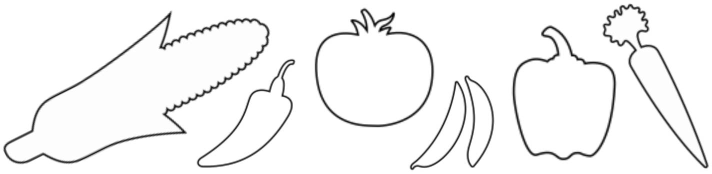
The Herbivores

Herbivores also digest grass with the help of bacteria. Some herbivores (cows) have 4 stomachs to provide even better homes for those important little microbes. Others, like the camel and hippopotamus, have 3 stomachs. And horses have just 1, plus a long "water gut" that provides the perfect place for the bacteria to do their work.



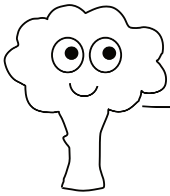
Why fiber is important

We can't digest cellulose, but does that mean we don't want to eat it? Not so fast! If you were able to digest absolutely *everything* you ate, well, that would be a bit of a problem. How would you get rid of things your body didn't want, like extra cholesterol in your blood? If you have enough fiber, the fiber binds to the extra cholesterol and takes it out with the trash. If you don't have enough fiber, the cholesterol is reabsorbed into the bloodstream. Too much cholesterol can cause a heart attack. And that's just one of the many benefits of having enough fiber in your diet.



YOUR DOODLE SPACE

Draw your favorite moment from class or write a cool fact!



CURIOUS WHICH FOODS HAVE THE MOST FIBER? THAT WOULD BE DARK GREEN VEGETABLES, LEGUMES, AND WHOLE GRAINS!

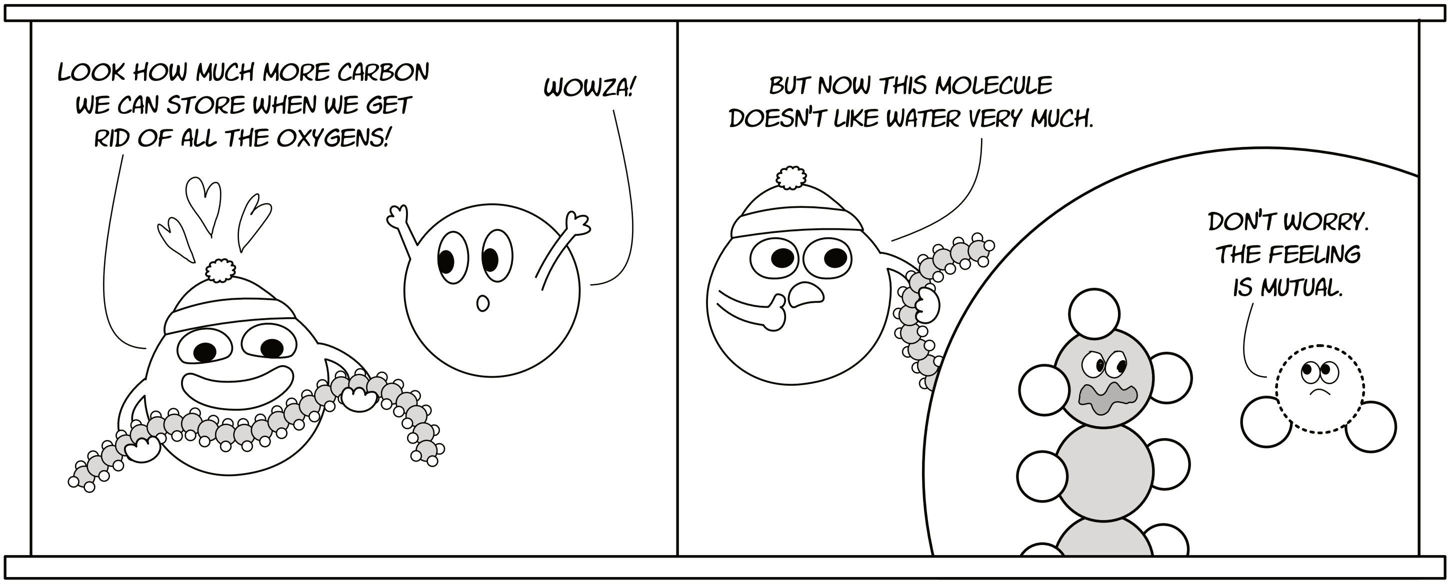
Your notes:

LIPIDS AKA fat & oil

FILL IN THE BLANKS USING THESE WORDS:

oils carbon hydrogen
fats hydrophobic determines

Lipids are fats and OILS. Chemically, they are very long strands of carbon and HYDROGEN. How long the strand is and what types of bonds it has (single or double) DETERMINES what type of oil or fat it is. But all FATS and oils are mostly made of just two atoms: CARBON and hydrogen. Because these long strands don't have any charged areas, they are HYDROPHOBIC which means water fearing. This is why oil and water don't mix together!



