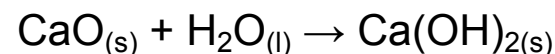




EXAMPLE GIBBS CALCULATIONS

1. Calcium Oxide reacts with water to form Calcium Hydroxide



Thermodynamic Data:	$\Delta H_f / \text{kJ.mol}^{-1}$	$S / \text{J.K.mol}^{-1}$
$\text{CaO}_{(s)}$	-636.5	39.7
$\text{H}_2\text{O}_{(l)}$	-285.9	70.0
$\text{Ca(OH)}_{2(s)}$	-986.6	76.1

a) Calculate the Enthalpy changes for the reactions.

b) Calculate the Entropy Change for the reactions.

c) At what temperature would this reaction **NOT** be feasible?