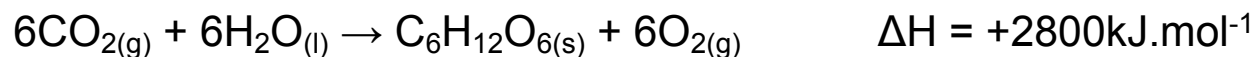




EXAMPLE GIBBS CALCULATIONS

2. Plants are able to produce Glucose from Carbon Dioxide and water.



	$\text{CO}_{2(g)}$	$\text{H}_2\text{O}_{(l)}$	$\text{C}_6\text{H}_{12}\text{O}_{6(s)}$	$\text{O}_{2(g)}$
S / J.K.mol ⁻¹	214	70	218	205

- a) Calculate the Entropy Change for the reaction.
- b) Calculate ΔG for the reaction at 298K.
- c) Explain why this reaction is **NOT** feasible at **ANY** temperature.