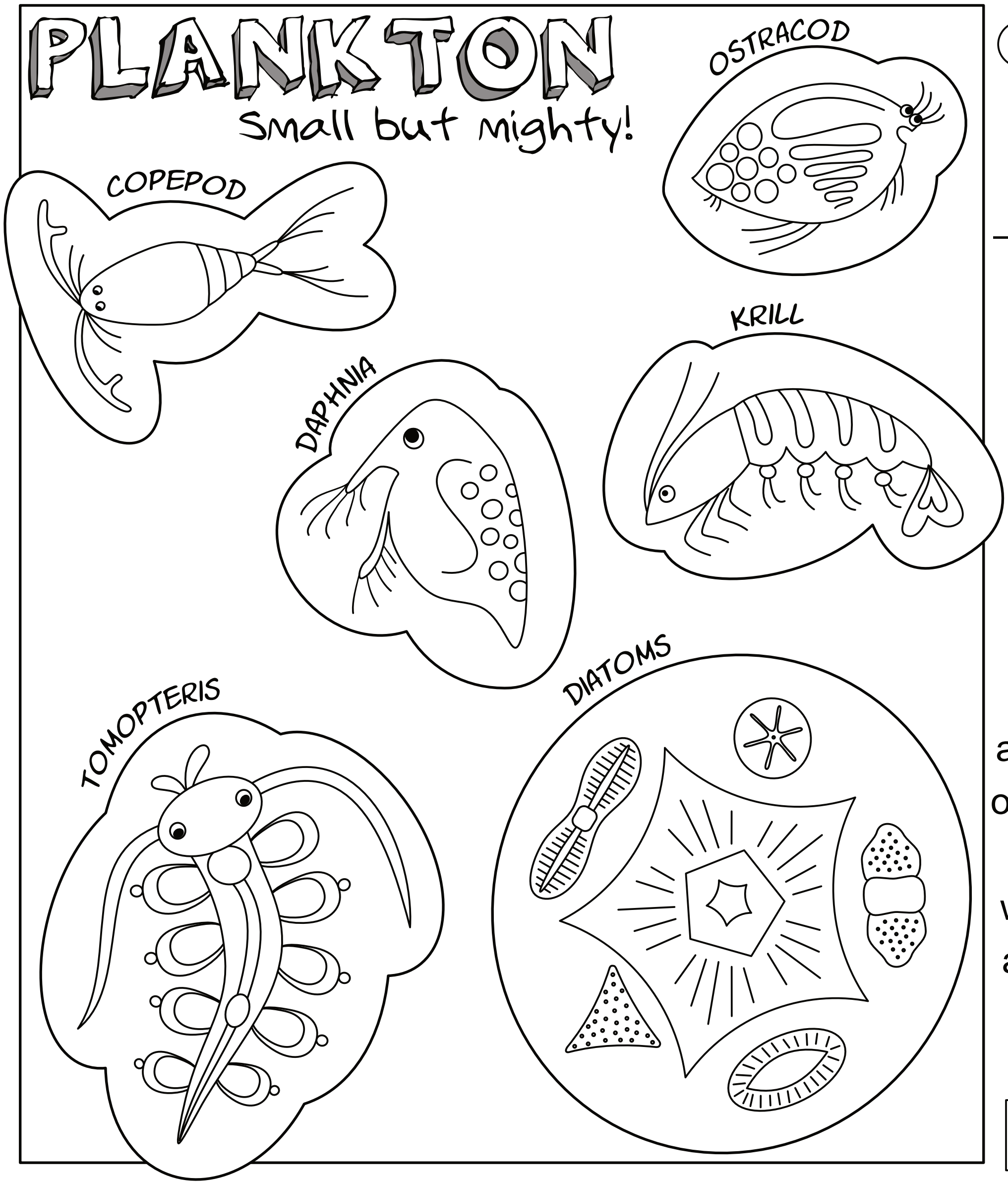
YOUR DOODLE SPACE

Draw your favorite moment from class or write a cool fact!

Your notes: ______________________________

PLANKTON

Small but mighty!



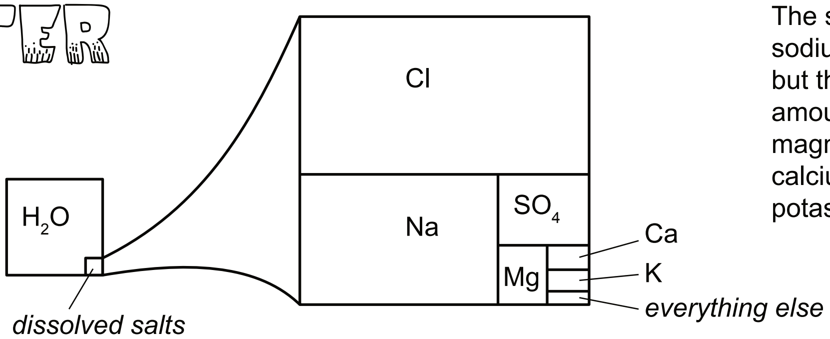
Plankton are the microscopic organisms floating in seawater or __FRESH__ water. They can be plants, animals, viruses, or fungi! __PHYTOPLANKTON__, like diatoms, are photosynthetic. __ZOOPLANKTON__, like copepods or krill, are small animals that eat other plankton. All animals on Earth depend on plankton because they produce most of the world's __OXYGEN__ and absorb large amounts of ____________.

FILL IN THE BLANKS ABOVE USING THESE WORDS:

zooplankton phytoplankton
oxygen fresh CO_2

SEAWATER

96.5% water
3.5% dissolved salts



The salts are mostly sodium chloride (~85%) but there's a decent amount of sulfate (7%), magnesium (4%), calcium (1.2%) and potassium (1.1%) too.

Your notes: ____

____

____

____

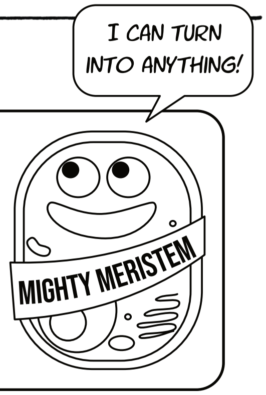
Hands-on Activity

Plant propagation

MATERIALS:

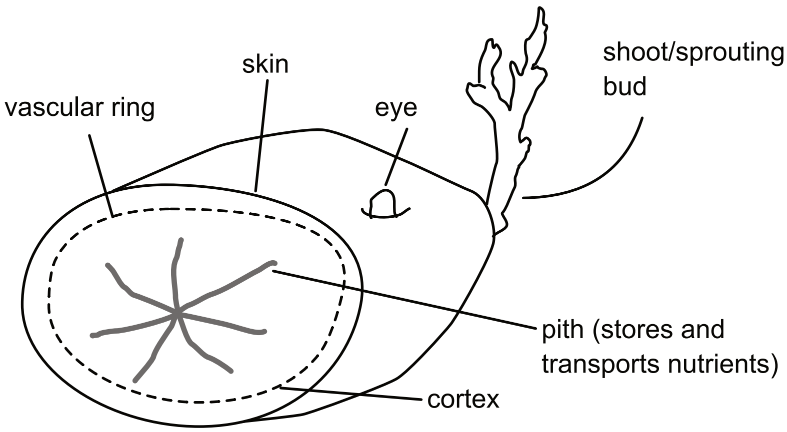
- A TUBER (SUCH AS A POTATO)
- A ROOT VEGETABLE (CARROT)
- A PINEAPPLE
- CUPS
- TOOTHPICKS
- WATER

A leaf is very different than a root, and those differences come from the cells and how they behave. Not all plant cells can grow into roots or leaves or a new plant. But certain cells called meristems can! Try regrowing plants from foods in the kitchen to learn more about these amazing meristems and different types of plants.



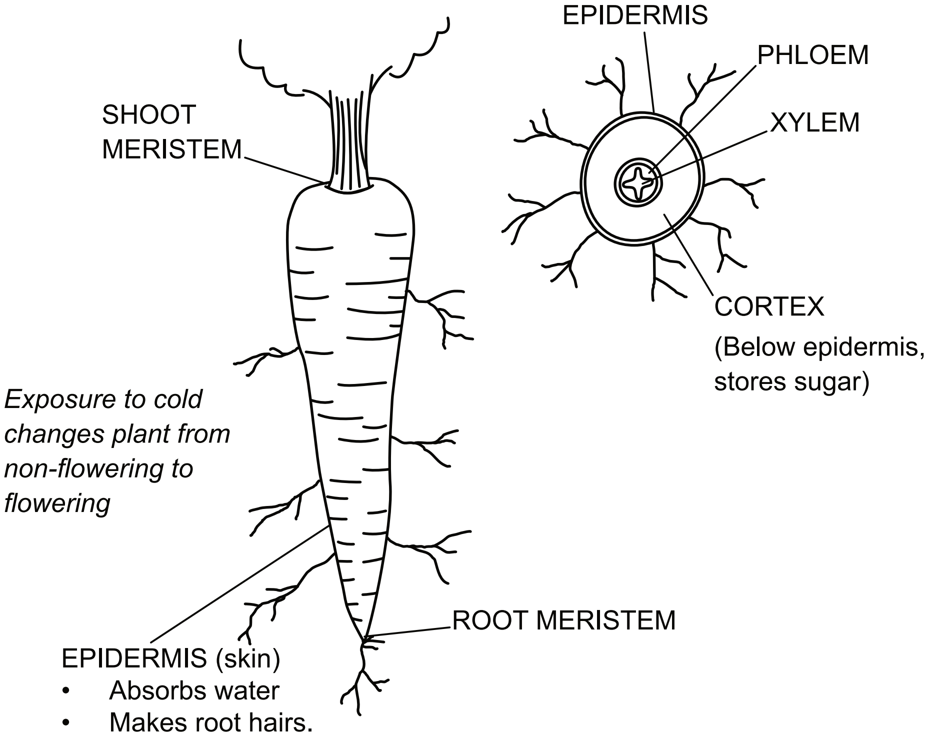
TUBERS

In a tuber like a potato, sweet potato, turmeric, or ginger, there are small "eyes" or nodes where a new shoot and root can grow. If you cut a potato or other tuber and place toothpicks in it so that it is half submerged in water, a shoot and roots will grow from the node.



ROOTS

Root vegetables like a beets, carrots, turnips, rutabagas, parsnips, or onions can regrow from the top of the vegetable. Select a vegetable with some green at the top. Place the top in a cup that is partially filled with water. Use toothpicks to suspend it so that the bottom part is wet but the top is exposed to air. The plant will regrow from the top but it won't grow a new root vegetable -- only the leaves and potentially flowers and seeds.

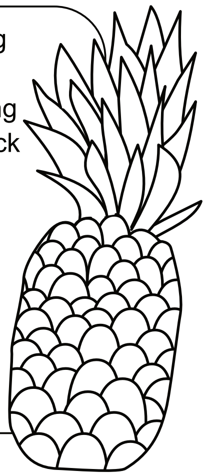


Hands-on Activity

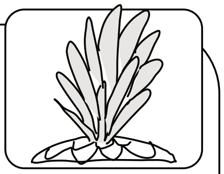
Plant propagation continued...

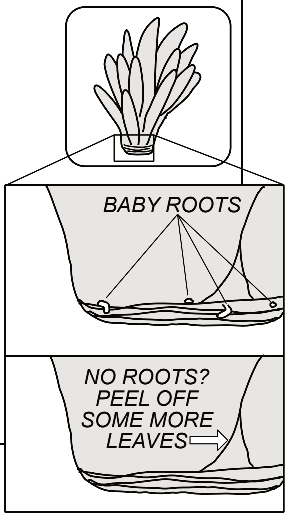
PINEAPPLES!

Pineapples are a very interesting fruit because they have a full miniature pineapple plant growing on top of the fruit! If you peel back the leaves from the top of a pineapple, you will see tiny little rootlets. Place this top of the pineapple into a cup half full of water. Soon roots will emerge. Congrats! Your new pineapple plant is growing.

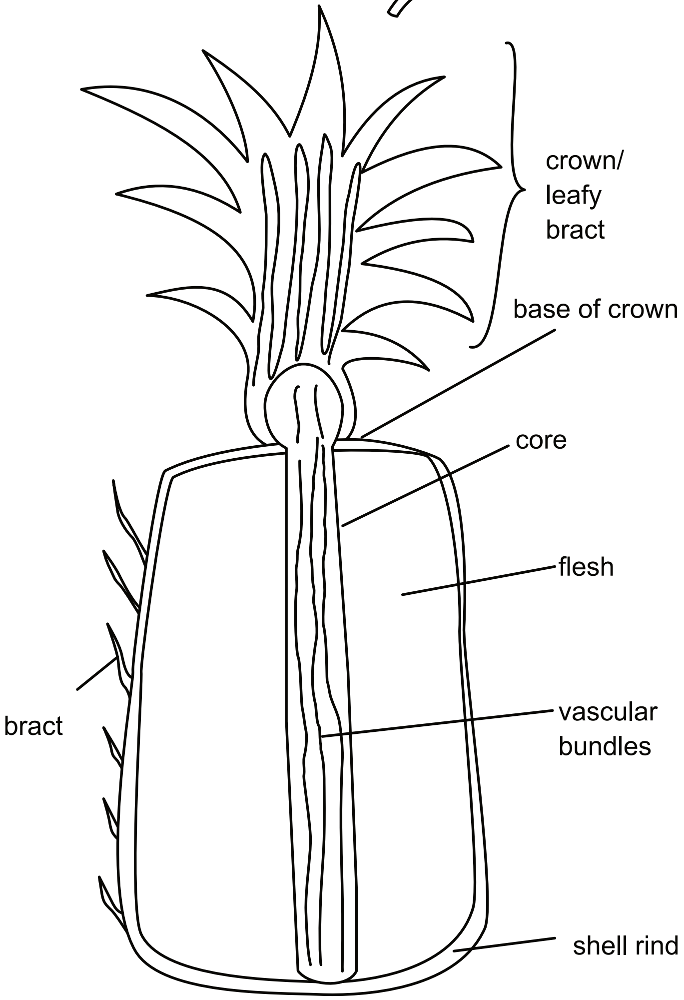


1. Cut off the pineapple top.
2. Pull off fruit and first two layer of leaves from the base.
3. Look closely. Do you see tiny little nubs of roots? Yes? Put your pineapple in a cup with some water. No? Keep removing leaves until you find the small nubs of roots.

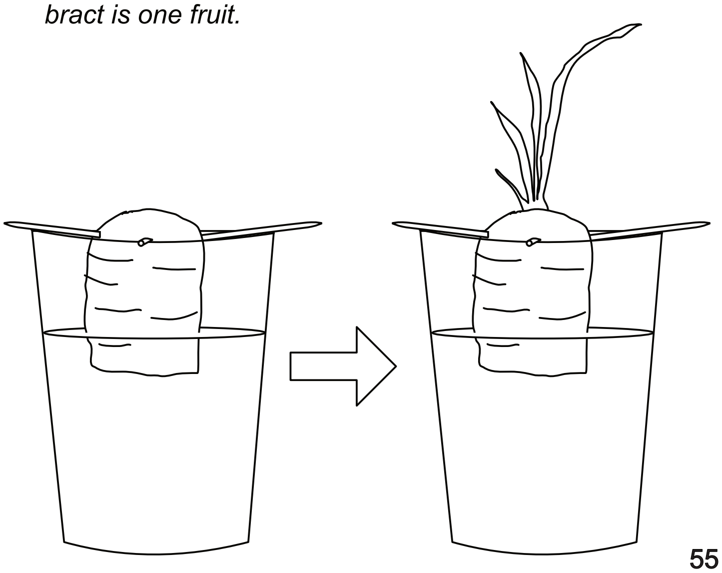




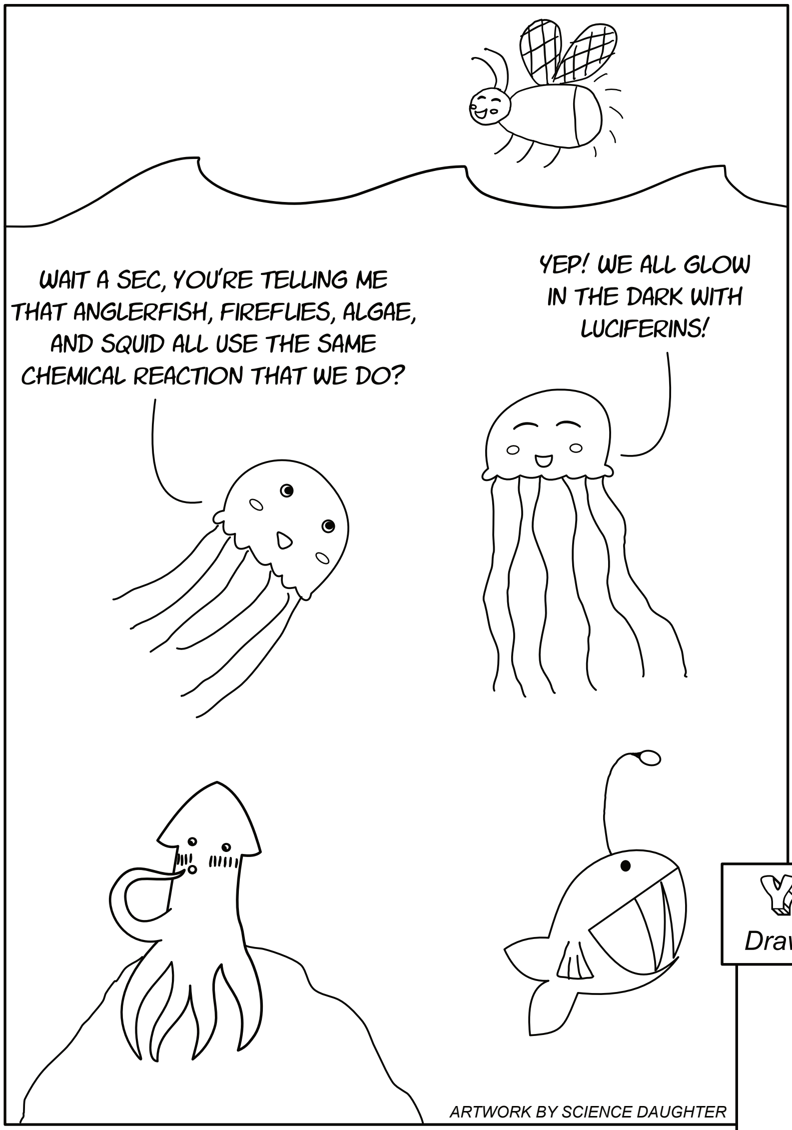
You can also try: an avocado pit, basil leaves, the base of a bok choy, cabbage leaves, celery base, cilantro stems, garlic, green onions, mushroom stocks, onion base, or romaine lettuce base.



A pineapple is a “sorosus.“ That’s greek for heap. It’s a cluster of berries that all grew together! Each bract is one fruit.



BIOLUMINESCENCE



ARTWORK BY SCIENCE DAUGHTER

When living animals make light it is called BIOLUMINESCENCE. Certain types of insects, fish, invertebrates, and algae can create LIGHT using a special CHEMICAL reaction. While there are several different types of chemicals are used, the main idea is that OXYGEN provides energy for a chemical reaction that creates light. The ANIMAL can control when the reaction starts and stops, creating FLASHES of light to communicate, lure in prey, or escape from predators.

FILL IN THE BLANKS ABOVE USING THESE WORDS:

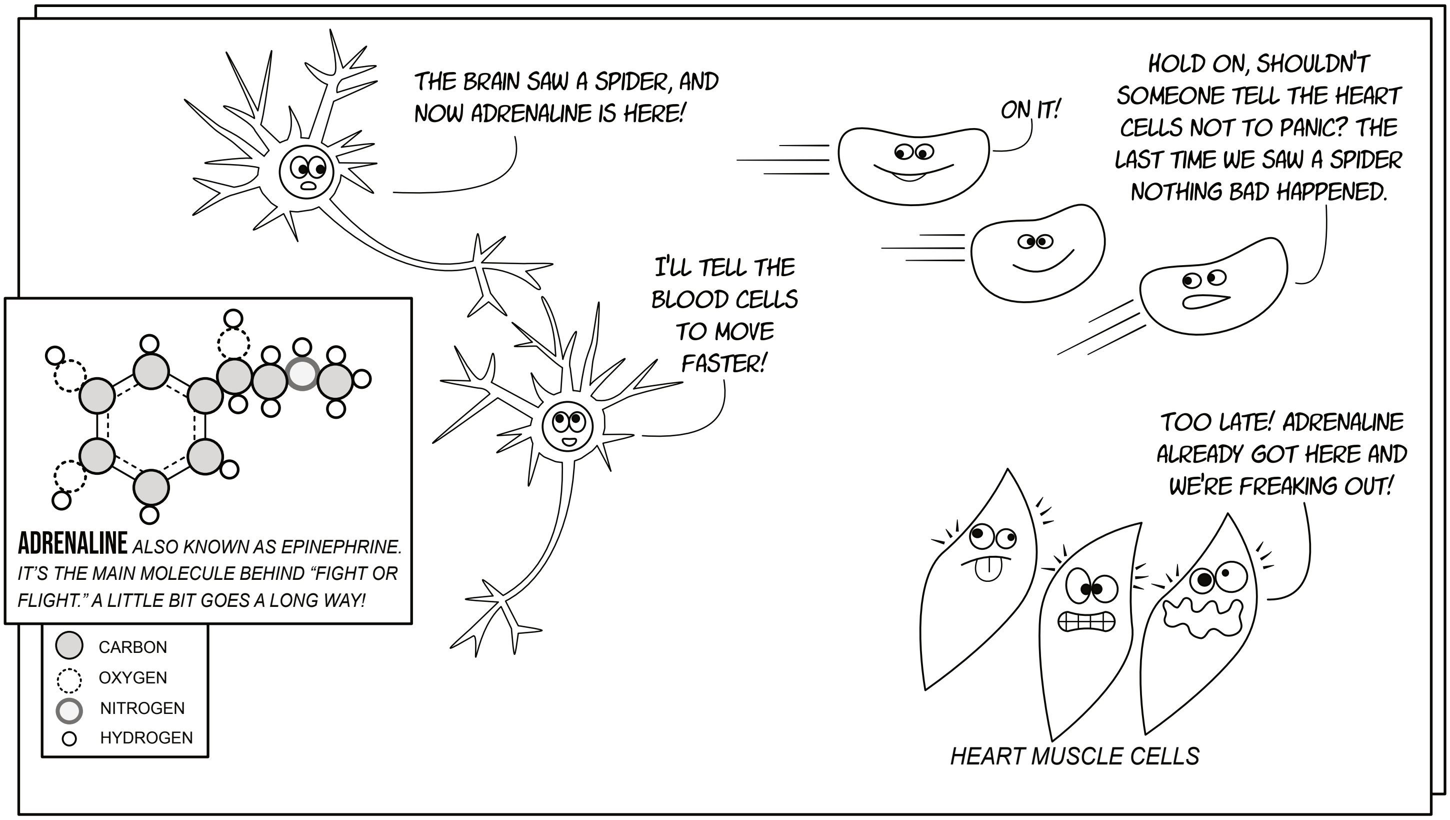
light
animal
bioluminescence
chemical
oxygen
flashes

YOUR DOODLE SPACE

Draw your favorite moment from class or write a cool fact!

Your notes:

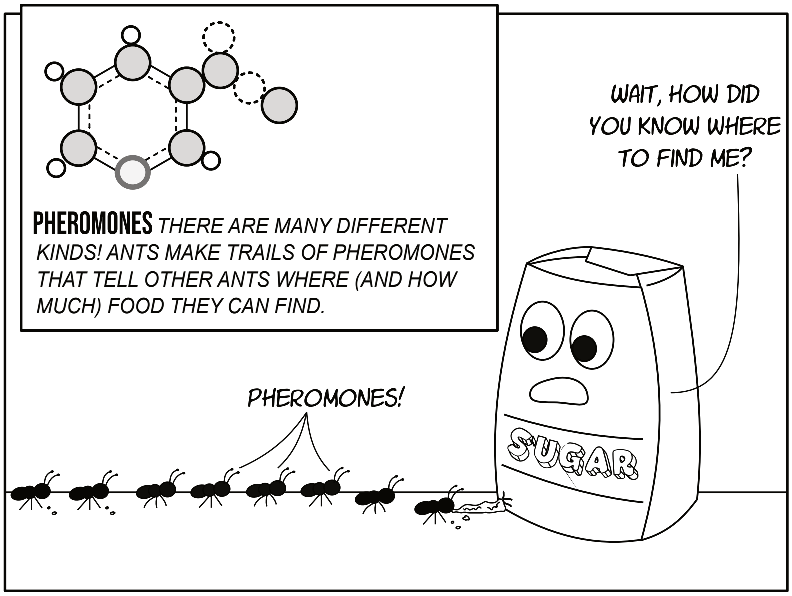
From CELLS to COLONIES



ADRENALINE *ALSO KNOWN AS EPINEPHRINE. IT'S THE MAIN MOLECULE BEHIND "FIGHT OR FLIGHT." A LITTLE BIT GOES A LONG WAY!*

CHEMISTRY IS EVERYWHERE!

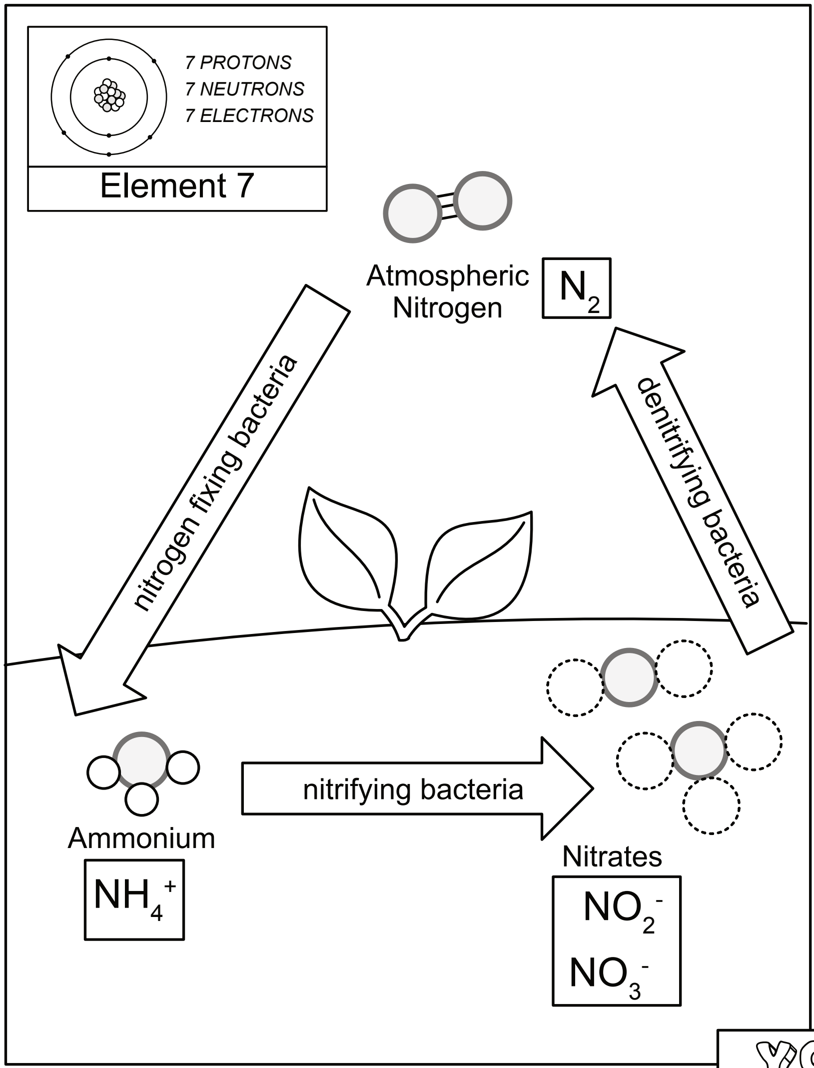
How does a nerve cell learn to send signals, or a blood cell to carry oxygen? How does one ant tell the rest of the colony where to find food? The answer to both questions comes down to chemistry! Chemical reactions power the life of the cell, and control how it communicates with other cells. And it's not just cells that communicate with chemicals - insect colonies do too!



PHEROMONES *THERE ARE MANY DIFFERENT KINDS! ANTS MAKE TRAILS OF PHEROMONES THAT TELL OTHER ANTS WHERE (AND HOW MUCH) FOOD THEY CAN FIND.*

Your notes:

NITROGEN Essential nutrient & most of our air



Nitrogen is an essential element used to make PROTEINS and DNA. Every ANIMAL and plant needs it, and 78% of Earth's ATMOSPHERE is nitrogen. So you might think that it would be easy to get, but the nitrogen in the air is N_2. It's two atoms bound with a very strong TRIPLE bond and that bond is very hard to break! No animals can do it. No plants can do it. Only BACTERIA can change atmospheric nitrogen into a form that PLANTS and animals can use. We call this "fixing" nitrogen.

FILL IN THE BLANKS ABOVE USING THESE WORDS:

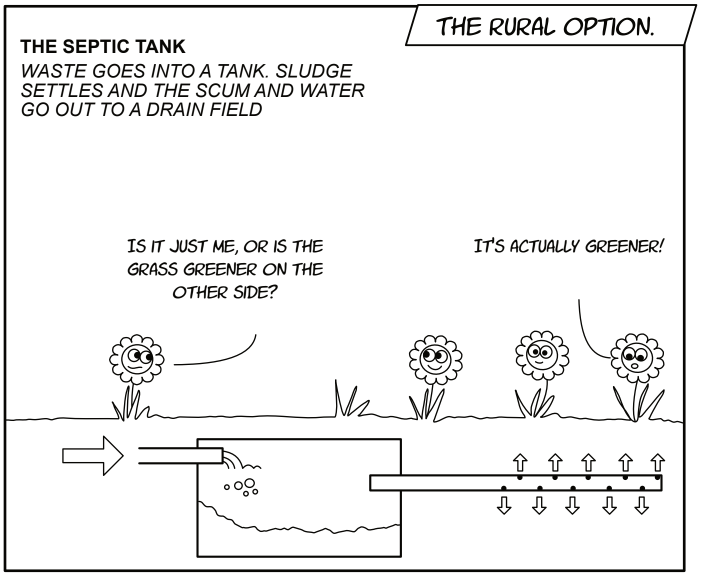
bacteria
proteins
animal
triple
plants
atmosphere

Your notes: ____________________

YOUR DOODLE SPACE

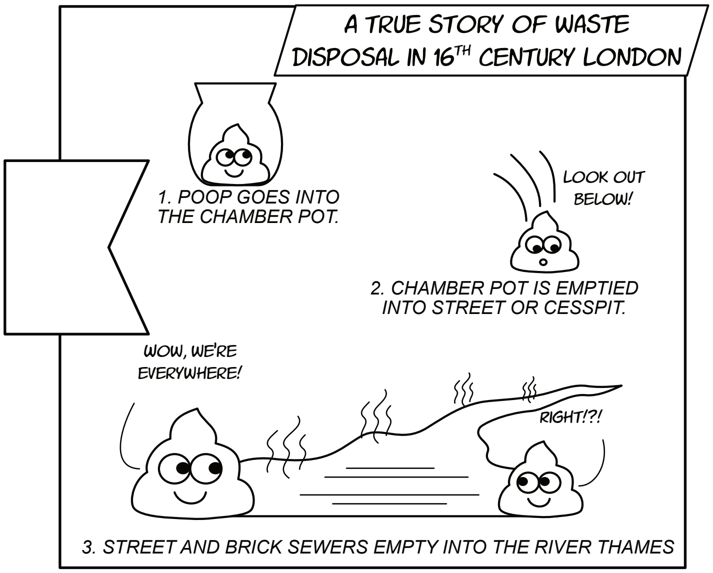
Draw your favorite moment from class or write a cool fact!

Water Reclamation



THE SEPTIC TANK

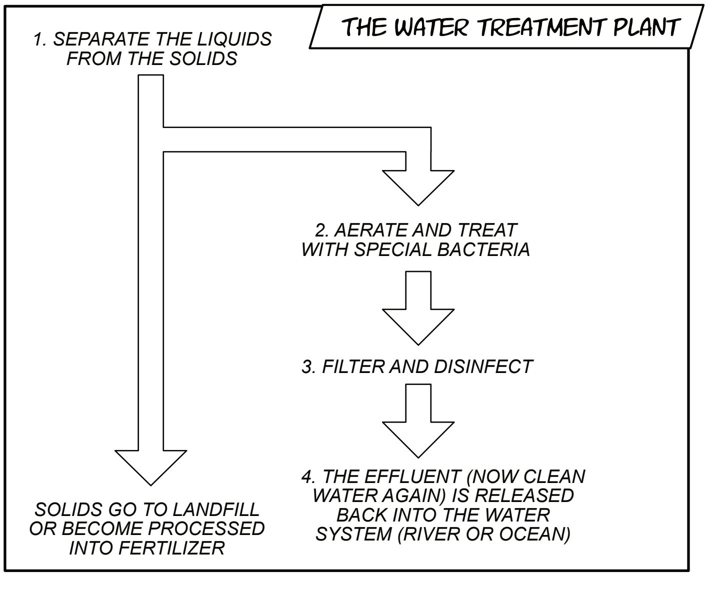
WASTE GOES INTO A TANK. SLUDGE SETTLES AND THE SCUM AND WATER GO OUT TO A DRAIN FIELD



1. POOP GOES INTO THE CHAMBER POT.

2. CHAMBER POT IS EMPTIED INTO STREET OR CESSPIT.

3. STREET AND BRICK SEWERS EMPTY INTO THE RIVER THAMES



1. SEPARATE THE LIQUIDS FROM THE SOLIDS

2. AERATE AND TREAT WITH SPECIAL BACTERIA

3. FILTER AND DISINFECT

4. THE EFFLUENT (NOW CLEAN WATER AGAIN) IS RELEASED BACK INTO THE WATER SYSTEM (RIVER OR OCEAN)

SOLIDS GO TO LANDFILL OR BECOME PROCESSED INTO FERTILIZER

YOUR DOODLE SPACE

Draw your favorite moment from class or write a cool fact!

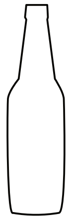
Water is our most precious and interconnected resource.

Your notes: ____________________

Hands-on Activity

DIY water filter

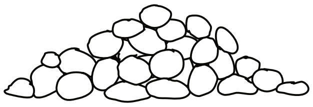
MATERIALS:



A PLASTIC BOTTLE (16 OR 12 OUNCES)

COTTON BALLS OR A PIECE OF COTTON FABRIC

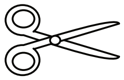
WATER



½ CUP GRAVEL

¼ CUP ACTIVATED CHARCOAL

2 COFFEE FILTERS

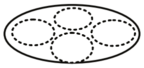


SCISSORS OR KNIFE TO CUT THE PLASTIC BOTTLE

½ CUP SAND

INSTRUCTIONS:

1. Carefully use scissors or knife to cut off the bottom of the water bottle to create a tall funnel.
2. Trim one of the coffee filters into four smaller circles.



3. Place one or two cotton balls on top of two of the coffee filter circles. Carefully place them in the neck of the water bottle. If they flip and turn sideways simply turn the bottle upside down and shake them back out and then try again. It may help to use a chopstick or wooden skewer or straw.
4. Once you have your coffee filter circles and cotton balls in place, put the remaining coffee filter circles on top to make a "coffee filter cotton ball sandwich." This is the lowest layer of filtration.
5. Next, carefully pour 1/4 cup of activated charcoal onto a coffee filter and lower it into the bottle. Then fold the coffee filter over the top of the charcoal to completely enclose it.

   Run a little bit of water through the filter to help the two lower layers compress and make sure that they are pressed against the sides of the bottle.

   Next add 1/2 cup of sand, then add the final layer of 1/2 cup of gravel.
6. Experiment by running different liquids through your filter. Start with relatively clean water such as the leftover water from cooking vegetables. If you run it through your filter, does it still smell like vegetables or have color to it? Or did the filter clean the water?
7. Next, add some food coloring to your water or go outside and get some mud. See how your filter does cleaning that water.

> *Warning! Only drink water that you know is safe to drink!*
> *While this filter is similar to modern filtration systems, it is small enough that contaminants can overwhelmed it and "sneak" through.*

Your observations: ____________________

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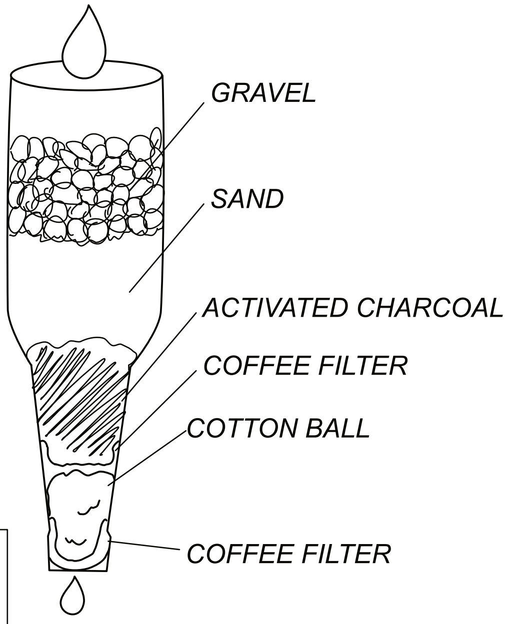
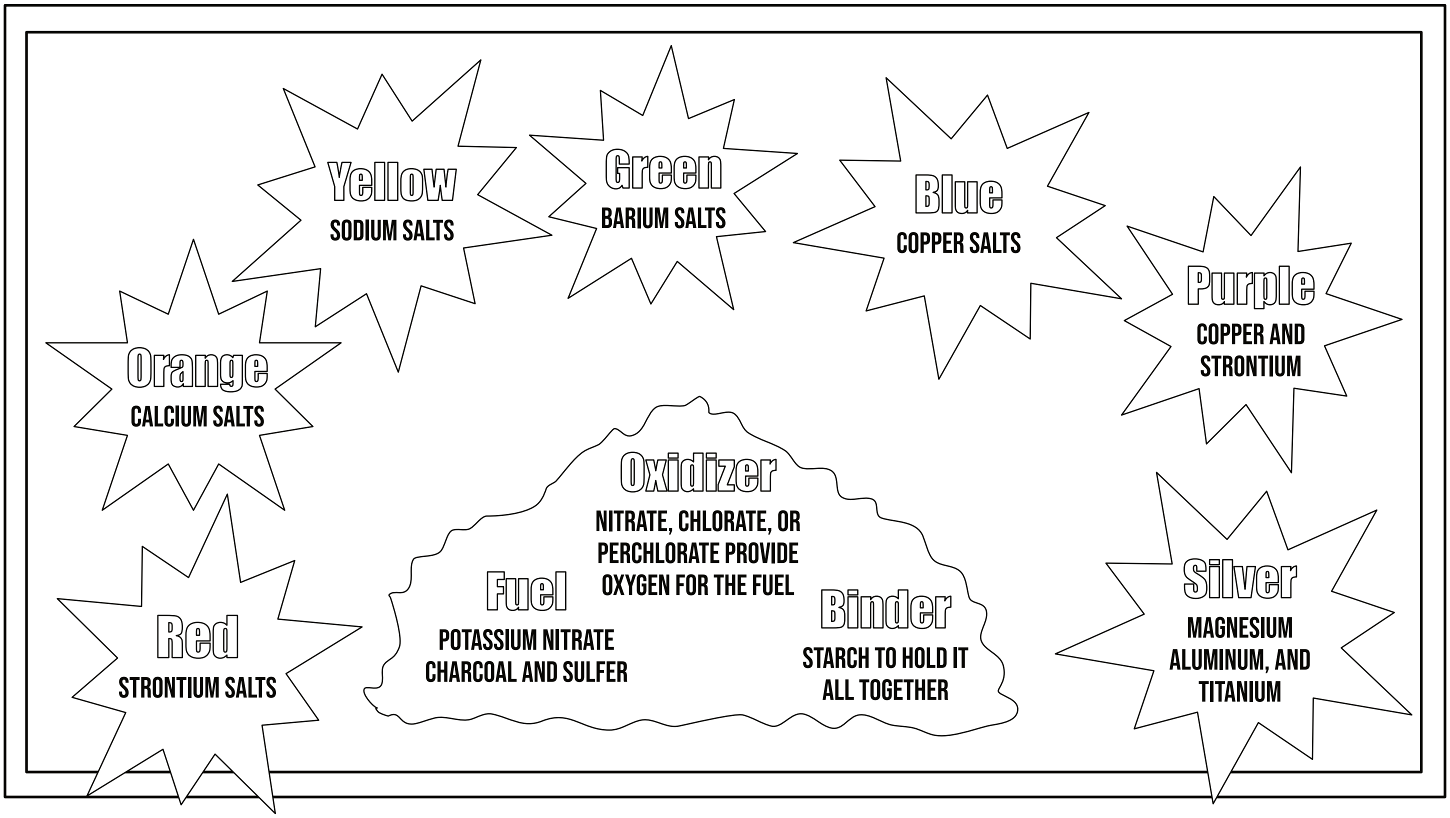


FIREWORKS and lab safety



FILL IN THE BLANKS USING THESE WORDS:

safety chemical reactions
pressure safe

Fireworks are controlled CHEMICAL REACTIONS. These explosive devices delight us with their bright colors on holidays around the world, and they're also a good reminder of the importance of SAFETY precautions. Can chemistry be a lot of fun? Absolutely. Can a little knowledge be a dangerous thing? Sometimes! Make sure you think ahead about what might happen during a reaction. If your future chemistry experiment will produce a gas, be extra careful because PRESSURE might build up. Always wear safety glasses, and make sure to clean up after yourself when your experiments are done! Keeping your laboratory space clean and organized isn't just good manners. It keeps you and your equipments SAFE.

Your notes:

The end

of our course... Hopefully the beginning of many more adventures in science!

We hope you enjoyed this chemistry course! These doodle notes were all drawn by Science Mom (with help from Math Dad, Science Daughter, and Science Moms Liza, Krista, and Emily). If you enjoyed this course, we think you'd also enjoy Theodore Gray's three books: Elements, Reactions, and Molecules.

Last but not least, we have two "go the extra mile" activities, which you'll see on the next few pages. If you complete either of these activities, take a picture of your work and send it to us at jenny@science.mom or tag us on social media.

Twitter: @jennyballif
Facebook: @TheScienceMom
Instagram: @the.science.mom

Work hard, grow smart, and stay curious!
-Science Mom

MEMORIZE THE PERIODIC TABLE CHALLENGE

Can you memorize the periodic table? Practice filling in this chart (print extra copies!) and see if you can learn the entire table. Color in the different families too!

1																	2
3	4											5	6	7	8	9	10
11	12											13	14	15	16	17	18
19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36
37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54
55	56		72	73	74	75	76	77	78	79	80	81	82	83	84	85	86
87	88		104	105	106	107	108	109	110	111	112	113	114	115	116	117	118

57	58	59	60	61	62	63	64	65	66	67	68	69	70	71
89	90	91	92	93	94	95	96	97	98	99	100	101	102	103

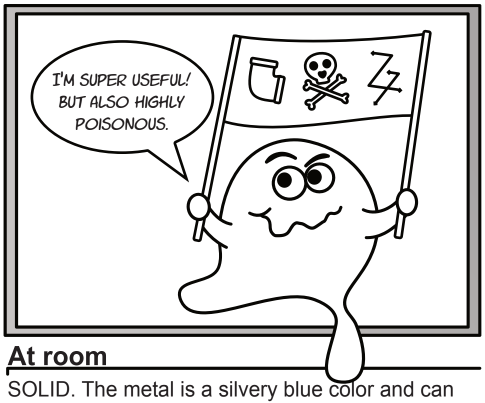
☐ **Nonmetals**
☐ **Alkali Metals**
☐ **Alkali Earth metals**
☐ **Transition metals**
☐ **Metals**

☐ **Metaloids**
☐ **Halogens**
☐ **Noble gases**
☐ **Transactinides**
☐ **Lanthanides**
☐ **Actinides**

MAKE A FULL DECK OF ELEMENTAL TRADING CARDS

LEAD

Pb
82



At room
SOLID. The metal is a silvery blue color and can be either shiny or dull.

Hazard
Caution! Poisonous. Can cause permanent nerve and brain damage. Take care that old pipes (which contain lead) don't leach it into the water.

Other: Used in plumbing, bullets, and to make radiation shields. Few things are as good as lead at absorbing dangerous radiation.

Remember the element trading cards from page 13 and 14? You made 4 of them, now here's a super challenge. Can you create a FULL DECK with all 118 known elements?

Print out extra copies of these templates or make your own! If you complete this epic challenge, email us. We'd love to see your work!

At room temp:

Hazard rating:

Other:

At room temp:

Hazard rating:

Other:

At room temp:

Hazard rating:

Other:

At room temp:

Hazard rating:

Other:

At room temp:

Hazard rating:

Other:

At room temp:

Hazard rating:

Other:

At room temp:

Hazard rating:

Other:

At room temp:

Hazard rating:

Other:

At room temp:

Hazard rating:

Other:

At room temp:

Hazard rating:

Other: